



CHEMICAL REACTIONS

11. Chemical Reactions

All living beings born in this beautiful world have their own life styles. Have you observed and analyzed your daily life from the view point of a chemist? Chemical reactions happen around us all the time and even in our body.

Any change can be classified as physical change and chemical change. Physical changes can be easily reversed but, it is not easy to reverse a chemical change. What is the reason? In chemical changes, new substances are formed and it is difficult to regenerate the original substances. Chemical changes are more permanent than physical changes. All chemical changes are accompanied by chemical reactions.

How do we come to know that a chemical reaction has taken place? Let us perform some activities to find out the answer to this question.

ACTIVITY 11.1

- Look at the new silver anklet of your mother or sister
- Note the colour of the anklet
- Observe the colour of an old anklet
- What change do you observe?

The lustrous white colour of the silver anklet slowly changes into slightly black colour. That is, silver anklet has got



Fig. 11.1 Silver Anklet

tarnished. Can you guess the reason behind it?

It is due to the formation of silver sulphide (Ag_2S), as a result of the reaction between silver and hydrogen sulphide in the air.

ACTIVITY 11.2

- Take lead nitrate solution in a beaker
- Take potassium iodide solution in a test tube. (Both solutions are colourless)
- Add potassium iodide solution slowly to the lead nitrate solution
- What do you observe?

You observe a deep yellow precipitate, don't you?

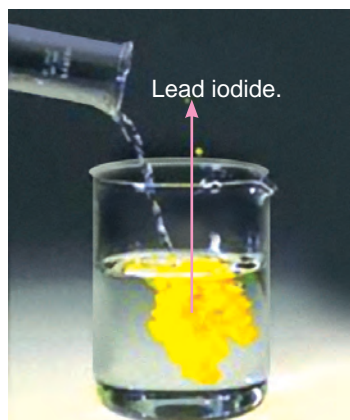


Fig. 11.2 Yellow precipitate of lead iodide.

It is lead iodide (PbI_2).

ACTIVITY 11.3

- Take 5g of calcium oxide (quick lime) in a beaker
- Add water to it slowly
- Touch the beaker
- What do you feel?

Do you feel hot? Let us see what happens

Calcium oxide reacts with water to produce slaked lime (calcium hydroxide). This reaction is exothermic and will be accompanied by hissing sound and bubbles leading to the release of considerable amount of heat.

ACTIVITY 11.4

- Take a pinch of calcium carbonate powder in a test tube
- Add dilute hydrochloric acid
- Note the changes in the test tube carefully

Do you observe any brisk effervescence? It is due to the evolution of carbon dioxide gas.



Fig. 11.3 Reaction of calcium carbonate with dil.HCl

These are some of the common observations in a chemical reaction. From the activities that we have discussed, it is clear that chemical reactions will bring about a permanent change resulting in the formation of new product(s).

The substances taking part in the reaction are known as reactants and those formed as a result of the reaction are called products.

MORE TO KNOW

A solution of slaked lime produced in the Activity 11.3 is used for white washing. Calcium hydroxide reacts slowly with carbon dioxide in air to form a thin layer of calcium carbonate on the walls. Calcium carbonate is formed after two to three days of white washing and gives a shiny finish to the walls. It is interesting to note that the chemical formula for marble is also $CaCO_3$.

11.1. TYPES OF CHEMICAL REACTIONS

Since there are numerous chemical reactions, the study of these reactions can be made easier by classifying them. All the chemical reactions are classified into six broad categories depending on the way the product formed.

Let us see the different types of classifications of chemical reactions.

1. COMBINATION REACTION



A combines with B to form a new product AB. It is the simple representation of combination reaction.

ACTIVITY 11.5

- Take a clean piece of magnesium ribbon
- Hold the ribbon with a pair of tongs
- Burn it in air using a burner (keeping Mg ribbon as far as possible from your eyes)
- Collect the ash

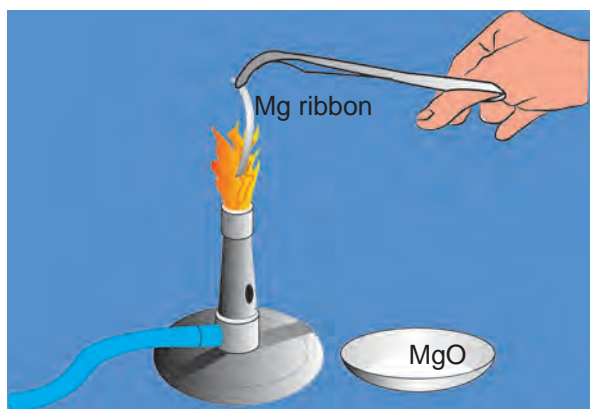


Fig. 11.4 Burning of Mg ribbon

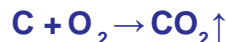
In the above activity, magnesium combines with oxygen to form a single product, magnesium oxide. **Such a reaction in which a single product formed from two or more reactants is known as combination reaction.**



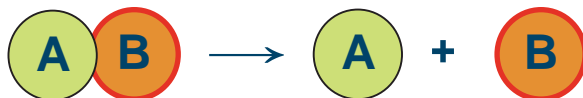
Repeat “Activity 11.3”. This reaction is also an example for COMBINATION REACTION. Attempt to write the equation yourself.

Let us discuss some more examples of combination reactions.

- Combustion of coal
- Combustion of hydrogen



2 DECOMPOSITION REACTION



AB splits into A and B. It is the representation of decomposition reaction.

ACTIVITY 11.6

- Take about 2 g of copper carbonate powder in a dry test tube
- Note the colour of copper carbonate
- Heat the test tube over the flame
- Observe the change after heating

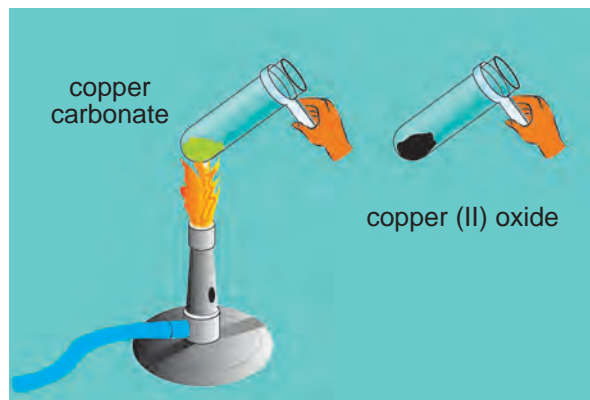


Fig. 11.5 Heating the test tube containing copper carbonate

Change of colour from green to black is observed. This is due to the decomposition of copper carbonate to copper (II) oxide.



ACTIVITY 11.7

- Take lead nitrate in a test tube
- Heat it over the flame
- Observe the changes

Liberation of a reddish brown gas (NO_2) is observed. This is because of the decomposition of lead nitrate into lead oxide, nitrogen dioxide and oxygen.



From the above two activities (11.6 and 11.7), It can be noted that a **single compound breaks down to produce two or more substances**. Such type of reaction is called **decomposition reaction**.

Some other examples for decomposition reaction:

1. Decomposition of lime stone



2. Decomposition of ammonium dichromate



MORE TO KNOW

At very high temperature, ammonium dichromate decomposes immediately to green vapours which gets released along with the steam. It seems as if a volcano erupts and is termed as chemical volcano.

3. DISPLACEMENT REACTION



In the reaction between **A** and **BC**, **A** displaces **B** from **BC** to form **AC**. This shows that **A** is more reactive than **B**.

ACTIVITY 11.8

- Take 20 ml of copper sulphate solution in a beaker
- Drop an iron nail into the beaker
- Leave it for few days
- Observe the colour of the copper sulphate solution and the iron nail

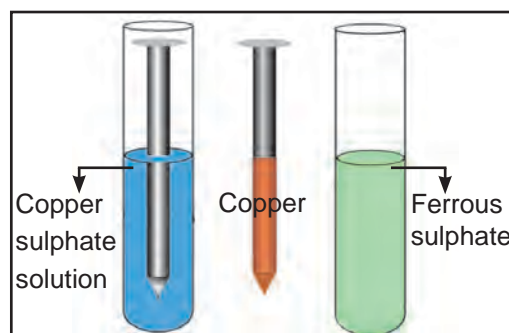
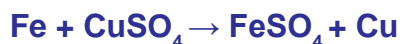


Fig. 11.6 Iron displaces copper from copper sulphate solution

Blue colour of the copper sulphate solution changes into green colour and the iron nail acquires a brownish look. It is a noticeable change. Is it not? This change confirms that iron is more reactive than copper. The following chemical reaction takes place in this activity.



In this reaction, iron displaces copper from CuSO_4 solution.

Repeat "Activity 11.8" but use zinc rod instead of an iron nail. What colour changes do you observe on the rod and in the solution? Write the chemical equation.

Other example:



Lead can displace copper from its salt solutions. Can copper displace zinc or lead from their salt solutions? No, because copper is less reactive than zinc and lead.

The reaction in which, a more reactive element displaces a less reactive element from its compound is called displacement reaction.

4. DOUBLE DECOMPOSITION REACTION (DOUBLE DISPLACEMENT REACTION)



In the reaction between **AB** and **CD**, both the reactants decompose to form **AD** and **CB** through the rearrangement of ions.

ACTIVITY 11.9

- Take 5ml of sodium sulphate solution in a test tube
- In another test tube, take 5ml of barium chloride
- Mix both the solutions
- What do you observe?

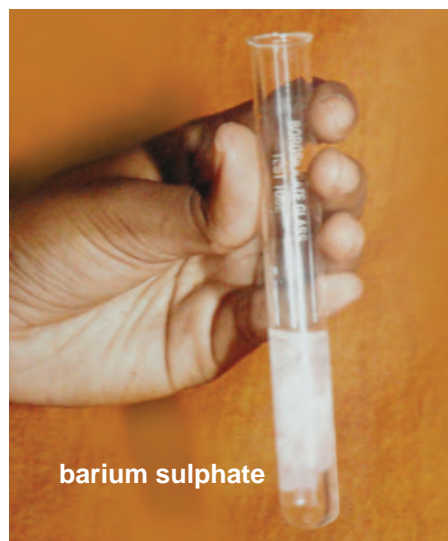


Fig. 11.7 Formation of barium sulphate

You will observe formation of a white substance, which is insoluble in water. The insoluble substance formed is known as *precipitate*. Any reaction that produces a precipitate is called a **precipitation reaction**. The formed white precipitate of barium sulphate, is due to the reaction of SO_4^{2-} and Ba^{2+} ions. The other product formed is sodium chloride.



Repeat "Activity 11.2" for double decomposition reaction. Attempt to write the equation by yourself.

Double decomposition reaction is any reaction in which exchange of ions between two reactants occur, leading to the formation of two different products.

Other example :



5. OXIDATION AND REDUCTION

We are all aware of the fact that oxygen is the most essential element for sustaining life. One can live without food or even water for a number of days, but not without oxygen. In our daily life we come across phenomena like fading of the colours of the clothes, burning of combustible substances like cooking gas, wood and coal, and also rusting of iron articles. All such processes fall in the category of a specific type of chemical reaction called oxidation – reduction reaction (redox reaction). A large number of industrial processes like electroplating, extraction of metals like aluminium, are based upon the redox reaction.

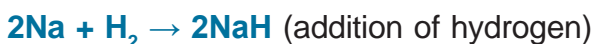
Oxidation:

A chemical reaction which involves addition of oxygen or removal of hydrogen or loss of electron(s) is called as oxidation.



Reduction:

A chemical reaction which involves addition of hydrogen or removal of oxygen or gain of electron(s) is called as reduction.



Redox reaction:

A chemical reaction in which oxidation and reduction take place simultaneously is called redox reaction.



Attempt to write any other redox reaction

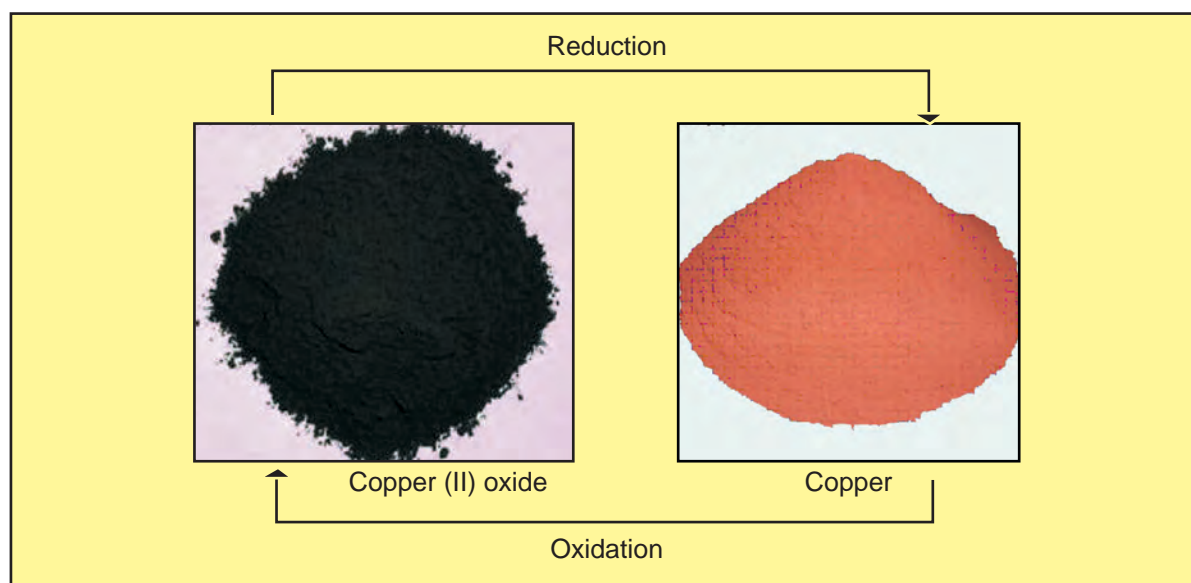


Fig. 11.8 Redox reaction

During the conversion of copper(II) oxide to copper, the copper(II) oxide is losing oxygen and is being reduced. The hydrogen is gaining oxygen and is being oxidised. In other words, one reactant gets oxidised while the other gets reduced during the reaction. Such reactions are called oxidation – reduction reactions or redox reactions.

Oxidation is Gain of oxygen Loss of hydrogen Loss of electron(s)	Reduction is Loss of oxygen Gain of hydrogen Gain of electron(s)
Oxidation and reduction always takes place together, so the reaction is called redox reaction.	

DON'T FORGET

Loss of electron is oxidation.
Gain of electron is reduction.
The term LEO, GER will help you to remember.

MORE TO KNOW

Oxidation also has damaging effects on food and eatables. When food containing fat and oil is left as such for a long time, it becomes stale. The stale food develops bad taste and smell. This is very common in curd or cheese particularly in summer. Oils and fats are slowly oxidised to certain bad smelling compounds.

6. EXOTHERMIC AND ENDOTHERMIC REACTIONS

During chemical reactions one of the most common change is a change in temperature. When detergent is

dissolved in water to wash clothes, heat is given out. When glucose is kept on our tongue, a chilling effect is felt. During these processes, heat is either given out or absorbed from the surroundings. In the same way, in most of the chemical reactions, energy is either taken up or given out.

a. Exothermic reactions

The chemical reactions which proceed with the evolution of heat energy are called exothermic reactions.



All combustion reactions are exothermic. Heat energy is liberated as the reaction proceeds.

b. Endothermic reactions

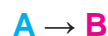
The chemical reactions which proceed with the absorption of heat energy are called endothermic reactions.



11.2 RATE OF THE CHEMICAL REACTION

Rate of the chemical reaction is defined as change in concentration of any one of the reactants or product per unit time.

Consider the reaction



Rate of the reaction is given by

$$\text{Rate} = - \frac{d[\text{A}]}{dt} = + \frac{d[\text{B}]}{dt}$$

[A] - concentration of reactant A

[B] - concentration of product B

- ve sign indicates decrease in concentration of A with time.

+ ve sign indicates increase in concentration of B with time.

11.2.1 FACTORS INFLUENCING THE RATE OF THE CHEMICAL REACTION

1. NATURE OF THE REACTANTS

ACTIVITY 11.10

- Take magnesium ribbon in two test tubes A and B
- Add hydrochloric acid to test tube A
- Add acetic acid to test tube B
- Observe the changes in two test tubes

Magnesium ribbon reacts with both hydrochloric acid and acetic acid but reaction is faster in hydrochloric acid than in acetic acid. Do you know why? Hydrochloric acid is more reactive than acetic acid. It shows that **nature of the reactant influences the rate of the reaction.**

2. CONCENTRATION OF THE REACTANTS

ACTIVITY 11.11

- Take 3g of granulated zinc in the test tube A and B
- Add 5 ml of 1 M hydrochloric acid in test tube A
- Add 5 ml of 2 M hydrochloric acid in test tube B
- Observe the changes

Granulated zinc reacts with both 1M hydrochloric acid and 2M hydrochloric acid, the rate of evolution of hydrogen gas is more from the test tube B than from the test tube A. This is because, 2M hydrochloric acid is more concentrated than 1M hydrochloric acid. That is, **greater the concentration of the reactant, greater will be the rate of the reaction.**

3. SURFACE AREA OF THE REACTANTS

ACTIVITY 11.12

- Take powdered calcium carbonate in beaker A
- Take marble chips (calcium carbonate) in beaker B
- Add hydrochloric acid in both beakers A and B
- Observe the changes

Powdered calcium carbonate reacts more quickly with hydrochloric acid than marble chips. What is the reason?.

Powdered calcium carbonate offers large surface area for the reaction to occur at a faster rate. **This shows that greater the surface area, greater is the rate of the reaction.**

4. TEMPERATURE

ACTIVITY 11.13

- Take 3g of marble chips in a beaker
- Add 5 ml of 1M hydrochloric acid
- Observe the changes
- Heat the beaker
- Observe the changes

Calcium carbonate present in marble chips react slowly with hydrochloric acid at room temperature and evolves carbon dioxide at slower rate, whereas on heating, the evolution of carbon dioxide is made faster. **This shows that increase in temperature increases the rate of the reaction.**

5. CATALYST

ACTIVITY 11.14

- Take potassium chlorate in a test tube
- Heat the test tube
- Observe what happens
- Add manganese dioxide as a catalyst
- Observe the changes

When potassium chlorate is heated, oxygen is evolved very slowly whereas after the addition of manganese dioxide to the reactant, oxygen is liberated at a faster rate. **This shows that manganese dioxide acts as a catalyst and influences the rate of the reaction.**

GROUP ACTIVITY

- From dawn to dusk observe any 10 chemical changes taking place around you and classify them
- Prepare volcano using ammonium dichromate (vigorous)
- Prepare volcano using baking soda (silent).

MORE TO KNOW

A substance which alters the rate of the reaction without undergoing any change in mass and composition is known as catalyst.

ACIDS, BASES AND SALTS

Nivi : Hai Vini, you look tired.
Take this fresh lime juice.

Vini : No, it has sour taste.

Nivi : Do you know why is it sour?

Vini : Sorry, I have no idea at all.

Nivi : It is due to the presence of acid. Ok let's get set to learn about this.

Acids, bases and salts are used in everyday life. Let it be a fruit juice or a detergent or a medicine. They play a key role in our day-to-day activities. Our body metabolism is carried out by means of hydrochloric acid secreted in our stomach.

11.3. ACIDS

Acid is a substance which furnishes H^+ ions or H_3O^+ ions when dissolved in water. Acids have one or more replaceable hydrogen atoms. The word acid is derived from the Latin name 'acidus' which means sour taste. Substances with 'sour taste' are acids. Lemon juice, vinegar and grape juice have sour taste, so they are acidic. They change blue litmus to red. They are

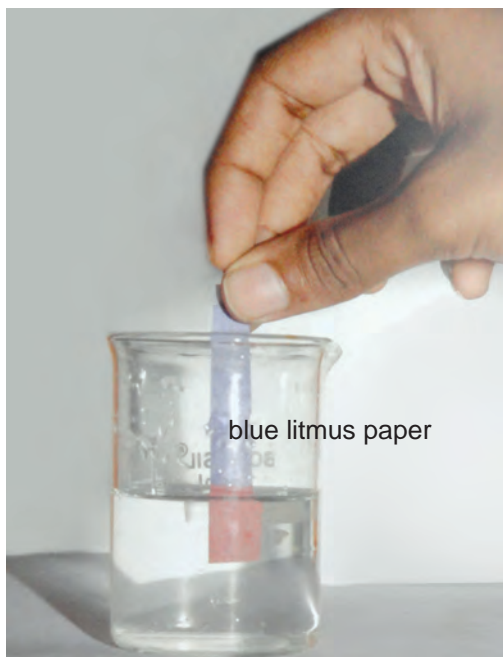


Fig. 11.9 Acid solution turns blue litmus paper red

colourless with phenolphthalein and pink with methyl orange. Many organic acids are naturally present in food items.

11.3.1 CLASSIFICATION OF ACIDS

1. **Based on their sources :** Acids are classified into two types namely organic acids and inorganic acids.

Organic acids:- Acids present in plants and animals (living beings) are **organic acids** eg. HCOOH , CH_3COOH (Weak acids).

Inorganic acids:- Acids from rocks and minerals are **inorganic acids** or mineral acids eg. HCl , HNO_3 , H_2SO_4 (Strong acids).

2. **Based on their basicity**

Monobasic acid: - It is an acid which gives one hydrogen ion per molecule of the acid in solution eg. HCl , HNO_3 .

Dibasic acid:- It is an acid which gives

Source	Acid present
Apple	Malic acid
Lemon	Citric acid
Grape	Tartaric acid
Tomato	Oxalic acid
Vinegar (food preservative)	Acetic acid
Curd	Lactic acid



What is the acid present in it?

two hydrogen ions per molecule of the acid in solution eg. H_2SO_4 , H_2CO_3 .

Tribasic acid:- It is an acid which gives three hydrogen ions per molecule of the acid in solution. eg. H_3PO_4 ,

MORE TO KNOW

For acids, we use the term basicity which means the number of replaceable hydrogen atoms present in one molecule of an acid. For example acetic acid has four hydrogen atoms but only one can be replaced. Hence it is monobasic.

3. Based on ionisation

Acids are classified into two types based on ionisation.

Strong acids:- These are acids which ionise completely in water eg. HCl

Weak acids:- These are acids which ionise partially in water eg. CH_3COOH

4. Based on concentration:-

Depending on the percentage or amount of acid dissolved in water acids are classified into concentrated acid and dilute acid.

Concentrated acid:- It is an acid having a relatively high percentage of acid in its aqueous solution.

Dilute acid:- It is an acid having a relatively low percentage of acid in aqueous solution.

MORE TO KNOW

Care must be taken while mixing any concentrated mineral acid with water. The acid must always be added slowly to water with constant stirring. If water is added to a concentrated acid the large amount of heat is generated which may cause burns. The mixture splashes out of the container.

11.3.2 CHEMICAL PROPERTIES OF ACIDS

1. REACTION OF METALS WITH ACID

Note that zinc reacts with dilute hydrochloric acid to form zinc chloride

ACTIVITY 11.15

- Take 5 g of zinc granules in a test tube
- Add 10 ml of dilute hydrochloric acid through thistle funnel
- During the course of addition, what do you observe?

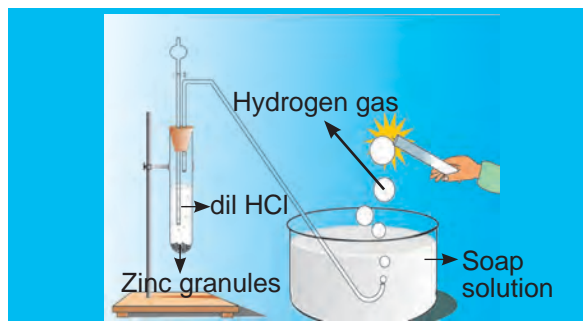


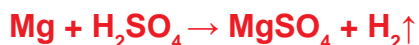
Fig. 11.10 Reaction of Zn granules with dilute HCl and hydrogen gas.



When a burning candle is brought near the bubble containing hydrogen gas, the flame goes off with a 'pop'ing sound. This confirms that metal displaces hydrogen from the dilute acid. (Hydrogen gas burns with a 'pop'ing sound)



Another example



MORE TO KNOW

- All metals do not liberate hydrogen gas on reaction with acids. eg., Ag, Cu.
- Lime stone, chalk and marble are different physical forms of calcium carbonate. They react with acids giving corresponding salt, carbon dioxide and water.

2. REACTION OF METAL CARBONATE AND METAL BICARBONATE WITH ACIDS

ACTIVITY 11.16

- Take two test tubes, label them as I and II
- Take small amount of washing soda (Na_2CO_3) in test tube I and small amount of baking soda (NaHCO_3) in test tube II
- Add dilute hydrochloric acid to both the test tubes
- What do you observe?
- Pass the gas produced in each case, through lime water [$\text{Ca}(\text{OH})_2$] solution and record your observations

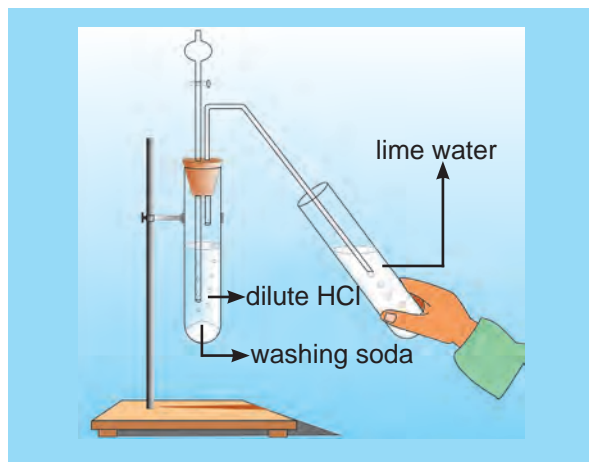


Fig. 11.11 Testing of carbon dioxide

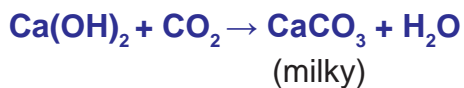
Test tube I



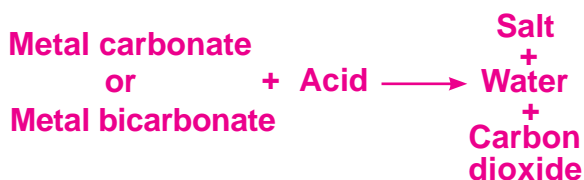
Test tube II



When carbon dioxide is passed through lime water, it turns milky.



From the above activity the reaction can be summarized as



Other examples



MORE TO KNOW

Since metal carbonates and metal bicarbonates are basic they react with acids to give salt and water with the liberation of carbon dioxide.

3. REACTION OF METALLIC OXIDES WITH ACIDS

ACTIVITY 11.17

- Take about 2g copper (II) oxide in a watch glass and add dilute hydrochloric acid slowly
- Note the colour of the salt
- What has happened to the copper (II) oxide?

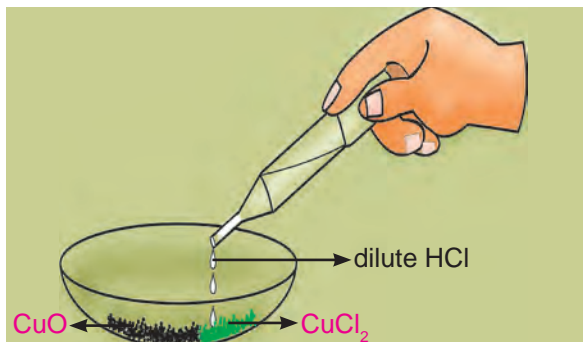


Fig. 11.12 Reaction of copper(II) oxide with dilute hydrochloric acid

The colour changes from **black to green**. This is due to the formation of copper (II) chloride in the reaction. Since metal oxides are basic, they react with acid to form salt and water.



From the above activity we conclude that



Another example



4 . ACTION OF ACIDS WITH WATER.

An acid produces hydrogen ions in water.



Hydrogen ions cannot exist alone, but they exist in the form of hydronium (H_3O^+) ions. When water is absent, the separation of hydrogen ions from an acid does not occur.

11.3.3. USES OF ACIDS

1. Sulphuric acid (King of chemicals) is used in car battery and in the preparation of many other compounds.
2. Nitric acid is used in the production of ammonium nitrate which is used as fertilizer in agriculture.
3. Hydrochloric acid is used as cleansing agent in toilet.
4. Tartaric acid is a constituent of baking powder.
5. Salt of benzoic acid (sodium benzoate) is used in food preservation.
6. Carbonic acid is used in aerated drinks.

MORE TO KNOW

The atmosphere of Venus is made up of thick white and yellowish clouds of sulphuric acid. Do you think life can exist on this planet?

11.4. BASES

Base is a substance which releases hydroxide ions when dissolved in water. It is a substance which is bitter in taste and soapy to touch (e.g. Washing soda, caustic soda and caustic potash). They change red litmus to blue. They are pink with phenolphthalein and yellow with methyl orange.

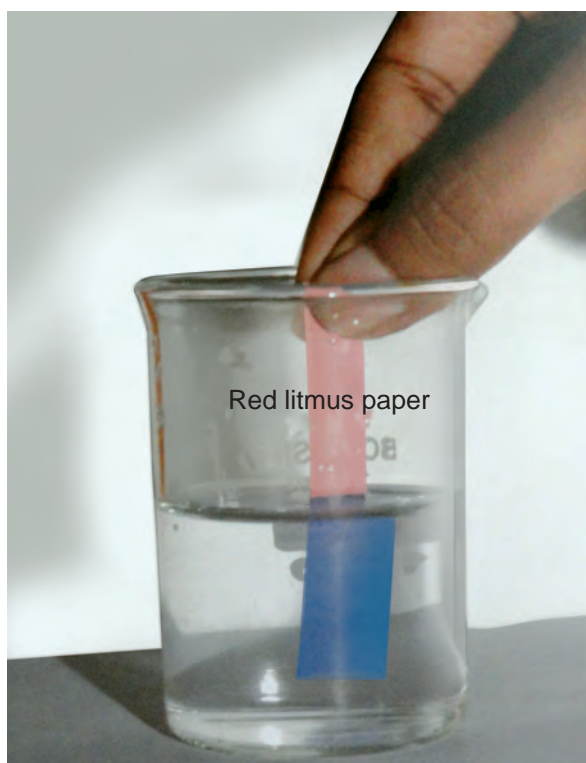


Fig. 11.13 Bases turns red litmus paper blue

11.4.1. Classification of bases

1. Based on ionisation

Strong bases:- These are bases which ionise completely in aqueous solution eg. NaOH, KOH.

Weak bases:- These are bases which ionise partially in aqueous solution eg. NH_4OH , $\text{Ca}(\text{OH})_2$.

2. Based on their acidity

Monoacidic base:- It is a base which ionises in water to give

one hydroxide ion per molecule eg. NaOH, KOH.

Diacidic base:- It is a base which ionises in water to give two hydroxide ions per molecule eg. $\text{Ca}(\text{OH})_2$, $\text{Mg}(\text{OH})_2$.

Triacidic base:- It is a base which ionises in water to give three hydroxide ions per molecule eg. $\text{Al}(\text{OH})_3$, $\text{Fe}(\text{OH})_3$.

MORE TO KNOW

The term acidity is used for base which means the number of replaceable hydroxyl groups present in one molecule of a base.

3. Based on the concentration:

Depending on the percentage or amount of base dissolved in water, bases are classified as concentrated alkali and dilute alkali.

Concentrated alkali:- It is an alkali having a relatively high percentage of alkali in its aqueous solution.

Dilute alkali:- It is an alkali having a relatively low percentage of alkali in its aqueous solution.

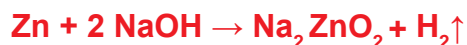
MORE TO KNOW

Bases which dissolve in water are called alkalies. All alkalies are bases, but not all bases are alkalies. NaOH and KOH are alkalies whereas $\text{Al}(\text{OH})_3$ and $\text{Zn}(\text{OH})_2$ are bases.

11.4.2. Chemical Properties Of Bases

1. REACTION OF BASE WITH METALS

Zinc reacts with sodium hydroxide to form sodium zincate with the liberation of hydrogen gas.



Another example



MORE TO KNOW

All metals do not react with sodium hydroxide eg. Cu, Ag, Cr

2. REACTION OF NON METALLIC OXIDES WITH BASES

Sodium hydroxide reacts with carbon dioxide gives sodium carbonate and water.



The above reaction confirms that



Another example



3. ACTION OF BASES WITH WATER

Bases generate hydroxide (OH^-) ions when dissolved in water.



4. REACTION OF ACIDS WITH BASES

ACTIVITY 11.18

- Indira takes 20 ml of 0.1N sodium hydroxide solution in a conical flask and adds few drops of phenolphthalein.
- What colour does she observe?
- She is adding 20 ml of 0.1N hydrochloric acid solution to the above conical flask drop by drop.
- Does she observe any colour change in the reaction mixture?

In the above activity, Indira observed that the effect of a base is nullified by an acid.



The above reaction between an acid and a base is known as neutralisation reaction.



NaOH solution NaOH Solution + Phenolphthalein NaOH solution + Phenolphthalein + HCl Solution

Fig. 11.14 Reaction of sodium hydroxide with hydrochloric acid

11.4 USES OF BASES

1. Sodium hydroxide is used in the manufacture of soap.
2. Calcium hydroxide is used in white washing the buildings.
3. Magnesium hydroxide is used as a medicine for stomach troubles.
4. Ammonium hydroxide is used to remove grease stains from clothes.

11.5 IDENTIFICATION OF ACIDS AND BASES

ACTIVITY 11.19

- Collect lemon juice, washing soda solution, soap solution and soft drinks.
- Take 2 ml of each solution in a test tube and test with a litmus paper or indicator.
- What change in colour do you observe with red litmus, blue litmus, phenolphthalein and methyl orange?
- Tabulate your observations.

Sample solution	Red litmus	Blue litmus	Phenolphthalein	Methyl orange
Lemon Juice				
Washing soda Solution				
Soap solution				
Soft drinks				

Same activity can be repeated for dilute hydrochloric acid, dilute sulphuric acid, sodium hydroxide solution and potassium hydroxide solution with the help of your teacher.

INDICATOR	COLOUR IN ACID	COLOUR IN BASE
Litmus	Red	
Phenolphthalein		
Methyl orange	Red	Yellow

11.6 pH SCALE

pH stands for the power of hydrogen ion concentration in a solution. pH values decide whether a solution is acidic or basic or neutral. pH scale was introduced by S.P.L. Sorenson. It is mathematically expressed as

$$\text{pH} = -\log_{10} [\text{H}^+]$$

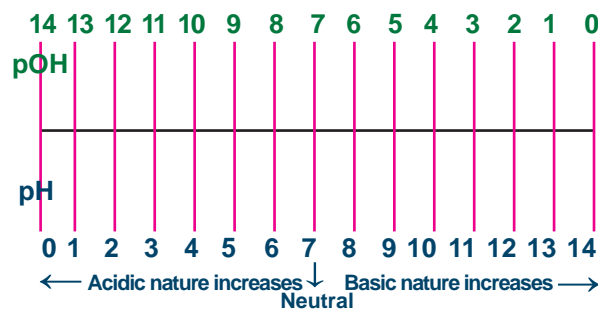
For neutral solution $[\text{H}^+] = 10^{-7}\text{M}$; $\text{pH} = 7$

For acidic solution $[\text{H}^+] > 10^{-7}\text{M}$; $\text{pH} < 7$

For basic solution $[\text{H}^+] < 10^{-7}\text{M}$; $\text{pH} > 7$

When OH^- ions are taken into account the pH expression is replaced by pOH

$$\text{pOH} = -\log_{10} [\text{OH}^-]$$



Problems

1. The hydrogen ion concentration of a solution is 0.001M. What is the pH of the solution?

Solution

$$\text{pH} = -\log_{10} [\text{H}^+]$$

$$\text{pH} = -\log_{10} (0.001)$$

$$\text{pH} = -\log_{10} (10^{-3})$$

$$= -(-3) \log_{10} 10 \quad [\log 10 = 1]$$

$$\text{pH} = 3$$

2. The hydrogen ion concentration of a solution is 1.0×10^{-9} M. What is the pH of the solution? Predict whether the given solution is acidic, basic or neutral.

Solution

$$\text{pH} = -\log_{10} [\text{H}^+]$$

$$\text{pH} = -\log_{10} (1.0 \times 10^{-9})$$

$$\text{pH} = -(\log_{10} 1.0 + \log_{10} 10^{-9}) \quad [\log_{10} 1 = 0]$$

$$= -(0 - 9 \log_{10} 10)$$

$$\text{pH} = -(0 - 9) = 9$$

$$\text{pH} = 9 \text{ ie } \text{pH} > 7$$

Therefore the given solution is **basic**.

3. The hydroxyl ion concentration of a solution is 0.001M. What is the pH of the solution?

Solution

$$\text{pOH} = -\log_{10} [\text{OH}^-]$$

$$\text{pOH} = -\log_{10} (10^{-3})$$

$$\text{pOH} = 3$$

$$\text{pH} = 14 - \text{pOH}$$

$$\text{pH} = 14 - 3 = 11$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{pH} = 14 - \text{pOH}$$

4. The hydroxyl ion concentration of a solution is 1.0×10^{-9} M. What is the pH of the solution?

Solution

$$\text{pOH} = -\log_{10} [\text{OH}^-]$$

$$\text{pOH} = -\log_{10} (1.0 \times 10^{-9})$$

$$\text{pOH} = 9$$

$$\text{pH} = 14 - \text{pOH}$$

$$\text{pH} = 14 - 9 = 5$$

11.6.1 pH paper

A more common method of measuring pH in a school laboratory is by using pH paper. pH paper contains a mixture of indicators, which gives different colours across the entire pH range. pH value of the various solutions are given in the table.

$$\text{pH} = -\log_{10} [\text{H}^+]$$

$$\text{pH} = -\log_{10} \left[\frac{1}{\text{H}^+} \right]$$

$$[\text{H}^+] = 10^{-\text{pH}}$$

$$[\text{H}^+] = 1 \times 10^{-7}; \text{pH} = 7$$

$$[\text{H}^+] = 1 \times 10^{-2}; \text{pH} = 2$$

$$[\text{H}^+] = 1 \times 10^{-14}; \text{pH} = 14$$

Solution	Approximate pH
Lemon juice	2.2 - 2.4
Tomato juice	4.1
Coffee	4.4 - 5.5
Human saliva	6.5 - 7.5
House hold ammonia	12.0



Fig. 11.15 pH paper

ACTIVITY 11.20

- Take lemon juice, orange juice, 1M NaOH, 1M HCl, pure water and vinegar
- Dip pH paper into these solutions
- Observe the changes

Sl. No.	Sample	Colour of pH paper	Approximate pH	Nature of substance
1.	Lemon juice			
2.	Orange juice			
3.	1M NaOH			
4.	1M HCl			
5.	Pure H ₂ O			
6.	Vinegar			

11.6.2 Importance of pH in everyday life

1. pH in human body

- (i) Using pH factor the healthiness of our body is predicted. At pH level 6.9, the body becomes prone to viral infections like colds, cough and flu. Cancer cells thrive inside the body at a pH of 5.5.
- (ii) The pH of a normal, healthy human skin is 4.5 to 6. Proper skin pH is essential for a healthy complexion.
- (iii) pH of stomach fluid is approximately 2.0. This fluid is essential for the digestion of food.
- (iv) Human blood pH range is 7.35 to 7.45. Any increase or decrease in this value, leads to diseases. The ideal pH for blood is 7.4.
- (v) pH of normal saliva ranges between 6.5 to 7.5.
- (vi) White enamel coating in our teeth is calcium phosphate, hardest substance in our body. It does not dissolve in water. If pH of mouth falls below 5.5, the enamel gets corroded. Toothpastes are generally basic, and is used for cleaning the teeth, can neutralize the excess acid and prevent tooth decay.

2. pH in soil

In agriculture, the pH of soil is very important. Citrus fruits require slightly alkaline soil, while rice requires acidic soil and sugar cane requires neutral soil.

3. pH in rain water

pH of rain water is approximately 7 showing high level of its purity and neutrality. If rain water is polluted by SO_2 and NO_2 , acid rain occurs, bringing the pH value less than 7.

11.7 SALT

When you say salt, you may think of white stuff put on chips. But that is just one salt called common salt. There are many other salts used in other fields. Salts are the products of the reaction between acids and bases (see reaction of acids and bases), which produce positive ions and negative ions when dissolved in water.

11.7.1 Classification of salts

1. Normal salts

A normal salt is obtained by complete neutralization of an acid by a base



2. Acid salts

Acid salts are derived by the partial replacement of hydrogen ions of an acid by a metal. When a calculated amount of a base is added to a polybasic acid, acid salt is obtained, as follows.



3. Basic salts

Basic salts are formed by the partial replacement of hydroxide ions of a diacidic or triacidic base by an acid radical.

A basic salt may further reacts with an acid to give a normal salt.



(Diacidic base) Basic salt

4. Double salts

Double salts are formed by the combination of saturated solution of two simple salts in equimolar ratio followed by crystallization.

e.g. potash alum

11.7.2 USES OF SALTS

Common salt (NaCl)

It is used in our daily food and as preservative.

Washing soda (Na₂CO₃)

1. It is used in softening hard water.
2. It is used as a cleaning agent for domestic purposes.

Baking soda (NaHCO₃)

1. It is used in making baking powder, which is the mixture of baking soda and tartaric acid. Baking powder is

used to make cake and bread soft and spongy .

2. It is an ingredient in antacid. Being alkaline, it neutralises excess of acid in the stomach.

Bleaching powder (CaOCl₂)

1. It is used for disinfecting drinking water to make it free from microorganisms.
2. It is used for bleaching cotton and linen in the textile industry

Plaster of paris(CaSO₄· 1/2H₂O)

It is used for plastering fractured bones and in making casts for statues

GROUP ACTIVITY

Prepare the following salt in the laboratory

1. Sodium chloride
2. Potash alum

EVALUATION

PART A



The above reaction is an example of

- a) Combination reaction
 - b) Double displacement reaction
 - c) Displacement reaction
 - d) Decomposition reaction.
2. A reddish brown coloured element 'X' on heating in air becomes black coloured compound 'Y'. X and Y are _____ and _____(Cu, CuO / Pb, PbO).
 3. A student tested the pH of pure water using a pH paper. It showed green colour. If a pH paper is used after adding lemon juice into water, what color will he observe? (Green / Red / Yellow)
 4. Chemical volcano is an example of (combination reaction / decomposition reaction)
 5. When crystals of lead nitrate on heating strongly produces a _____ gas and the colour of the gas is _____.

6. When aqueous solution of silver nitrate and sodium chloride are mixed _____ precipitate is immediately formed (white / yellow / red).
7. Zinc can displace aluminium metal from aqueous solution of aluminium sulphate (zinc is more reactive than aluminium / aluminium is more reactive than zinc).
8. To protect tooth decay, we are advised to brush our teeth regularly. The nature of the tooth paste commonly used is _____ in nature.
9. Vinegar is present in acetic acid. Curd contains _____ acid (Lactic acid / Tartaric acid).
10. $\text{pH} = -\log_{10} [\text{H}^+]$. The pH of a solution containing hydrogen ion concentration of 0.001M solution is _____ (3 / 11 / 14)

PART B

11. What type of chemical reaction takes place when i) limestone is heated ii) a magnesium ribbon is burnt in air
12. The pH values of certain familiar substances are given below

Substance	pH value
Blood	7.4
Baking soda	8.2
Vinegar	2.5
Household ammonia	12

analyse the data in the table and answer the following questions

- a) Which substance is acidic in nature?
- b) Which substances are basic in nature?
13. Why does the colour of copper sulphate change when an iron nail is kept in it? Justify your answer.
14. The hydroxyl ion concentration of a solution is $1.0 \times 10^{-8}\text{M}$. What is the pH of the solution?
15. Equal lengths of magnesium ribbons are taken in test tubes A and B. Hydrochloric acid is added to test tube A, while acetic acid is added to test tube B. Amount and concentration taken for both the acids are same in which test tube reaction occurs more vigorously and why.

FURTHER REFERENCE

Books:

1. Text book of Inorganic Chemistry—**P.L. Soni - S.Chand & sons publishers**
2. Principles of Physical Chemistry —**B.R. Puri, L.R. Sharma Vishal publishers**

Websites:

- [www. ask.com](http://www.ask.com)
www.chem4kids.com

Chapter

12



PERIODIC CLASSIFICATION OF ELEMENTS

12. Periodic classification of elements

Have you ever visited a library? There are thousands of books in a large library. If you ask for a book in general it is very difficult to trace. Whereas if you ask for a particular book, the library staff can locate it very easily. How is it possible? In library the books are classified into various categories and sub categories. They are arranged on shelves accordingly. Therefore locating books become very easy.

As on date one hundred and eighteen elements are known. It is difficult to identify each and every element individually and to know its property and uses. Therefore they have been classified on the basis of their similarities in properties. One of



Henry Gwyn-Jeffreys Moseley, an English physicist (1887–1915), used X-rays to determine the atomic numbers of the elements.

the important instincts of mankind is to be systematic. Scientists felt the necessity to group elements of similar characteristics together so that if the properties of one of them are known, those of the others could be guessed and related.

When a large number of elements were discovered, several attempts were being made to arrange them on the basis of their properties, nature, character, valency, etc., (Real credit for preparing the periodic table goes to Mendeleev).

12.1. MODERN PERIODIC LAW

A large number of scientists made attempts to eliminate the drawbacks of Mendeleev's periodic table. In 1912, Moseley, an English physicist measured the frequencies of X-rays emitted by a metal, when the metal was bombarded with high speed electrons. He plotted square roots of the frequencies against atomic numbers. The plot obtained was a straight line. He found that the square root of the frequency of the prominent X-rays emitted by a metal was proportional to the atomic number and not to the atomic weight of the atom of that metal.

MORE TO KNOW

Atomic number is number of protons in the nucleus or number of electrons revolving around the nucleus in an atom.

Moseley suggested that atomic number (Z) should be the basis of the classification of the element. Thus, he gave modern periodic law as follows:

Modern periodic law states that “**the physical and chemical properties of elements are the periodic function of their atomic numbers.**”

Thus, according to the modern periodic law, if elements are arranged in the increasing order of their atomic numbers, the elements with similar properties are repeated after certain regular intervals.

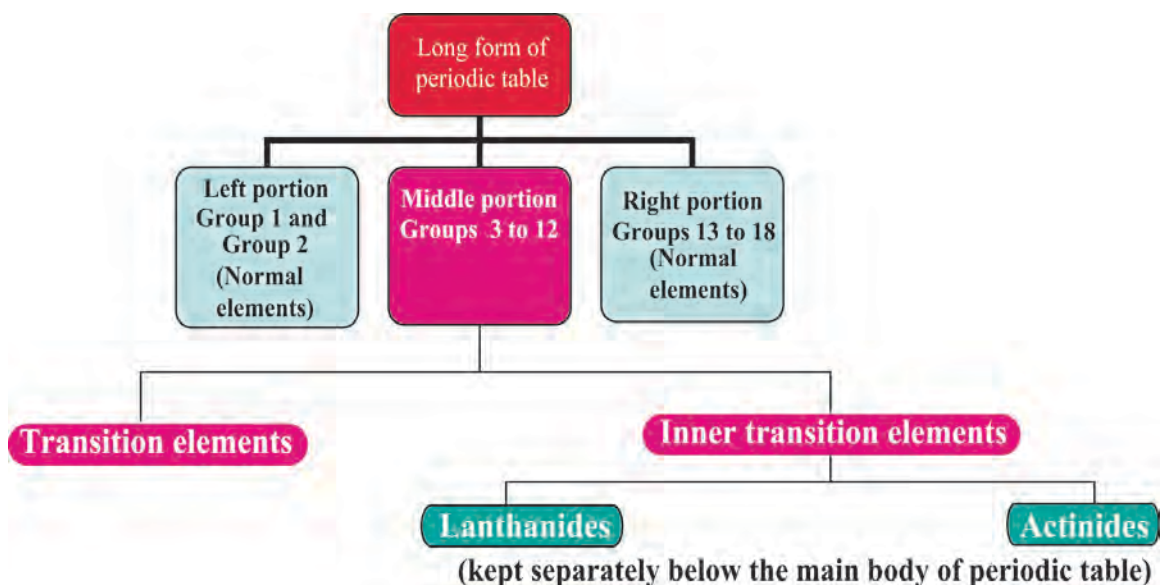
12.2. MODERN PERIODIC TABLE

Based on the modern periodic law, a number of forms of periodic table have been proposed from time to time but general plan of the table remained the same as proposed by Mendeleev. The table which is most commonly used and which is based upon the **electronic configuration of elements** is called the **long form of the periodic table**. This is called the **modern periodic table**.

12.2.1. Description of modern or long form of the periodic table

Long form of the periodic table is a chart of elements in which the elements have been arranged in the increasing order of their atomic numbers. This table consists of **horizontal rows called periods** and **vertical columns called groups**.

12.2.2. Different portions of long form of periodic table



12.2.3. Study of periods

The **horizontal rows** are called

MORE TO KNOW

The modern periodic table has also been divided into four blocks known as s,p,d and f blocks.

periods. There are **seven** horizontal rows in the periodic table.

- **First period** (Atomic number 1 and 2): This is the shortest period. It contains only two elements (Hydrogen and Helium).
- **Second period** (Atomic number 3 to 10): This is a short period. It contains eight elements (Lithium to Neon).
- **Third period** (Atomic number 11 to 18): This is also a short period. It contains eight elements (Sodium to Argon).

- **Fourth period** (Atomic number 19 to 36):
This is a long period. It contains eighteen elements (Potassium to Krypton). This includes 8 normal elements and 10 transition elements.
- **Fifth period** (Atomic number 37 to 54):
This is also a long period. It contains 18 elements (Rubidium to Xenon). This includes 8 normal elements and 10 transition elements.
- **Sixth period** (Atomic number 55 to 86):
This is the longest period. It contains 32 elements (Cesium to Radon). This includes 8 normal elements, 10 transition elements and 14 inner transition elements (Lanthanides).
- **Seventh period** (Atomic number 87 to 118):
As like the sixth period, this period also can accommodate 32 elements. Till now only 26 elements have been authenticated by IUPAC

12.2.4. Study of groups

- Vertical columns in the periodic table starting from top to bottom are called groups. There are 18 groups in the periodic table.
- First group elements are called alkali metals.
- Second group elements are called alkaline earth metals.
- Groups three to twelve are called transition elements .
- Group 1, 2 and 13- 18 are called normal elements or main group elements or representative elements .
- Group 13 is Boron family.
- Group 14 is Carbon family.

- Group 15 is Nitrogen family.
- Group 16 elements are called chalcogen family (except polonium).
- Group 17 elements are called halogen family.
- Group 18 elements are called noble gases or inert gases.
- The Lanthanides and actinides which form part of the group 3 are called inner transition elements.

12.3. CHARACTERISTICS OF MODERN PERIODIC TABLE

12.3.1. Characteristics Of Periods

- In a period, the electrons are filled in the same valence shell of all elements.
- As the electronic configuration changes along the period, the chemical properties of the elements also change.
- Atomic size of the elements in a period decrease from left to the right.
- In a period, the metallic character of the element decreases while their non-metallic character increases.

12.3.2. Characteristics of Groups

- The elements present in 2 and 18 groups differ in atomic number by 8,8,18,18,32.
- The elements present in 13 – 17 groups differ in atomic number by 8,18,18,32.
- The elements present in 4 - 12 groups differ in atomic number by 18,32,32.

- The elements present in a group have the same number of electrons in the valence shell of their atoms.
- The elements present in a group have the same valency.
- The elements present in a group have identical chemical properties.
- The physical properties of the elements in group such as melting point, boiling point, density vary gradually.
- Atomic radii of the elements present in a group increases downwards.

12.3.3. Advantages of the Modern Periodic Table

- The table is based on a more fundamental property i.e., atomic number.
- It correlates the position of the element with its electronic configuration more clearly.
- The completion of each period is more logical. In a period as the atomic number increases, the energy shells are gradually filled up until an inert gas configuration is reached.
- It is easy to remember and reproduce.
- Each group is an independent group and the idea of sub-groups has been discarded.
- One position for all isotopes of an element is justified, since the isotopes have the same atomic number.
- The position of eighth group

(in Mendeleev's table) is also justified in this table. All transition elements have been brought in the middle as the properties of transition elements are intermediate between left portion and right portion elements of the periodic table.

- The table completely separates metals from non-metals. The non-metals are present in upper right corners of the periodic table.
- The positions of certain elements which were earlier misfit (interchanged) in the Mendeleev's periodic table are now justified because it is based on atomic number of the elements.
- Justification has been offered for placing lanthanides and actinides at the bottom of the periodic table.

12.3.4. Defects in the Modern Periodic Table

- Position of hydrogen is not fixed till now.
- Position of lanthanides and actinides has not been given inside the main body of periodic table.
- It does not reflect the exact distribution of electrons of some of transition and inner transition elements.

MORE TO KNOW

The last element authenticated by IUPAC is Cn112 [Copernicium]. However, the number of elements discovered so far is 118.

12.4. METALLURGY



I (Al) am a light silvery white metal to build aircraft.
So, I am great.

I (Fe) am a lustrous steel metal to make machineries and bridges. So, I am great.

I (Cu) am a reddish brown metal to make coins.
So, I am great.



Individually you are great in your aspect. You will become the **GREATEST IF YOU ARE ALLOYED TOGETHER.**
Unity is strength.



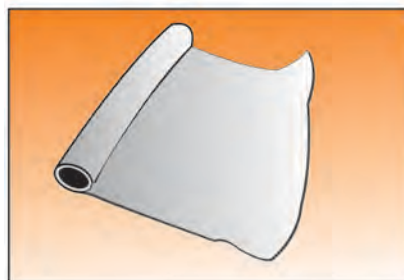
INTRODUCTION

Metallurgy is as old as our civilization. Copper was the first metal to be used for making utensils, weapons and for other works. Metals play a significant role in our life. They constitute the mineral wealth of a country which is the measure of prosperity.

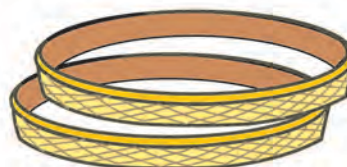
Metals like titanium, chromium, manganese, zirconium etc. find their applications in the manufacture of defence equipments. These are called **strategic metals**. The metal uranium plays, a vital role in nuclear reactions releasing enormous energy called nuclear energy. Copper, silver and gold are called **coinage metals** as they are used in making coins, jewellery etc.



Vietnamese Craft Work in silver



Aluminium foil



Bangles

MORE TO KNOW

Purity of gold is expressed in carat.

24 carat gold = pure gold.

For making ornaments 22 carat gold is used which contains 22 parts of gold by weight and 2 parts of copper by weight. The percentage of purity is $\frac{22}{24} \times 100 = 91.6\%$ (**916 Make gold**)

From one gram of gold, nearly 2km of wire can be drawn. Its an amazing fact indeed!

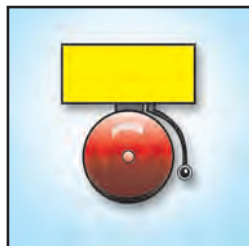
MORE TO KNOW

THE VITALITY OF METALS FOR THE TOTALITY OF LIFE

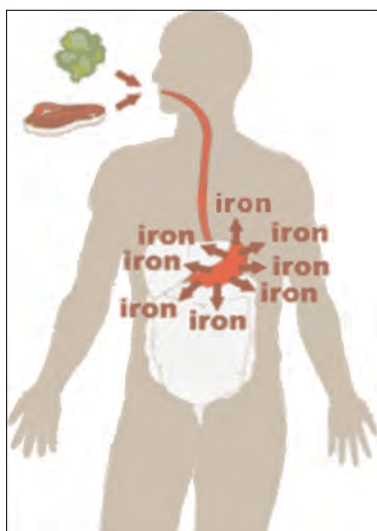
Metals in minute amounts are essential for various biological purposes. **Fe** – a constituent of blood pigment (haemoglobin).

Ca - a constituent of bone and teeth. **Co** - a constituent of vitamin B-12

Mg - constituent of chlorophyll.



METALS AROUND US



12.4. TERMINOLOGIES IN METALLURGY

12.4.1. Minerals: A mineral may be a single compound or complex mixture of various compounds of metals which are found in earth.

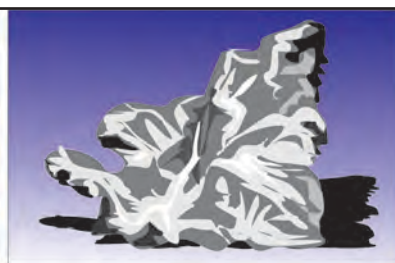
12.4.2. Ores: The mineral from which a metal can be readily and economically

extracted on a large scale is said to be a ore.

For example, clay ($\text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2 \cdot 2\text{H}_2\text{O}$) and bauxite ($\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$) are the two minerals of aluminium. But aluminium can be profitably extracted only from bauxite. Hence **bauxite is an ore of aluminium and clay is its mineral.**



Gold



Silver



Aluminium

12.4.3. Differences between minerals and ores

- Minerals contain a low percentage of metal while ores contain a large percentage of metal.
- Metals cannot be extracted easily from mineral. On the other hand, ores can be used for the extraction of metals.
- All minerals cannot be called as ores, but all ores are minerals.

Mining: The process of extracting the ores from the earth crust is called mining.

Metallurgy: Various steps involved in the extraction of metals from their ores as well as refining of crude metal are collectively known as metallurgy.

Gangue or Matrix: The rocky impurity, associated with the ore is called gangue or matrix.

Flux: It is the substance added to the ore to reduce the fusion temperature

Slag: It is the fusible product formed when flux reacts with gangue during the extraction of metals.



Smelting: Smelting is the process of reducing the roasted oxide to metals in the molten condition.

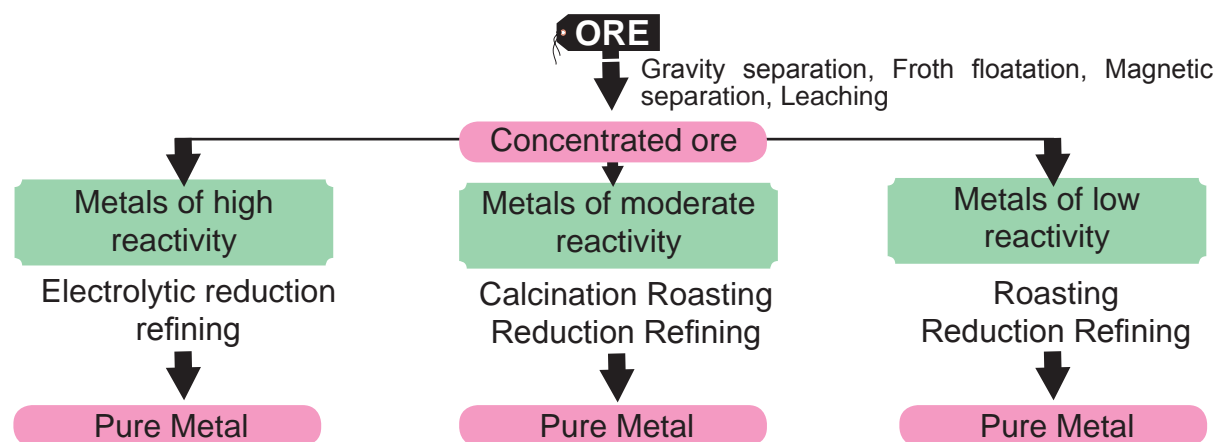
12.5. OCCURRENCE OF METALS

Nearly 80 metallic elements are obtained from mineral deposits on or beneath the surface of the earth. Metals which have low chemical reactivity are found in **free state, or in native state**.

Gold, silver and platinum are examples of metals that are partly found in a free state. Most of the other metals are found in a combined state in the form of their oxide ores, carbonate ores, halide ores, sulphide ores, sulphate ores and so on.

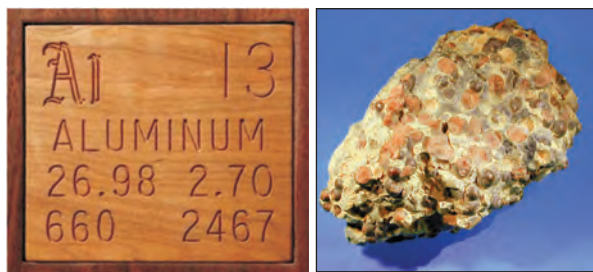
Oxide Ores	Carbonate Ores	Halide Ores	Sulphide Ores
Bauxite (Al ₂ O ₃ .2H ₂ O)	Marble (CaCO ₃)	Cryolite (Na ₃ AlF ₆)	Galena (PbS)
Cuprite (Cu ₂ O)	Magnesite (MgCO ₃)	Fluorspar (CaF ₂)	Iron pyrite (FeS ₂)
Haematite (Fe ₂ O ₃)	Siderite (FeCO ₃)	Rock salt (NaCl)	Zinc blende (ZnS)

Flow Chart (Extraction of Metal from its ore)



12.6. METALLURGY OF ALUMINIUM, COPPER AND IRON

12.6.1. Metallurgy of aluminium



Symbol : Al
Colour : Silvery white
Atomic number : 13
Electronic configuration: 2, 8, 3
Valency : 3
Atomic mass : 27

Position in the periodic table: period=3, group=13 (III A)

Aluminium is the most abundant metal in the earth's crust. Since it is a reactive metal it occurs in the combined state. The important ores of aluminium are as follows:

Name of the ore	Formula
Bauxite	$\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$
Cryolite	Na_3AlF_6
Corundum	Al_2O_3

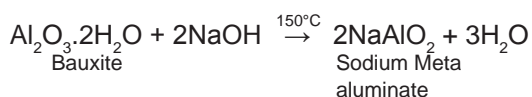
The chief ore of aluminium is bauxite ($\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$).

Extraction of aluminium from bauxite involves two stages:

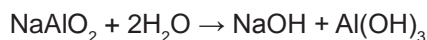
I. Conversion of Bauxite into Alumina by Baeyer's Process

The conversion of Bauxite into Alumina involves the following steps:

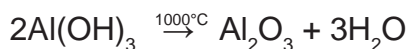
i. Bauxite ore is finely grounded and heated under pressure with concentrated caustic soda solution at 150°C to obtain sodium meta aluminate.



ii. On diluting sodium meta aluminate with water, aluminium hydroxide precipitate is obtained.



iii. The precipitate is filtered, washed, dried and ignited at 1000°C to get alumina.



2. Electrolytic reduction of Alumina by Hall's process

Aluminium is produced by the electrolytic reduction of fused alumina (Al_2O_3) in the electrolytic cell.

Cathode : Iron tank lined with graphite

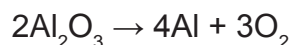
Anode : A bunch of graphite rods suspended in molten electrolyte

Electrolyte : Pure alumina + molten cryolite + fluorspar (fluorspar lowers the fusion temperature of electrolyte)

Temperature : $900-950^\circ\text{C}$

Voltage used : 5-6V

The overall equation for aluminium extraction is



Aluminium deposits at cathode and oxygen gas is liberated at anode

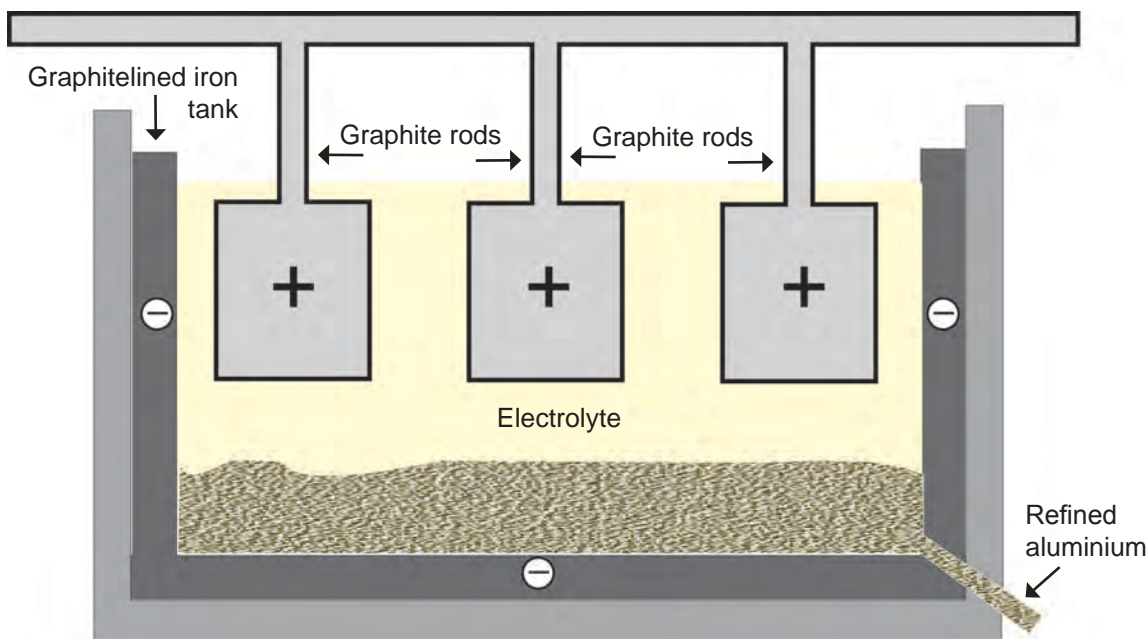


Fig 12.6.3 Electrolytic refining of aluminium

Properties of Aluminium

Physical properties:

- i. It is a silvery white metal.
- ii. It has low density and it is light
- iii. It is malleable and ductile.
- iv. It is a good conductor of heat and electricity.

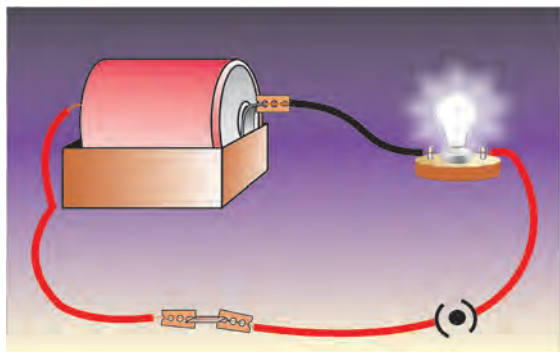


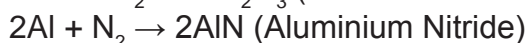
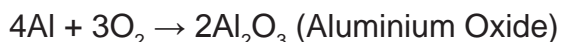
Fig. 12.6.4 Electric conductivity of metal

- v. Melting point: 660°C
- vi. It can be well polished to produce attractive shiny appearance.

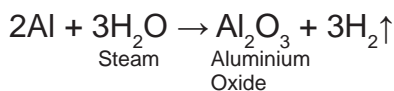
Chemical properties:

1. Reaction with air: It is not affected by dry air. On heating at 800°C , aluminium burns

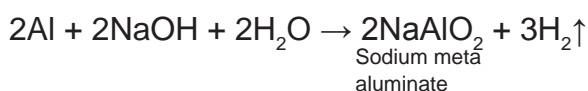
very brightly forming its oxide and nitride.



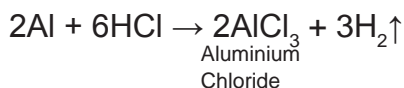
2. Reaction with water: Water has no reaction on aluminium due to the layer of oxide on it. When steam is passed over red hot aluminium, hydrogen is produced.



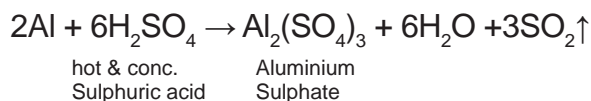
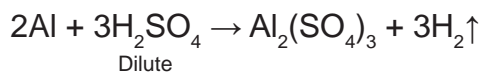
3. Reaction with alkalis: It reacts with strong caustic alkalis forming aluminates.



4. Reaction with acids: With dilute and con. HCl it liberates H_2 gas.



Aluminium liberates hydrogen on reaction with dilute sulphuric acid. Sulphur dioxide is liberated with hot concentrated sulphuric acid.

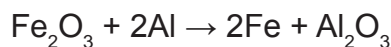


MORE TO KNOW

MORE TO KNOW

Dilute or concentrated nitric acid does not attack aluminium. But it renders aluminium passive due to the formation of an oxide film on its surface.

5. Reducing action : Aluminium is a powerful reducing agent. When a mixture of aluminium powder and iron oxide is ignited, the latter is reduced to metal. This process is known as aluminothermic process.



Uses of Aluminium

USES	FORM	REASON
1. Household utensils	Aluminium metal	It is light, cheap, corrosion resistant, and good conductor of heat.
2. Electrical cable industry	Aluminium wires	It is a good conductor of electricity.
3. Aeroplanes and other industrial parts	Duralumin Al, Cu, Mg, Mn Magnalium Al, Mg	Its alloys are light, have high tensile strength and are corrosion resistant.
4. Thermite welding	Al powder and Fe_2O_3	Its powder is a strong reducing agent and reduces Fe_2O_3 to iron.



AirCraft - An alloy of aluminium

INDUSTRIAL VISIT



Fig 12.6.6

Make an industrial visit to the place where **Thermite welding** is actually done and record your observations on joining the gap between the broken pieces of rails.

12.6.2 Metallurgy of Copper



Symbol : Cu

Atomic mass : 63.55

Atomic number : 29

Electronic

configuration : 2, 8, 18, 1

Valency : 1 and 2

Occurrence: It was named as cuprum by the Romans because they used to get it from the island of Cyprus. Copper is found in the **native state** as well as in the combined state.

Ores of copper	Formula
i. Copper pyrite	CuFeS_2
ii. Cuprite or ruby copper	Cu_2O
iii. Copper glance	Cu_2S

The chief ore of copper is copper pyrite. It yields nearly 76% of the world production of copper.

Extraction from copper pyrites:

Extraction of copper from copper pyrites involves the following steps.

1. Crushing and concentration: The ore is crushed and then concentrated by froth-floatation process.

2. Roasting: The concentrated ore is roasted in excess of air. During roasting,

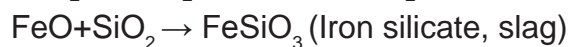
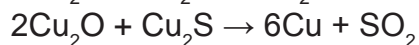
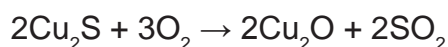
i. moisture and volatile impurities are removed.

ii. copper pyrite is partly converted into sulphides of copper and iron.



3. Smelting: The roasted ore is mixed with powdered coke and sand and is heated in a blast furnace to obtain matte and slag. (**Matte = Cu₂S + FeS**) The slag is removed as a waste.

4. Bessemerisation: The molten matte is transferred to Bessemer converter in order to obtain **blister copper**. Ferrous sulphide from matte is oxidised to ferrous oxide which is removed as slag using silica.



5. Refining: Blister copper contains 98% pure copper and 2% impurities and are purified by electrolytic refining.

Electrolytic refining.

This method is used to get metal of high degree of purity. For electrolytic refining of copper, we use

Cathode: A thin plate of pure copper metal.

Anode: A block of impure copper metal.

Electrolyte: Copper sulphate solution acidified with sulphuric acid. When electric current is passed through the electrolytic

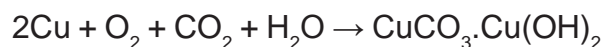
solution pure copper gets deposited at the cathode, impurities settled at the bottom of the anode in the form of sludge called **anode mud**.

Properties

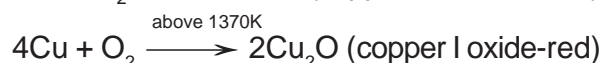
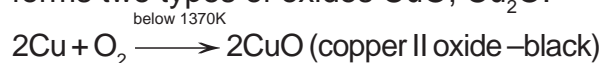
Physical properties: Copper is a reddish brown metal, with high lustre, high density and high melting point (1356°C).

Chemical properties:

i. Action of air and moisture: Copper gets covered with a green layer of basic copper carbonate in the presence of CO₂ and moisture.

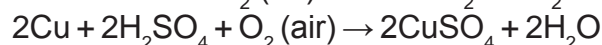


ii. Action of Heat: On heating at different temperatures in the presence of oxygen it forms two types of oxides CuO, Cu₂O.

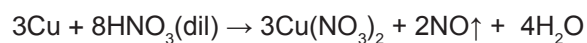


iii. Action of Acids: a) with dil. HCl and dil. H₂SO₄

Dilute acids such as HCl and H₂SO₄ have no action on these metals in the absence of air. Copper dissolves in these acids in the presence of air.

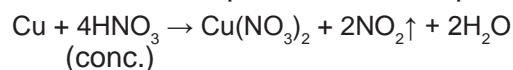


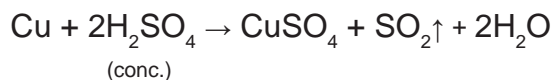
b) with dil. HNO₃ Copper reacts with dil. HNO₃ with the liberation of Nitric Oxide gas.



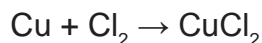
c) with con. HNO₃ and con. H₂SO₄

Copper reacts with con. HNO₃ and con. H₂SO₄ with the liberation of nitrogen dioxide and sulphur dioxide respectively.





iv. Action of chlorine: Chlorine reacts with copper, resulting in the formation of copper (II) chloride.



v. Action of alkalis: Copper is not attacked by alkalis.

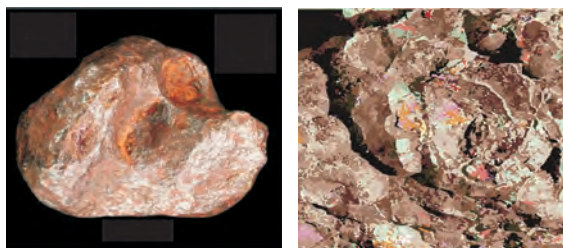
Uses

- It is extensively used for making electric cables and other electric appliances.
- It is used for making utensils, containers, calorimeters, coins.
- It is used in electroplating.
- It is alloyed with gold and silver for making coins and jewels.

PROJECT

Students may be asked to submit a project report on the important applications of copper in everyday life along with the samples.

12.6.3 METALLURGY OF IRON



Symbol	:	Fe
Colour	:	Greyish white
Atomic mass	:	55.9
Atomic number:		26
Valency	:	2 & 3
Electronic configuration	:	2, 8, 14, 2

Occurrence:

Iron is the second most abundant metal after aluminium. It occurs in nature as oxides, sulphides and carbonates. The ores of iron are given in the following table:

Ores of iron	Formula
i. Red haematite	Fe_2O_3
ii. Magnetite	Fe_3O_4
iii. Iron pyrites	FeS_2

Extraction of Iron from haematite ore (Fe_2O_3)

1. Concentration by gravity separation

The powdered ore is washed with stream of water. As a result, the lighter sand particles and other impurities are washed away and heavier ore particles settle down.

2. Roasting and calcination

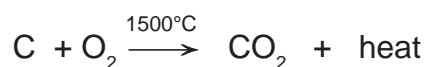
The concentrated ore is strongly heated in a limited supply of air in a reverberatory furnace. As a result, moisture is driven out and sulphur, arsenic, phosphorus impurities are oxidised off.

3. Smelting (in Blast furnace)

The **charge** consisting of roasted ore, coke and limestone in the ratio **8 : 4 : 1** is smelted in a blast furnace by introducing it through the **cup and cone** arrangement at the top. There are three important regions in the furnace.

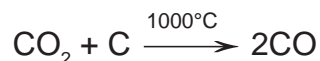
i. The lower region (combustion zone)-temperature is at 1500°C .

In this region, coke burns with oxygen to form CO_2 when the charge comes in contact with the hot blast of air.



It is an exothermic reaction since heat is liberated.

ii. The middle region (fusion zone)-The temperature prevails at 1000°C . In this region CO_2 is reduced to CO .



Limestone decomposes to calcium oxide and CO_2 .



These two reactions are endothermic due to the absorption of heat. Calcium oxide combines with silica to form calcium silicate slag.



iii. The upper region (reduction zone)-temperature prevails at 400°C . In this region carbon monoxide reduces ferric oxide to form a fairly pure spongy iron.



The molten iron is collected at the bottom of the furnace after removing the slag.

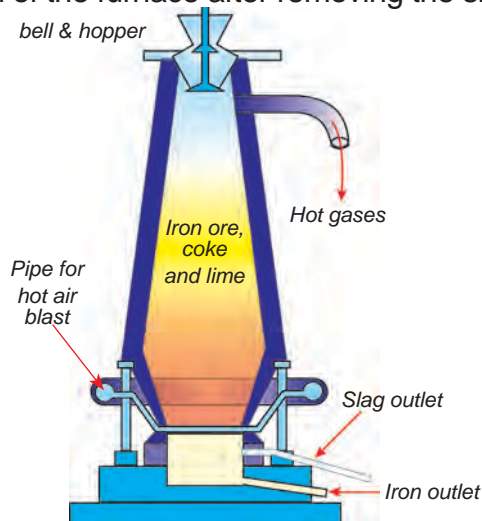


Fig. 12.8.3 Blast furnace

The iron thus formed is called **pig iron**. It is remelted and cast into different moulds. This iron is called **cast iron**.

MORE TO KNOW

CALCINATION AND ROASTING

CALCINATION: It is a process in which ore is heated in the absence of air. As a result of calcinations the carbonate ore is converted into its oxide.

ROASTING: It is a process in which ore is heated in the **presence of excess of air**. As a result of roasting the sulphide ore is converted into its oxide.

MORE TO KNOW

Depending upon the carbon content iron is classified into 3 types.

Pig iron with carbon content of 2- 4.5%

Wrought iron with carbon content <0.25%

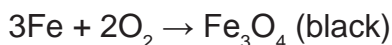
Steel with carbon content of 0.25-2%.

Physical properties

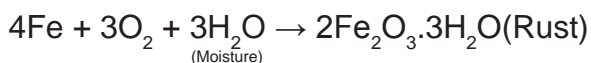
- It is a heavy metal of specific gravity 7.9 g/cc
- It is a lustrous metal and greyish white in colour.
- It has high tensility, malleability and ductility.
- It is a good conductor of heat and electricity.
- It can be magnetised.

Chemical properties

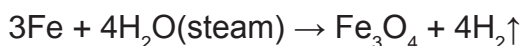
1.Reaction with air or oxygen: Only on heating in air, iron forms magnetic oxide



2.Reaction with moist air: When iron is exposed to moist air, it forms a layer of brown hydrated ferric oxide on its surface. This compound is known as rust and the phenomenon of forming this rust is known as rusting.



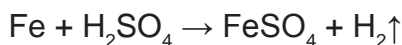
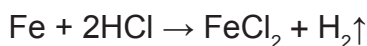
3.Reaction with steam: When steam is passed over red hot iron, magnetic oxide of iron is formed.



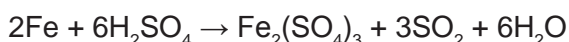
4.Reaction with chlorine: Iron combines with chlorine to form ferric chloride.



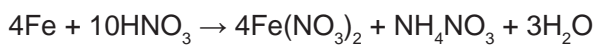
5.Reaction with acids: With dilute HCl and dilute H_2SO_4 it evolves H_2 gas



With conc. H_2SO_4 it forms ferric sulphate



With dilute HNO_3 in cold condition it gives ferrous nitrate



When iron is dipped in conc. HNO_3 it becomes chemically **inert or passive** due to the formation of a layer of iron oxide (Fe_3O_4) on its surface.

Uses of iron

i.Pig iron is used in making pipes, stoves, radiators, railings, man hole covers and drain pipes.

ii. Steel is used in the construction of

buildings, machinery, transmission and T.V towers and in making alloys.

iii.Wrought iron is used in making springs, anchors and electromagnets.

12.7 ALLOYS

An alloy is a homogeneous mixture of a metal with other metals or with non-metals that are fused together.

Alloys are solid solutions. Alloys can be considered as solid solutions in which the metal with high concentration is **solvent** and the metal with low concentration is **solute**. For example, brass is an alloy of zinc(solute) in copper(solvent).

12.7.1 Methods of making alloys:

- 1.By fusing the metals together.
- 2.By compressing finely divided metals one over the other.

Amalgam:An amalgam is an alloy of mercury with metals such as sodium, gold, silver, etc.,

MORE TO KNOW

DENTAL AMALGAMS

It is an alloy of mercury with silver and tin metals. It is used in dental filling.



Dental amalgam

12.7.2 Copper Alloys

Name of the alloy	Reason for alloying	Uses
i.Brass(Cu,Zn)	Lustrous,easily cast,malleable, ductile,harder than Cu.	Electrical fittings, medals, hard ware, decorative items.
ii.Bronze(Cu,Sn,Zn)	Hard,brittle,takes up polish.	Statues, coins, bells, gongs.

12.7.3 Aluminium Alloys

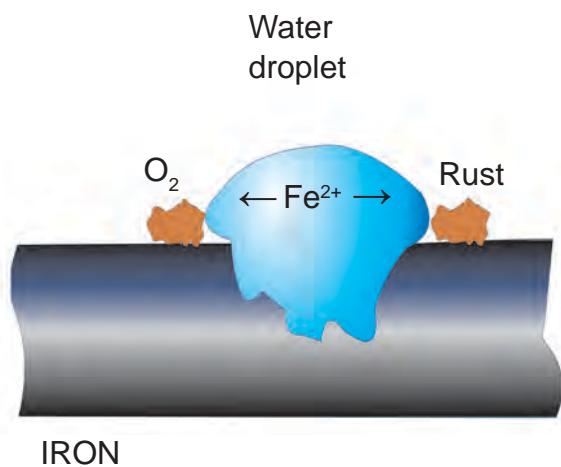
Name of the alloy	Reason for alloying	Uses
i.Duralumin(Al,Mg,Mn,Cu)	Light,strong,resistant to corrosion stronger than aluminium.	Aircraft,tools,pressure cookers
ii.Magnalium(Al,Mg)	Light,hard,tough,corrosion resistant.	Aircraft,scientific instrument

12.7.4 Iron Alloys

Name of the alloy	Reason for alloying	Uses
i.Stainless steel (Fe,C,Ni,Cr)	Lustrous,corrosion resistant,high tensile strength.	Utensils,cutlery,automobile parts.
ii.Nickel steel (Fe,C,Ni)	Hard, corrosion resistant,elastic.	Cables,aircraft parts,propeller.

12.8 CORROSION

Corrosion is defined as the slow and steady destruction of a metal by the environment. It results in the deterioration of the metal to form metal compounds by means of chemical reactions with the environment.



Rusting of iron

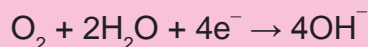
MORE TO KNOW

MECHANISM OF CORROSION

Corrosion is a simple electrochemical reaction.

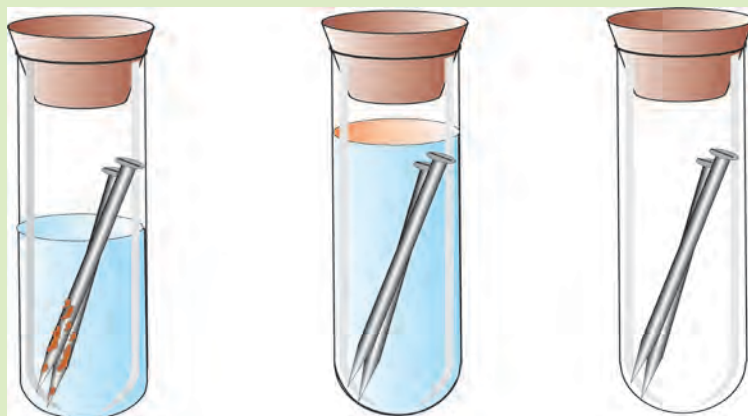
When the surface of iron is in contact with a piece of carbon and water, iron acts as the anode and the carbon acts as a cathode. CO₂ from air dissolves in water to form carbonic acid (H₂CO₃). This acid acts as an electrolyte.

The electrochemical reactions are as follows:



The Fe²⁺ ions are oxidised to Fe³⁺ ions. The Fe³⁺ ions combine with OH⁻ ions to form Fe(OH)₃. This becomes **rust (Fe₂O₃.xH₂O)** which is hydrated ferric oxide.

ACTIVITY 9.1



The conditions for rusting

Take three test tubes provided with rubber corks and label them as A, B and C. Place few iron nails of same size in these tubes. Pour some water in test tube A, some boiled water along with turpentine oil in test tube B and anhydrous CaCl_2 in test tube C. Keep them under observation for few days. Notice the changes.

The nails in A are rusted while the nails in B and C are unaffected.

The rusting of nails in A is due to air and water. In B, the oily layer above water does not allow air to come in contact with nails. In C, the substance anhydrous CaCl_2 has absorbed moisture completely. This activity shows that rusting of iron requires air and water.

12.8.1 Methods of preventing corrosion:

Corrosion of metals is prevented by not allowing them to come in contact with moisture, CO_2 and O_2 . This is achieved by the following methods:

- **By coating with paints:** Paint coated metal surfaces keep out air and moisture.
- **By coating with oil and grease:** Application of oil and grease on the surface of iron tools prevents them from moisture and air.
- **By alloying with other metals:** Alloyed metal is more resistant to corrosion.
- **Example:** stainless steel.
- **By the process of galvanization:** This is a process of coating zinc on iron sheets by using electric current. In this zinc forms a protective layer of zinc carbonate on the surface of iron. This prevents corrosion.
- **Electroplating:** It is a method of coating one metal with another by passing electric current. Example: silver plating, nickel plating. This method not only lends protection but also enhances the metallic appearance.
- **Sacrificial protection:** Magnesium is more reactive than iron. When it is coated on the articles made of steel it sacrifices itself to protect the steel.

EVALUATION

PART A

- In the modern periodic table periods and groups are given. Periods and groups indicate——
a) Rows and Columns b) Columns and rows
- Third period contains 8 elements, out of these elements how many elements are non-metals?
- An element which is an essential constituent of all organic compounds belongs to _____ group. (14th group / 15th group)
- Ore is used for the extraction of metals profitably. Bauxite is used to extract aluminium, it can be termed as _____. (ore / mineral)
- Gold does not occur in the combined form. It does not react with air (or) water. It is in _____. (native state / combined state)

PART B

- Assertion:** Greenish layer appears on copper vessels if left uncleaned.
Reason: It is due to the formation of layer of basic copper carbonate
Give your correct option
a) assertion and reason are correct and relevant to each other
b) assertion is true but reason is not relevant to the assertion
- A process employed for the concentration of sulphide ore is

- _____.
- (froth floatation / gravity separation)
- Coating the surface of iron with other metal prevents it from rusting. If it is coated with thin layer of zinc it is called _____ (galvanization / painting / cathodic protection)
 - Any metal mixed with mercury is called amalgam. The amalgam used for dental filling is _____. (Ag – Sn amalgam / Cu – Sn amalgam)
 - Assertion:** In thermite welding, aluminium powder and Fe_2O_3 are used. **Reason:** Aluminium powder is a strong reducing agent. Does the reason satisfy the assertion?

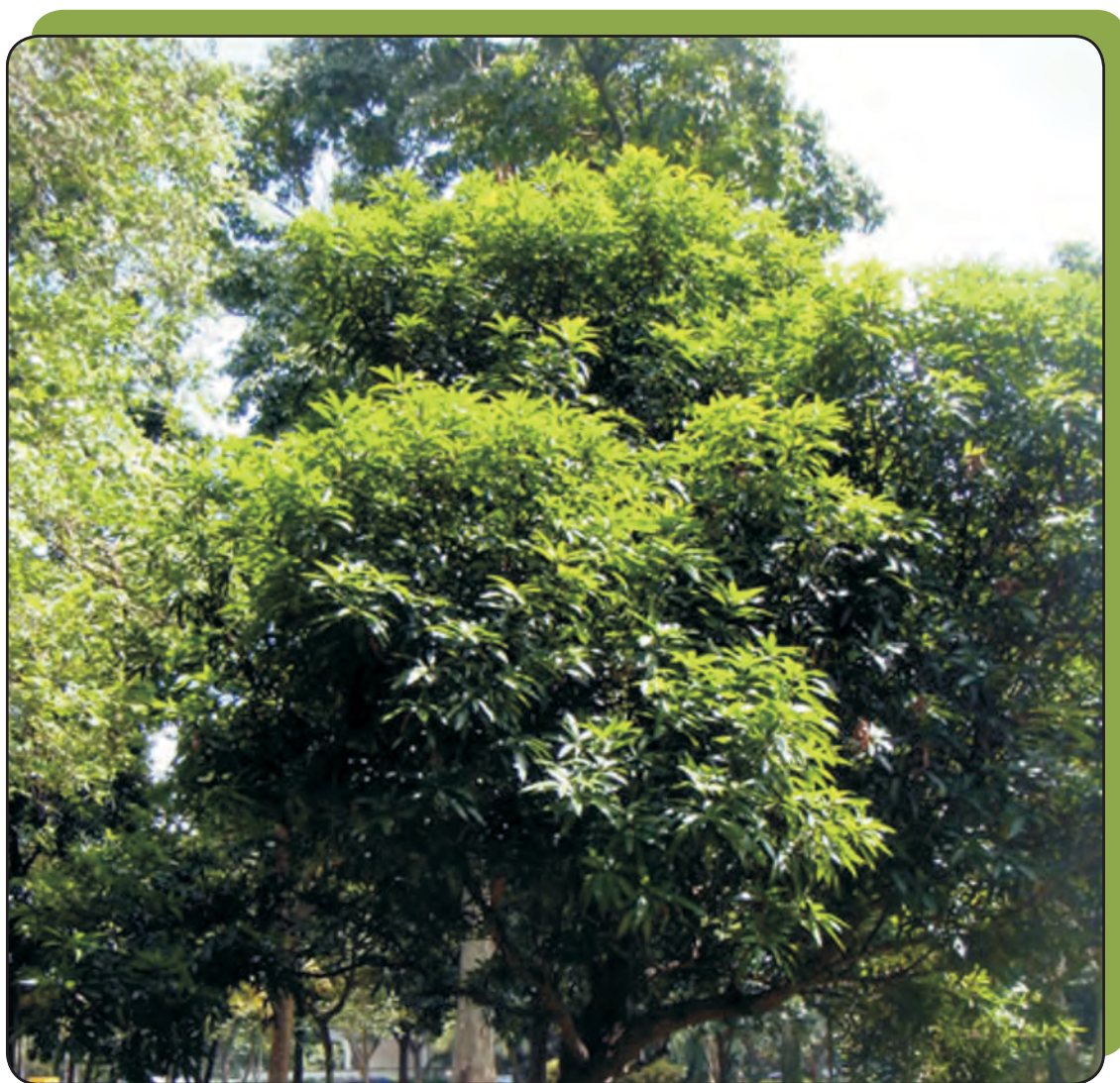
PART C

- Can rusting of iron nail occur in distilled water. Justify your answer.
- Why cannot aluminium metal be obtained by the reduction of aluminium oxide with coke?
- Iron reacts with con. HCl and con. H_2SO_4 . But it does not react with con. HNO_3 . Suggest your answer with proper reason.
- To design the body of the aircraft aluminium alloys are used. Give your reason.
- X is a silvery white metal. X reacts with oxygen to form Y. The same compound is obtained from the metal on reaction with steam with the liberation of hydrogen gas. Identify X and Y.

FURTHER REFERENCE:

Books: Text Book of Inorganic chemistry – P.L. Soni S.Chand Publishers

Website: www.tutorvista.com. www.sciencebyjones.com
<http://www.khanacademy.org>



CARBON AND ITS COMPOUNDS

13. Carbon and its compounds

Symbol	:	C
Atomic Number	:	6
Atomic Mass	:	12
Valency	:	4

The electronic configuration of carbon is $K=2, L=4$. It has four electrons in the valence shell and belongs to group IV A (group 14) of the periodic table.

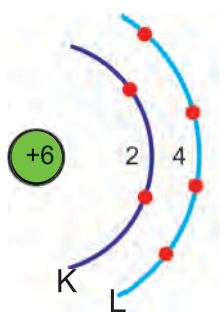


Fig. 13.1 electronic configuration of carbon

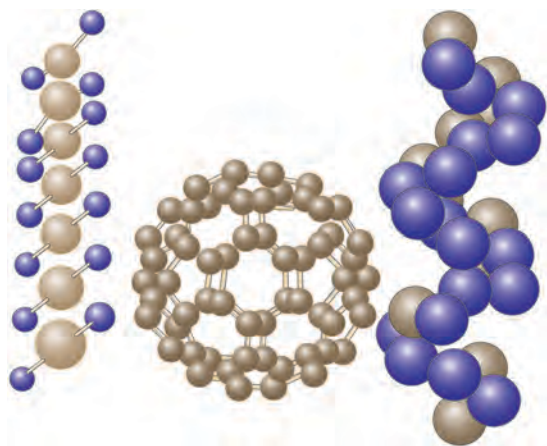


Fig. 13.2 An arrangement depicting carbon and its compounds.

INTRODUCTION

Without carbon, no living thing could survive. Human beings are made up of carbon compounds. Carbon is a non metal. In nature, it occurs in its pure form as **diamond and graphite**. When fuels burn, the carbon in them reacts with oxygen to form carbon dioxide.

Carbon compounds hold the key to plant and animal life on earth. Hence, carbon chemistry is called **Living Chemistry**. Carbon circulates through air, plants, animals and soil by means of complex reactions. This is called **carbon cycle**.

13.1. COMPOUNDS OF CARBON

In the beginning of 19th century scientists classified the compounds of carbon into two types, based on their source of occurrence:

- Inorganic compounds (obtained from non living matter)
- Organic compounds (obtained from living matter, such as plant and animal sources) however the basis of classification was subjected to alteration after wöhler synthesis.

LIVING CHEMISTRY

All living organisms are made of carbon atoms. This means that, carbon atoms form the building blocks for living organisms. These carbon atoms, in combination with other atoms decide life on earth. Hence **carbon chemistry** is also called as **living chemistry**.



Fig. 13.3



Fig. 13.4

FRIEDRICH WOHLER

A creator of revolution in **ORGANIC CHEMISTRY**

MORE TO KNOW

ORGANIC CHEMISTRY:

The word organic signifies life. The term organic chemistry was used by the Swedish chemist Berzelius. This refers to the chemistry of living things. However, the German chemist Wohler succeeded in creating an organic compound (urea) from an inorganic compound (ammonium cyanate) in his laboratory. This has dealt a severe blow to the Vital force theory (a theory of life process).



FRIEDRICH WOHLER
A German Chemist

13.2. MODERN DEFINITION OF ORGANIC CHEMISTRY

Organic chemistry is defined as the branch of chemistry that deals with organic compounds which are made up of the hydrocarbons and their derivatives. It gives a thorough insight into the nature of bonding, synthesis, characteristics and their usefulness in various fields.

MORE TO KNOW



A polished diamond



The most precious diamond is a crystalline allotrope of carbon. KOHINOOR DIAMOND is a 105 carat diamond (21.68g) It was seized by the EAST INDIA COMPANY and became the part of British Crown Jewels. May it be an ordinary coal or the most precious Kohinoor diamond, it is an allotropic modification of carbon indeed!

13.3. BONDING IN CARBON AND ITS COMPOUNDS

The atomic number of carbon is 6 and its ground state electronic configuration is $1s^2 2s^2 2p^2$. Since it has four electrons in its outermost shell, its valency is four. To achieve noble gas configuration, carbon atom has to lose or gain four electrons to form C^{4+} and C^{4-} ions.

1. It could gain four electrons forming C^{4-} anion, but it would be difficult for the nucleus with six protons to hold on to ten electrons i.e. four extra electrons.
2. It could lose four electrons to form C^{4+} cations, but it would require a large amount of energy to remove four electrons leaving behind the carbon cations with six protons in its nucleus holding on to just two electrons.

Carbon overcomes this problem by sharing its valence electrons with other atoms of carbon or with atoms of other elements. *This characteristic of carbon atom by virtue of which it forms four covalent bonds is generally referred as tetra valency of carbon.*

A molecule of methane (CH_4) is formed when four electrons of carbon are shared with four hydrogen atoms.

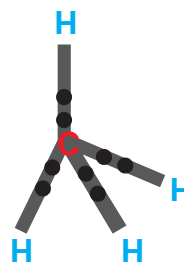


Fig. 13.5 Structure of methane

•• Represents shared pair of electrons

13.4 ALLOTROPY

Allotropy is defined as the property by which an element can exist in more than one form that are physically different but chemically similar.

Allotropes of carbon

- Carbon exists in three allotropic forms. They are crystalline form (diamond and graphite), amorphous form (coke, charcoal) and fullerene.

- In diamond each carbon atom is bonded to four other carbon atoms forming a rigid three dimensional structure, accounting for its hardness and rigidity.

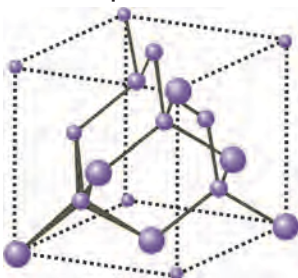


Fig. 13.6

Structure of diamond

- In graphite each carbon atom is bonded to three other carbon atoms in the same plane giving hexagonal layers held together by weak **vander Waals forces** accounting for softness.
- Graphite is a good conductor of electricity unlike other non-metals since it has free electrons in it.

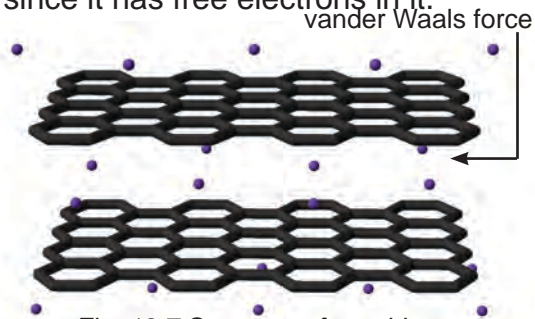


Fig. 13.7 Structure of graphite

- Fullerenes form another type of carbon allotropes. The first one was identified to contain 60 carbon atoms in the shape of a football. (C-60).

Since this looks like the geodesic dome designed by the US architect Buckminster Fuller, it is named as

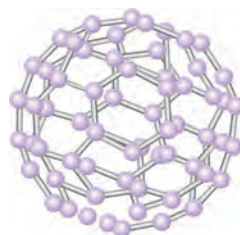


Fig. 13.8 Fullerene



Fig. 13.9 Foot ball

Buckminster Fullerene.

13.5 Physical nature of carbon and its compounds :

- Carbon has the ability to form covalent bonds with other atoms of carbon giving rise to large number of molecules through self linking property. This property is called **catenation**. Since the valency of carbon is four, it is capable of bonding with four other atoms.
- Carbon combines with oxygen, hydrogen, nitrogen, sulphur, chlorine and many other elements to form various stable compounds.
- The stability of carbon compounds is due to the small size of carbon which enables the nucleus to hold on to the shared pair of electrons strongly.
- Carbon compounds show **isomerism**, the phenomenon by which two or more compounds to have same molecular formula but different structural formula with difference in properties. i.e the formula C_2H_6O represents two different compounds namely ethyl alcohol (C_2H_5OH) and dimethyl ether (CH_3OCH_3).
- Carbon compounds have low melting

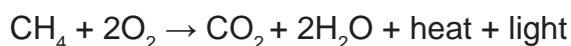
and boiling points because of their covalent nature.

- The reactions shown by carbon compounds involve breaking of old bonds in the reacting molecules and the formation of new bonds in the product molecules.
- Carbon compounds are easily combustible.

13.6 CHEMICAL PROPERTIES

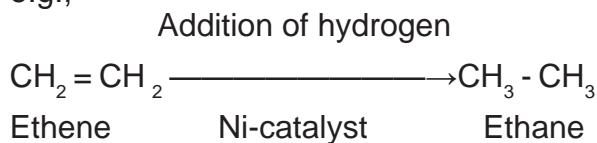
- Carbon and its compounds burn in oxygen to give carbon dioxide along with heat and light.

e.g.,



- Carbon compounds can be easily oxidized using suitable oxidizing agent Alkaline potassium permanganate to form carboxylic acids.
- Unsaturated carbon compounds undergo addition reactions with hydrogen in the presence of palladium or nickel catalyst.

e.g.,



- Carbon compounds undergo substitution reactions in the presence of either sunlight or any other reagents. e.g., methane undergoes substitution reaction to form different types of products.
- Carbon compounds such as alcohols react with sodium to liberate hydrogen gas.

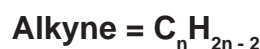
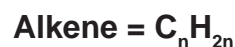


13.7 HOMOLOGOUS SERIES

A homologous series is a group or a class of organic compounds having similar structure and similar chemical properties in which the successive compounds differ by a CH_2 group.

13.7.1 Characteristics of homologous series

- Each member of the series differs from the preceding or succeeding member by a common difference of CH_2 and by a molecular mass of 14 amu (amu = atomic mass unit).
- All members of homologous series contain same elements and the same functional groups.
- All members of homologous series have same general molecular formula.



- The members in homologous series show a regular gradation in their physical properties with respect to increase in molecular mass.
- The chemical properties of the members of the homologous series are similar.
- All members of homologous series can be prepared by using same general method.

13.8 IMPORTANCE OF HOMOLOGOUS SERIES

1. It helps to predict the properties of the members of the series that are yet to be prepared.
2. Knowledge of homologous series gives a systematic study of the members.
3. The nature of any member of the family can be ascertained if the properties of the first member are known.

13.9 HYDROCARBONS

The simplest organic compounds containing only carbon and hydrogen are called **Hydrocarbons**. These are regarded as the **parent organic compounds** and all other compounds are considered to be derived from them by the replacement of one or more hydrogen atoms by other atoms or groups of atoms.

Hydro carbons are classified into two types: *saturated and unsaturated hydrocarbons*.

13.9.1 Saturated hydrocarbons – Alkanes

General formula = C_nH_{2n+2} Suffix : ane

These are the organic compounds which contain carbon – carbon single bond. These were earlier named as **paraffins** (Latin : meaning little affinity) due to their least chemical reactivity. According to IUPAC system, these are named as **alkanes** (ane is suffix with root word).

Formula	Common name	IUPAC name
CH_4	Methane	Methane
CH_3CH_3	Ethane	Ethane
$CH_3CH_2CH_3$	Propane	Propane
$CH_3CH_2CH_2CH_3$	n-Butane	Butane

13.9.2 Unsaturated hydrocarbons

These are hydrocarbons which contain carbon to carbon double bonds $\left(\begin{array}{c} | \quad | \\ -C=C- \\ | \quad | \end{array} \right)$ or carbon to carbon triple bonds $-C\equiv C-$ in their molecules. These are further classified into two types: **alkenes and alkynes**.

i) **Alkenes: General formula: C_nH_{2n} Suffix: ene**

The hydrocarbons containing at least one carbon to carbon double bond are called **alkenes**. They have the general formula C_nH_{2n} . These were previously called **olefins** (Greek : olefiant – oil forming) because the lower gaseous members of the family form oily products when treated with chlorine.

In IUPAC system, the name of alkene is derived by replacing suffix **ane** of the corresponding alkane by **ene**. For example,

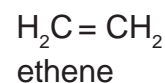
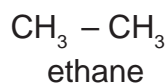


Fig. 14.0 Bromine Test

(Left) No change in colour - saturated,
(Right) Decolouration occurs - unsaturated

In higher alkenes, the position of the double bond, can be indicated by assigning numbers 1, 2, 3, 4,to the carbon atoms present in the molecule.

Alkene	Common name	IUPAC name
$\text{CH}_2 = \text{CH}_2$	Ethylene	Ethene
$\text{CH}_3\text{CH} = \text{CH}_2$	Propylene	Propene
$\text{CH}_3\text{CH}_2\text{-CH} = \text{CH}_2$	α -Butylene	But-1-ene
$\text{CH}_3\text{CH} = \text{CHCH}_3$	β -Butylene	But-2-ene

ii) **Alkynes: General formula: $\text{C}_n\text{H}_{2n-2}$ Suffix : yne**

The hydrocarbons containing carbon to carbon triple bond are called **alkynes**. Alkynes are named in the same way as alkenes i.e., by replacing suffix **ane** of alkane by **yne**. In higher members, the position of triple bond is indicated by giving numbers 1, 2, 3, 4,to the carbon atom in the molecule.

Alkyne	Common name	IUPAC name
$\text{HC} \equiv \text{CH}$	Acetylene	Ethyne
$\text{H}_3\text{C} - \text{C} \equiv \text{CH}$	Methyl acetylene	Propyne
$\text{H}_3\text{C} - \text{C} \equiv \text{C} - \text{CH}_3$	Dimethyl acetylene	But-2-yne
$\text{H}_3\text{C} - \text{CH}_2 - \text{C} \equiv \text{CH}$	Ethyl acetylene	But-1-yne

13.10. FUNCTIONAL GROUP

Functional group may be defined as an atom or group of atoms or reactive part which is responsible for the characteristic properties of the compounds.

The chemical properties of organic compounds are determined by the functional groups while their physical properties are determined by the remaining part of the molecule.

Example: $-\text{OH} \Rightarrow$ Alcohol

$-\text{CHO} \Rightarrow$ Aldehyde

$\text{>C=O} \Rightarrow$ Ketone

$-\text{COOH} \Rightarrow$ Carboxylic acid

13.10.1. Classification of organic compounds based on functional group

1. Alcohols

Alcohols are carbon compounds containing $-\text{OH}$ group attached to alkyl group. The general formula of alcohol is R-OH where 'R' is an **alkyl group** and $-\text{OH}$ is the **functional group**. The IUPAC name of alcohol is derived by replacing $-\text{e}$, in the word **alkane**, by the suffix $-\text{ol}$. Hence we get the name **alkanol**.

Molecular formula	Common name	IUPAC name
CH ₃ OH	Methyl alcohol	Methanol
CH ₃ -CH ₂ -OH	Ethyl alcohol	Ethanol
CH ₃ -CH ₂ -CH ₂ -OH	n-Propyl alcohol	1-Propanol
$\begin{array}{c} \text{CH}_3\text{-CH-CH}_3 \\ \\ \text{OH} \end{array}$	Isopropyl alcohol or secondary propyl alcohol	2-Propanol
CH ₃ -CH ₂ -CH ₂ -CH ₂ -OH	n-Butyl alcohol	1-Butanol
$\begin{array}{c} \text{CH}_3\text{-CH-CH}_2\text{-OH} \\ \\ \text{CH}_3 \end{array}$	Isobutyl alcohol	2-Methyl-1-propanol

2. Aldehydes

Aldehydes are carbon compounds containing **-CHO** group attached to alkyl group or hydrogen atom. The general formula of aldehydes is **R – CHO** where '**R**' is an **alkyl group** or **hydrogen atom** and **-CHO** is the **functional group**. The IUPAC name of aldehyde is derived by replacing **-e**, in the word alkane, by the suffix **-al**. Hence we get the name "**alkanal**".

Molecular formula	Common name	IUPAC name
HCHO	Formaldehyde	Methanal
CH ₃ -CHO	Acetaldehyde	Ethanal
CH ₃ -CH ₂ -CHO	Propionaldehyde	Propanal
CH ₃ -CH ₂ -CH ₂ -CHO	Butyraldehyde	Butanal

3. Ketones

Ketones are carbon compounds containing carbonyl **-CO-** group attached to two alkyl groups. The general formula of ketone is **R-CO-R'** where **R** and **R'** are **alkyl groups** and **-CO-** is the **functional group**. The IUPAC name of ketone is derived by replacing **-e**, in the word alkane, by the suffix **-one**. Hence we get the name "**alkanone**".

Molecular formula	Common name	IUPAC name
CH_3COCH_3	Dimethyl ketone (Acetone)	Propanone
$\text{CH}_3\text{COCH}_2\text{CH}_3$	Ethyl methyl ketone	2-Butanone
$\text{CH}_3\text{CH}_2\text{COCH}_2\text{CH}_3$	Diethyl ketone	3-Pentanone

4. Carboxylic Acids

Carboxylic acids are carbon compounds containing $-\text{COOH}$ group attached to a hydrogen atom or alkyl group. The general formula of acid is $\text{R}-\text{COOH}$ where 'R' is a **hydrogen atom** or **alkyl group** and $-\text{COOH}$ is the **functional group**. The IUPAC name of acid is derived by replacing -e, in the word alkane, by the suffix -oic acid. Hence we get the name "**alkanoic acid**".

Molecular formula	Common name	IUPAC name
HCOOH	Formic acid	Methanoic acid
CH_3-COOH	Acetic acid	Ethanoic acid
$\text{CH}_3-\text{CH}_2-\text{COOH}$	Propionic acid	Propanoic acid
$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{COOH}$	n-Butyric acid	Butanoic acid

SOME IMPORTANT ORGANIC COMPOUNDS

Almost all the compounds are useful to us in a number of ways. Most of the fuels, medicines, paints, explosives, synthetic polymers, perfumes and detergents are basically organic compounds. In fact, organic chemistry has made our life colourful and also comfortable. Two commercially important compounds, ethanol and ethanoic acid are briefly discussed here.

13.11 ETHANOL ($\text{C}_2\text{H}_5\text{OH}$)

Ethanol or ethyl alcohol or simply alcohol is one of the most important members of the family of alcohols.

(1) Manufacture of ethanol from molasses

Molasses is a dark coloured syrupy liquid left after the crystallization of sugar from the concentrated sugar cane juice. Molasses still contain about 30% of sucrose which cannot be separated by crystallization. It is converted into ethanol by the following steps:

(i) Dilution

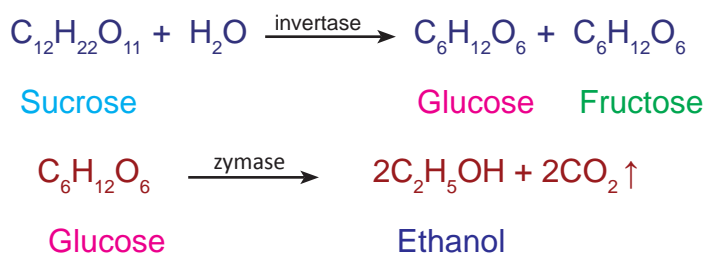
Molasses is first diluted with water to bring down the concentration of sugar to about 8 to 10 percent.

(ii) Addition of ammonium salts

Molasses usually contains enough nitrogenous matter to act as food for yeast during fermentation. If the nitrogen content of the molasses is poor, it may be fortified by the addition of ammonium sulphate or ammonium phosphate.

(iii) Addition of yeast

The solution from step (ii) is collected in large 'fermentation tanks' and yeast is added to it. The mixture is kept at about 303K for a few days. During this period, the enzymes invertase and zymase present in yeast, bring about the conversion of sucrose into ethanol.



The fermented liquid is technically called wash.

(iv) Distillation of wash

The fermented liquid containing 15 to 18 percent alcohol and the rest of the water, is now subjected to fractional distillation. The main fraction drawn, is an aqueous solution of ethanol which contains 95.5% of ethanol and 4.5% of water. This is called rectified spirit. This mixture is then heated under reflux over quicklime for about 5 to 6 hours and then allowed to stand for 12 hours. On distillation of this mixture, pure alcohol (100%) is obtained. This is called absolute alcohol.

MORE TO KNOW

FERMENTATION :

The slow chemical change taking place in an organic compound by the action of enzymes leading to the formation of smaller molecules is called fermentation.

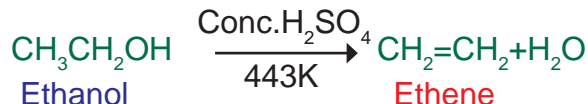
2. Physical properties

- (i) Ethanol is a clear liquid with burning taste.
- (ii) Its boiling point is 351.5 K which is higher than corresponding alkane.
- (iii) It is completely miscible with water in all proportions.

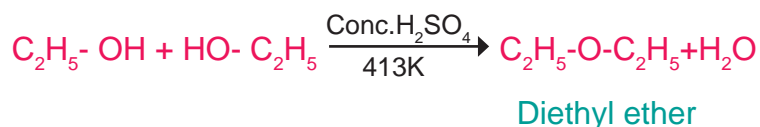
3. Chemical properties

(i) Dehydration

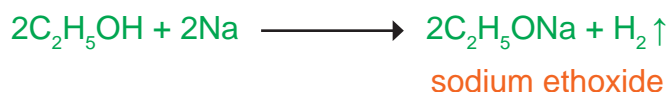
- (a) **Intra molecular dehydration** : Ethanol, when heated with excess conc. H_2SO_4 at 443 K undergoes intra molecular dehydration (i.e. removal of water within a molecule of ethanol).



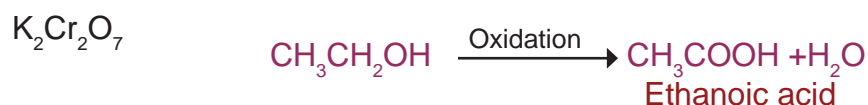
- (b) **Inter molecular dehydration** : When excess of alcohol is heated with conc. H_2SO_4 at 413K two molecules condense by losing a molecule of water to form ether (i.e. removal of water from two molecules of ethanol).



- (ii) **Reaction with sodium** : Ethanol reacts with sodium metal to form sodium ethoxide and hydrogen gas.

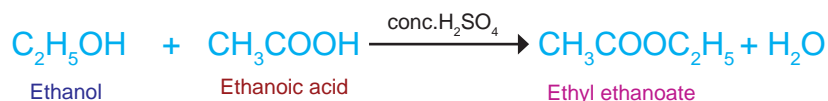


- (iii). **Oxidation** : Ethanol is oxidized to ethanoic acid with alkaline KMnO_4 or acidified

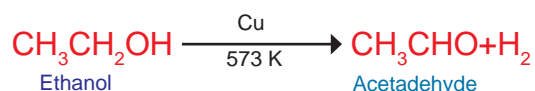


During this reaction, orange colour of $K_2Cr_2O_7$ changes to green. Therefore, this reaction can be used for the **identification of alcohols**.

(iv) Esterification : Ethanol reacts with ethanoic acid in the presence of $\text{conc. H}_2\text{SO}_4$ (catalyst) to form ethyl ethanoate and water. The compound formed by the reaction of an alcohol with carboxylic acid is known as ester (fruity smelling compound) and the reaction is called esterification.



(v). Dehydrogenation : When the vapour of ethanol is passed over reduced copper catalyst at 573 K, it is dehydrogenated to acetaldehyde.



4. Uses

Ethanol is used

1. as an anti-freeze in automobile radiators.
2. as a preservative for biological specimen.
3. as an antiseptic to sterilize wounds in hospitals.
4. as a solvent for drugs, oils, fats, perfumes, dyes, etc.
5. in the preparation of methylated spirit (mixture of 95% of ethanol and 5% of methanol), rectified spirit (mixture of 95.5% of ethanol and 4.5% of water), power alcohol (mixture of petrol and ethanol) and denatured spirit (ethanol mixed with pyridine).
6. in cough and digestive syrups.

Evil effects of consuming alcohol

- If ethanol is consumed, it tends to slow down metabolism of our body and depresses the central nervous system.
- It causes mental depression and emotional disorder.
- It affects our health by causing ulcer, high blood pressure, cancer, brain and liver damage.
- Nearly 40% accidents are due to drunken drive.

- Unlike ethanol, intake of methanol in very small quantities can cause death.
- Methanol is oxidized to methanal (formaldehyde) in the liver and methanal reacts rapidly with the components of cells.
- Methanal causes the protoplasm to get coagulated, in the same way an egg is coagulated by cooking. Methanol also affects the optic nerve, causing blindness.

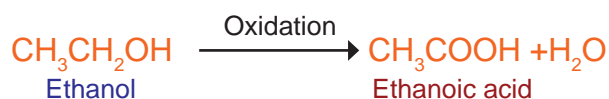
13.12. ETHANOIC ACID (CH₃COOH)

Ethanoic acid is most commonly known as acetic acid and belongs to a group of acids called carboxylic acids. Acetic acid is present in many fruits and sour taste of fruits is because of this acid.

1. Preparation of Ethanoic acid

Ethanol on oxidation in the presence of alkaline potassium permanganate or acidified potassium dichromate gives ethanoic acid.

2. Physical properties



- Ethanoic acid is a colourless liquid and has a sour taste.
- It is miscible with water in all proportions.
- Boiling point (391 K) is higher than corresponding alcohols, aldehydes and ketones.
- On cooling, pure ethanoic acid is frozen to form ice like flakes. They look like glaciers, so it is called glacial acetic acid.

3. Chemical properties

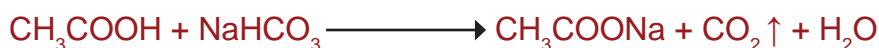
- Ethanoic acid is a weak acid but it turns blue litmus to red.
- Reaction with metal

Ethanoic acid reacts with metals like Na, K, Zn, etc to form metal ethanoate and hydrogen gas.



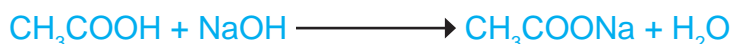
(iii) Reaction with carbonates and bicarbonates.

Ethanoic acid reacts with carbonates and bicarbonates and produces brisk effervescence due to the evolution of carbon dioxide.



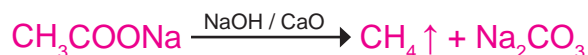
(iv) Reaction with base

Ethanoic acid reacts with sodium hydroxide to form sodium ethanoate and water.



(v) Decarboxylation (Removal of CO_2)

When sodium salt of ethanoic acid is heated with soda lime (Solid mixture of 3 parts of NaOH and 1 part of CaO) methane gas is formed.



4. USES

Ethanoic acid is used

1. for making vinegar which is used as a preservative in food and fruit juices.
2. as a laboratory reagent.
3. for coagulating rubber from latex.
4. in the preparation of dyes, perfumes and medicine.

EVALUATION

PART A

1. Assertion: Chemical bonds in organic compounds are covalent in nature.
Reason: Covalent bond is formed by the sharing of electrons in the bonding atoms.
Does the reason satisfy the given assertion.
2. Assertion: Diamond is the hardest crystalline form of carbon
Reason: Carbon atoms in diamond are tetrahedral in nature (Verify the suitability of reason to the given Assertion mentioned above)

3. Assertion: Due to catenation a large number of carbon compounds are formed.
Reason: Carbon compounds show the property of allotropy.
(Is the reason holding good for the given Assertion)
4. Buckminster fullerene is the allotropic form of (Nitrogen / Carbon / Sulphur)
5. Eventhough it is a non metal, graphite conducts electricity. It is due to the presence of(free electrons / bonded electrons)
6. Formula of methane is CH_4 and its succeeding member ethane is expressed in C_2H_6 . The common difference of succession between them is (CH_2 / C_2H_2)
7. IUPAC name of first member of alkyne is (ethene / ethyne)
8. Out of ketonic and aldehydic group which is the terminal functional group?
9. Acetic acid is heated with a solid 'X' kept in a test tube. A colourless and odourless gas (Y) is evolved. The gas turns lime water milky when passed through it. Identify X and Y.
10. Assertion: Denaturation of ethyl alcohol makes it unfit for drinking purposes.
Reason: Denaturation of ethyl alcohol is carried out by methyl alcohol.
Check whether the reason is correct for assertion.

PART B

11. Write down the possible isomers and give their IUPAC names using the formula C_4H_{10} .
12. Diamond is the hardest allotrope of Carbon. Give reason for its hardness.
13. An organic compound (A) is widely used as a preservatives in pickles and has a molecular formula $\text{C}_2\text{H}_4\text{O}_2$. This compound reacts with ethanol to form a sweet smelling compound (B).
 - (i) Identify the compound A and B.
 - (ii) Name the process and write corresponding chemical equation.
4. An organic compound (A) of molecular formula $\text{C}_2\text{H}_6\text{O}$ on oxidation with alkaline KMnO_4 solution gives an acid (B) with the same number of carbon atoms. Compound A is used as an antiseptic to sterilize wounds in hospitals. Identify A and B. Write the chemical equation involved in the formation of B from A.

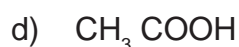
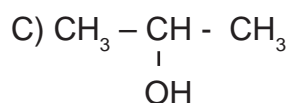
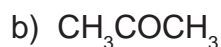
PART C

1. Fill in the blanks using suitable formula in the given table

No.	Alkane	Alkene	Alkyne
1.	C_2H_6 ethaneethene	C_2H_2 ethyne
2.Propane	C_3H_6 Propenepropyne
3.	C_4H_{10} ButaneButeneButyne

2. Homologous series predict the properties of the members of hydrocarbon. Justify this statement through its characteristics.

3. Write the common name and IUPAC name of the following.



FURTHER REFERENCE

Books: 1. Organic chemistry - **B.S. Bahl & Arun Bahl** S.Chand Publishers

2. Organic chemistry - **R.T. Morrison & R.N. Boyd** - Practice Hall Publishers.

Website: www.tutorvista.com, www.topperlearning.com



MEASURING INSTRUMENTS

14. Measuring Instruments

Physics is the most basic science, which deals with the study of nature and natural phenomena. It is a quantitative science. Therefore physicists measure things. The ultimate test of any physical quantity is its agreement with observations and measurement of physical phenomena. One of the major contributions of physics to other sciences and society are the many measuring instruments and techniques that physics has developed. One such instrument is screw gauge.

14.1 SCREW GAUGE

Screw Gauge is an instrument to measure the dimensions of very small objects upto 0.001 cm.

The Screw Gauge consists of 'U' shaped metal frame Fig. 14.1.

A hollow cylinder is attached to one end of the frame.

Grooves are cut on the inner surface of the cylinder through which a screw passes through.

On the cylinder parallel to the axis of the screw a scale is graduated in millimeter called Pitch Scale.

One end of the screw is attached to a sleeve.

The head of the sleeve is divided into 100 divisions called as the Head Scale.

The other end of the screw has a plane surface (s_1).

A stud (s_2) is attached to the other end of the frame, just opposite to the tip of the screw.

The screw head is provided with a ratchet arrangement (safety device) to prevent the user from exerting undue pressure.

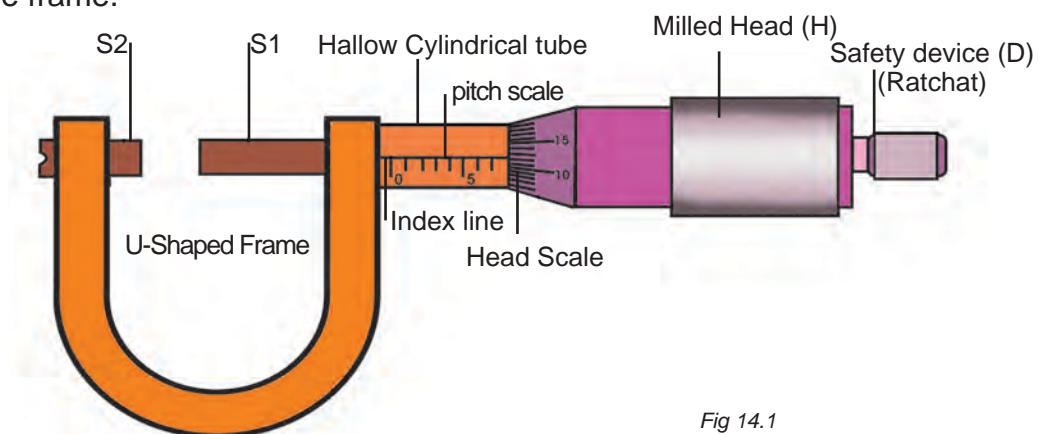


Fig 14.1

Principle of the Screw Gauge

Screw Gauge works under the principle of the screw. When a screw is rotated in a nut, the distance moved by the tip of the screw is directly proportional to the number of rotations.

Pitch of the screw

Pitch of the screw is the distance between two successive screw threads. It is also equal to the distance travelled by the tip of the screw for one complete rotation of the head.

$$\text{Pitch} = \frac{\text{Distance travelled on the pitch scale}}{\text{No. of rotations}}$$

Least Count of a Screw Gauge

The distance moved by the tip of the screw for a rotation of one division on the head scale is called the least count of the Screw Gauge.

$$\text{L.C} = \frac{\text{Pitch}}{\text{No. of divisions on the head scale}}$$

Zero Error of a Screw Gauge

The plane surface of the screw and the opposite plane stud on the frame are brought into contact.

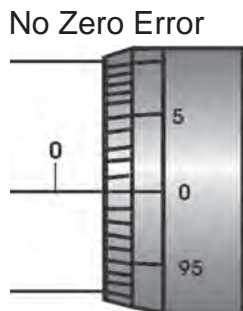


Fig. 14.2

If the zero of the head scale coincides with the pitch scale axis, there is no zero error. Fig. 14.2

Positive Zero Error

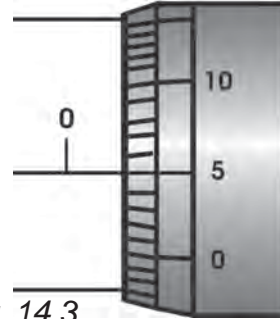


Fig. 14.3

If the zero of the head scale lies below the pitch scale axis, the zero error is positive. If the n^{th} division of the head scale coincides with pitch scale axis the zero error is positive. Fig. 14.3

$$\text{Z.E} = + (n \times \text{L.C}) ,$$

Then the Zero Correction

$$\text{Z.C} = - (n \times \text{L.C})$$

Negative Zero Error



Fig 14.4

If the Zero of the head scale lies above the pitch scale axis, the zero error is negative. If the n^{th} division coincides with the pitch scale axis, the zero error is negative. Fig. 14.4

$$\text{Z.E} = - (100 - n) \times \text{L.C},$$

Then the Zero Correction

$$\text{Z.C} = + (100 - n) \times \text{L.C}$$

To measure the diameter of a thin wire using Screw Gauge

- Determine the Pitch, the Least count and the Zero Error of the Screw Gauge.
- Place the wire between two studs.
- Rotate the head until the wire is held firmly but not tightly, with the help of ratchet.
- Note the reading on the pitch scale crossed by the head scale (P.S.R) and the head scale reading coincides with the head scale axis (H.S.C).
- The diameter of the wire is given by $P.S.R + (H.S.C \times L.C) \pm Z.C$
- Repeat the experiment for different portions of the wire.
- Tabulate the readings.
- The average of the last column reading gives the diameter of the wire.

S.No	P.S.R mm	H.S.C	H.S.C x L.C mm	Total Reading P.S.R + (H.S.C x L.C) ±Z.C mm
1				
2				
3				

Nowadays we have digital Screw Gauge to take the reading at once.

14.2 Measuring long distances

For measuring long distances such as distance of the moon or a planet from the earth, special methods are adopted. Radio echo method, laser pulse method and parallax method are used to determine very long distances. In order to measure such very long distances the units astronomical distance and light year are used.

Astronomical distance

Astronomical distance is the mean distance of the centre of the sun from the centre of the earth.

$$1 \text{ Astronomical unit (AU)} \\ = 1.496 \times 10^{11} \text{ m}$$

Light year

Light year is the distance travelled by light in one year in vacuum.

Distance traveled by light in one year in vacuum = Velocity of light x 1 year (in seconds)

$$= 3 \times 10^8 \times 365.25 \times 24 \times 60 \times 60 \\ = 9.467 \times 10^{15} \text{ m}$$

Therefore , 1 light year = $9.467 \times 10^{15} \text{ m}$

EVALUATION

PART A

1. Screw gauge is an instrument to measure the dimensions of very small objects upto_____

(0.1 cm., 0.01 cm., 0.1 mm., 0.01 mm)

2. In a screw gauge zero of the head scale

lies below the pitch scale axis, the zero error is _____

(positive, negative, nil)

3. Screw gauge is used to measure the diameter of _____

(crow bar, thin wire, cricket ball)

4. One light year is equal to _____

($365.25 \times 24 \times 60 \times 60 \times 3 \times 10^8$ m ,
 $1 \times 24 \times 60 \times 60 \times 3 \times 10^8$ m ,
 $360 \times 24 \times 60 \times 60 \times 3 \times 10^8$ m)

5. One astronomical unit is the distance between the centre of the earth and _____

(centre of the Moon, centre of the Sun, centre of the Mars)

PART B

1. Correct the mistakes if any, in the following statements.

Astronomical distance is the mean distance of the surface of the sun from the surface of the earth.

Light year is the distance travelled by light in one year in vacuum at a speed of 3×10^8 m. per minute

2. Match the items in group A with the items in group B

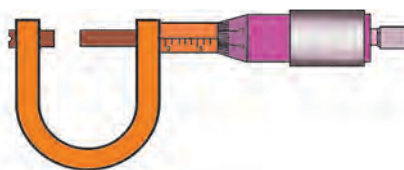
Group A	Group – B
Small dimensions	Kilo meter
Large dimensions	Screw gauge
Long distances	Scale
Small distances	Light year
	Altimeter

3. Fill in the blanks: Special methods adopted to determine very large distances are _____ and _____ (Laser pulse method, Light year method, Radio echo method)

4. Least count of a screw gauge is an important concept related to screw gauge. What do you mean by the term least count of a screw gauge.

5. Label the following parts of the screw gauge in the given screw gauge diagram.

- 1. Head scale 2. Pitch scale
- 3. Axis 4. Ratchet



FURTHER REFERENCE :

- Books:** 1. Complete physics for IGCSE - Oxford publications.
 2. Practical physics – Jerry. D. Wilson – Saunders college publishing
- Webste:** www.complere.com
 www.physlink.com



LAWS OF MOTION AND GRAVITATION

15. Laws of motion and gravitation

In our everyday life, we observe that some effort is required to put a stationary object into motion or to stop a moving object. Normally we have to push or pull or hit an object to change its state of motion.

The concept of force is based on this push, pull or hit. No one has seen, tasted, or felt force. However, we always see or feel the effect of a force. It can only be explained by describing what happens when a force is applied to an object. Push, pull or hit may bring objects into motion, because we make a force to act on them. Therefore, **force is one which changes or tends to change the state of rest or of uniform motion of a body.** Force is a vector quantity. Its SI unit is **newton**.

15.1. BALANCED AND UNBALANCED FORCES

Fig.15.1 shows a wooden block on a horizontal table. Two strings X and Y are tied to the two opposite faces of the block as shown.

If we apply a force by pulling the string 'X', the block begins to move to the right.

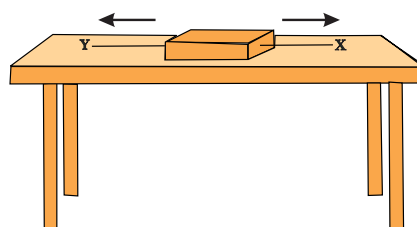


Fig. 15.1

Similarly, if we pull the string Y, the block moves to the left. But, if the block is pulled from both the sides with equal forces the block will not move and remains stationary. Forces acting on an object which do not change the state of rest or of uniform motion of it are called **balanced forces**. Now let us consider a situation in which two opposite forces of different magnitudes act on the block. The block moves in the direction of the greater force. The resultant of two forces acts on an object and brings it in motion. These opposite forces are called **unbalanced forces**.

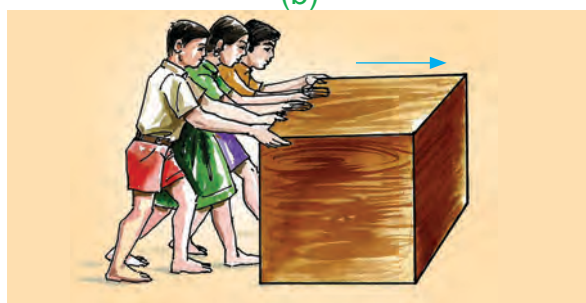
The following illustration clearly explains the concept of balanced and imbalanced forces. Some children are trying to push a box on a rough floor.



(a)



(b)



(c)

Fig. 15.2

If one boy pushes the box with a smaller force, the box does not move because of friction acting in a direction opposite to the push [Fig. 15.2(a)] This friction force arises between two surfaces in contact. In this case, between the bottom of the box and the floor. It balances the pushing force and therefore the box does not move. In [Fig.15.2(b)] two children push the box harder but the box still does not move. This is because the frictional force still balances the pushing force. If the children push the box harder still, the pushing force becomes bigger than the frictional force [Fig.15.2.

(c)]. There is an imbalanced force. So, the box starts moving.

15.2 First law of motion

Galileo observed the motion of objects on an inclined plane. He deduced that objects move with a constant speed when no force acts on them.



Name	: Galileo
Born	: 15 February 1564
Birth place	: Grand Duchy of Tuscany, Italy
Died	: 8 January 1642
Best known for	: Astronomy, physics and mathematics

Newton studied Galileo's ideas on force and motion and presented three fundamental laws that govern the motion of objects. These three laws are known as Newton's Laws of Motion. The first law of motion is stated as:

An object remains in the state of rest or of uniform motion in a straight line unless compelled to change that state by an applied unbalanced force.

In other words, all objects resist a change in their state of motion. The tendency of undisturbed objects to stay at rest or to keep moving with the same velocity is called inertia. This is why, the first law of motion is also known as the law of inertia.

Certain experiences that we come across while travelling in a motor car can be explained on the basis of the law of inertia. We tend to remain at rest with respect to the seat until the driver applies a braking force to stop the motor car. With the application of brakes, the car slows down but our body tends to continue in the same state of motion because of inertia. A sudden application of brakes may thus cause injury to us by collision with panels in front.

An opposite experience is encountered when we are standing in a bus which begins to move suddenly. Now we tend to fall backwards. This is because a sudden start of the bus brings motion to the bus as well as to our feet in contact with the floor of the bus. But the rest of our body opposes this motion because of its inertia.

When a motor car makes a sharp turn at a high speed, we tend to get thrown to one side. This can again be explained on the basis of the law of inertia. We tend to continue in our straight line motion. When an unbalanced force is applied by the engine to change the direction of motion of the motor car, we move to one side of the seat due to the inertia of our body.

Inertia of a body can be illustrated through the following activities.

ACTIVITY 15.1

Make a pile of similar carrom coins on a table as shown in Fig.15.3.

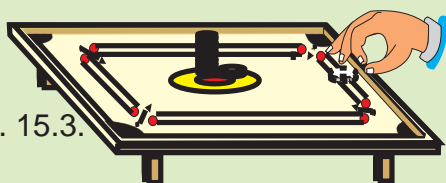


Fig. 15.3.

Attempt a sharp horizontal hit at the bottom of the pile using another carrom coin or the striker. If the hit is strong enough, the bottom coin moves out quickly. Once the lowest coin is removed, the inertia of the other coins makes them 'fall' vertically on the table.

15.3. INERTIA AND MASS

All the examples and activities given so far, illustrate that there is a resistance

offered by an object to change its state of motion. If it is at rest, it tends to remain at rest. If it is moving it tends to keep moving. This property of an object is called inertia. Therefore the ***inability of a body to change its state of rest or of uniform motion by itself is called inertia.***

Inertia of body depends mainly upon its mass. If we kick a foot ball, it flies away. But if we kick a stone of the same size with equal force, it hardly moves. We may, in fact get an injury in our foot. A force, that is just enough to cause a small carriage to pick up a large velocity, will produce a negligible change in the motion of a train. We say that train has more inertia than the carriage. Clearly, more massive objects offer larger inertia. The inertia of an object is measured by its mass.

15.4 MOMENTUM

Let us recount some observations from our everyday life. During the game of table tennis, if a ball hits a player, it does not hurt him. On the other hand, when fast moving cricket ball hits a spectator, it may hurt him. A truck at rest does not require any attention when parked along a roadside. But a moving truck, even at a very low speed, may kill a person standing in its path. A small mass such as a bullet may kill a person when fired from a gun. These observations suggest that the impact produced by an object depends on its mass and velocity. In other words, there appears to exist some quantity of importance that combines the object's mass and velocity. One such property called momentum was introduced by Newton. ***The momentum 'p' of an object is defined as the product of its mass 'm' and velocity 'v'. That is, $p=mv$***

Momentum has both direction and magnitude. It is a vector quantity. Its direction is same as that of the velocity. The SI unit of momentum is **kg ms⁻¹**.

15.5 SECOND LAW OF MOTION

Let us consider a situation in which a car with a dead battery is to be pushed along a straight road to give it a speed of 1 m s⁻¹ which is sufficient to start its engine. If one or two persons give a sudden push (unbalanced force) to it, it hardly starts. But a continuous push over it sometime results in a gradual acceleration of the car to the required speed. It means that the change of momentum of the car is not only determined by the magnitude of the force, but also by the time during which the force is exerted. It may then also be concluded that the force necessary to change the momentum of the object depends on the time rate at which the momentum is changed.

The second law of motion states that the **rate of change of momentum of an object is proportional to the applied unbalanced force in the direction of force**. Suppose an object of mass 'm' is moving along a straight line with an initial velocity 'u'. It is uniformly accelerated to velocity 'v' in time 't' by the application of constant force, 'F' throughout the time, 't'.

Initial momentum of the object = mu

Final momentum of the object = mv

The change in momentum = mv - mu = m(v - u) (1)

$$\begin{aligned} \text{Rate of change of momentum} &= \frac{\text{Change of momentum}}{\text{time}} \\ &= \frac{m(v-u)}{t} \quad (2) \end{aligned}$$

According to Newton II law of motion, this is nothing but applied force.

$$\text{Therefore the applied force, } F \propto \frac{m(v-u)}{t}$$

$$\text{But the acceleration, } a = \frac{v-u}{t}$$

(which is the rate of change of velocity).

The applied force, $F \propto ma$

$$F = Kma \quad (3)$$

'K' is known as the constant of proportionality. The SI unit of mass and acceleration are kg and m s⁻² respectively. The unit of force is so chosen that the value of the constant 'K' becomes one.

Therefore, $F = ma$ (4)

1 unit of force = (1 kg) x (1 m s⁻²)

The unit of force is kg m s⁻² or **newton** which has the symbol '**N**'.

One unit of force(1N) is defined as the amount of force that produces an acceleration of 1 m s⁻² in an object of 1 kg mass.

The second law of motion gives us a method to measure the force acting on an object as a product of its mass and acceleration.

Example:15.1

A constant force acts on an object of mass 10 kg for a duration of 4 s. It increases the objects velocity from 2 m s^{-1} to 8 m s^{-1} . Find the magnitude of the applied force.

Solution:

Given, mass of the object $m = 10 \text{ kg}$

Initial velocity $u = 2 \text{ m s}^{-1}$

Final velocity $v = 8 \text{ m s}^{-1}$

We know, force $F = \frac{m(v - u)}{t}$

$$F = \frac{10(8-2)}{4} = \frac{10 \times 6}{4} = 15 \text{ N}$$

Example:15.2

Which would require a greater force for accelerating a 2 kg of mass at 4 m s^{-2} or a 3 kg mass at 2 m s^{-2} ?

Solution

We know, force $F = ma$

Given $m_1 = 2 \text{ kg}$ $a_1 = 4 \text{ m s}^{-2}$

$m_2 = 3 \text{ kg}$ $a_2 = 2 \text{ m s}^{-2}$

Thus, $F_1 = m_1 a_1 = 2 \text{ kg} \times 4 \text{ m s}^{-2} = 8 \text{ N}$

and $F_2 = m_2 a_2 = 3 \text{ kg} \times 2 \text{ m s}^{-2} = 6 \text{ N}$

$$\Rightarrow F_1 > F_2$$

Thus, accelerating a 2 kg mass at 4 m s^{-2} would require a greater force.

15.6 THIRD LAW OF MOTION

Let us consider two spring balances connected together as shown in Fig. 15.4

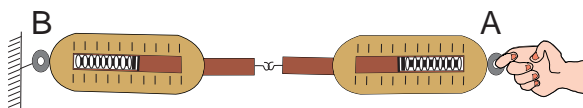


Fig. 15.4

The fixed end B of the balance is attached with a rigid support like a wall. When a force is applied through the free end of the spring balance A, it is observed that both the spring balances show the same readings on their scales. It means that the force exerted by spring balance A on balance B is equal but opposite in direction to the force exerted by the balance B on balance A. The force which balance A exerts on balance B is called action and the force of balance B on balance A is called the reaction.

Newton's third law of motion states that **for every action there is an equal and opposite reaction**. It must be remembered that the action and reaction always act on two different objects.

When a gun is fired it exerts forward force on the bullet. The bullet exerts an equal and opposite reaction force on the gun. This results in the recoil of the gun. Fig. 15.5

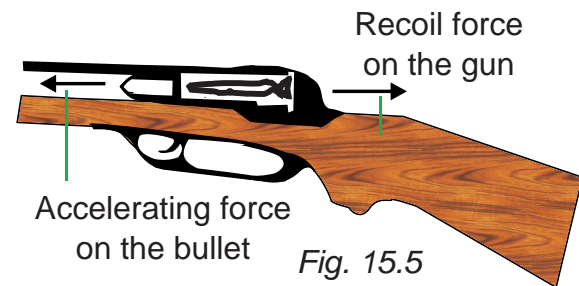


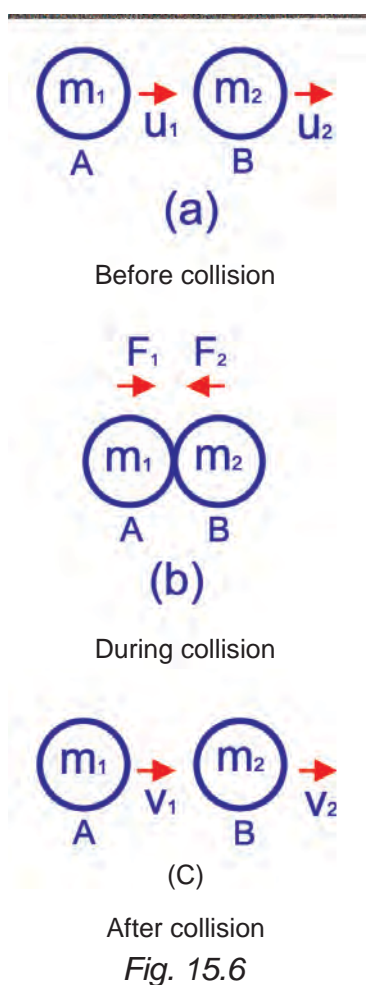
Fig. 15.5

Since the gun has a much greater mass than the bullet, the acceleration of the gun is much less than the acceleration of the bullet.

15.7 CONSERVATION OF MOMENTUM AND PROOF

The law of conservation of momentum states that, in the absence of external unbalanced force the total momentum of a system of objects remains unchanged or conserved by collision.

Consider two objects (two balls) A and B of masses ' m_1 ' and ' m_2 ' are traveling in the same direction along a straight line at different velocities ' u_1 ' and ' u_2 ' respectively Fig.15.6(a). There are no other external unbalanced forces acting on them. Let $u_1 > u_2$ and the two balls collide with each other as shown in Fig. 15.6(b). During collision which last for time ' t ', the ball A exerts a force F_1 on ball B, and the ball B exerts a force F_2 on ball A. Let v_1 and v_2 be the velocities of two balls A and B after collision respectively in the same direction as before collision, Fig 15.6(c).



According to Newton second law of motion

The force acting on B (action) $F_1 = \text{mass of B} \times \text{acceleration on B.}$

$$F_1 = \frac{m_2 (v_2 - u_2)}{t} \quad (1)$$

The force acting on A (reaction) $F_2 = \text{mass of A} \times \text{acceleration on A.}$

$$F_2 = \frac{m_1 (v_1 - u_1)}{t} \quad (2)$$

According to Newton's third law of motion

$$F_1 = -F_2$$

From equation (1) and (2)

$$\frac{m_2 (v_2 - u_2)}{t} = \frac{-m_1 (v_1 - u_1)}{t}$$

$$m_2 (v_2 - u_2) = -m_1 (v_1 - u_1)$$

$$m_2 v_2 - m_2 u_2 = -m_1 v_1 + m_1 u_1$$

$$m_1 v_1 + m_2 v_2 = m_1 u_1 + m_2 u_2$$

Therefore,

$$m_1 u_1 + m_2 u_2 = m_1 v_1 + m_2 v_2$$

The **total momentum before collision is equal to the total momentum after collision.** The total momentum of two objects remain unchanged due to collision in the absence of external force. This law holds good for any number of objects.

ACTIVITY 15.2

Take a big rubber balloon and inflate it fully. Tie its neck using a thread. Also using adhesive tape, fix a straw on the surface of this balloon.

- Pass a thread through the straw and hold one end of the thread in your hand or fix it on the wall.

ACTIVITY 15.2

- Ask your friend to hold the other end of the thread or fix it on a wall at some distance. This arrangement is shown in Fig.15.7
- Now remove the thread tied on the neck of the balloon. Let the air escape from the mouth of the balloon.
- Observe the direction in which the straw moves.

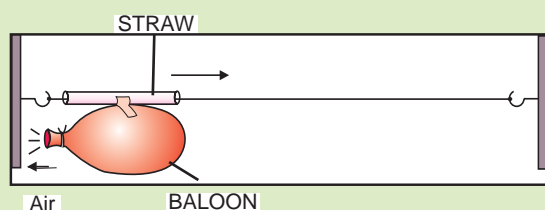


Fig. 15.7

Example:15.3

A bullet of mass 15g is horizontally fired with a velocity 100 m s^{-1} from a pistol of mass 2 kg what is the recoil velocity of the pistol?

Solution:

The mass of bullet, $m_1 = 15 \text{ g} = 0.015 \text{ kg}$

Mass of the pistol, $m_2 = 2 \text{ kg}$

Initial velocity of the bullet, $u_1 = 0$

Initial velocity of the pistol, $u_2 = 0$

Final velocity of the bullet, $v_1 = + 100 \text{ m s}^{-1}$

(The direction of bullet is taken from left to right-positive, by convention)

Recoil velocity of the pistol, = v

Total momentum of the pistol and bullet

before fire,

$$= (0.015 \times 0 + 2 \times 0) \text{ kg m s}^{-1}$$

$$= 0 \text{ kg m s}^{-1}$$

Total momentum of the pistol and bullet after fire,

$$= (0.015 \times 100 + 2 \times v)$$

$$= (1.5 + 2v) \text{ kg m s}^{-1}$$

According to the law of conservation of momentum,

Total momentum after fire = total momentum before fire

$$1.5 + 2v = 0$$

$$2v = -1.5$$

$$v = -0.75 \text{ m s}^{-1}$$

Negative sign indicates that the direction in which the pistol would recoil is opposite to that of the bullet, that is, right to left.

15.8 MOMENT OF FORCE AND COUPLE

Moment of a force

A force can rotate a nut when applied by a wrench or it can open a door while the door rotates on its hinges. In addition to the tendency to move a body in the direction of the application of a force, a force also tends to rotate the body about any axis which does not intersect the line of action of the force and also not parallel to it. This tendency of rotation is called turning effect of a force or moment of the force about the given axis. **The magnitude of the moment of force F about a point is defined as the product of the magnitude of force and the perpendicular distance of the point from the line of action of the force.**

Let us consider a force F acting at the point P on the body as shown in Fig. 15.8

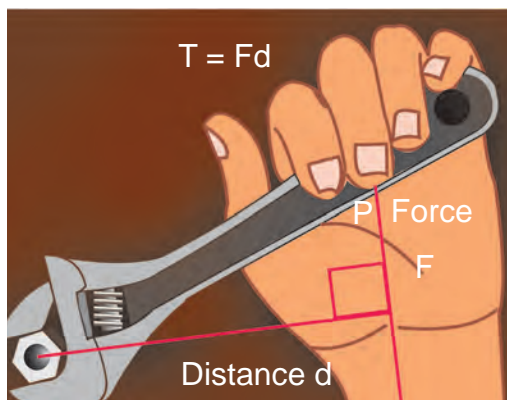


Fig. 15.8

Then, **the moment of the force F about the point $O = \text{Magnitude of the force } \times \text{ perpendicular distance}$** between the direction of the force and the point about which moment is to be determined $= F \times d$.

If the force acting on a body rotates the body in anticlockwise direction with respect to O then the moment is called anticlockwise moment. On the other hand, if the force rotates the body in clockwise direction then the moment is said to be clockwise moment. The unit of moment of the force is **N m**.

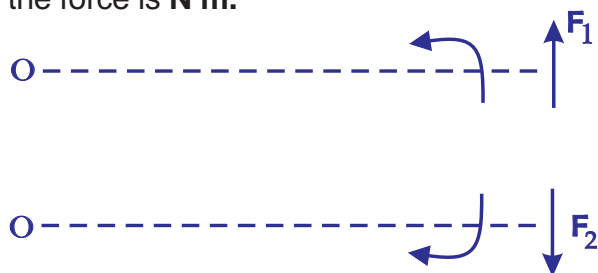


Fig. 15.9.

As a matter of convention, an anticlockwise moment is taken as positive and a clockwise moment as negative.

Couple

There are many examples in practice where two forces, acting together, exert a moment or turning effect on some object. As a very simple case, suppose two strings are tied to a wheel at the points X

and Y , and two equal and opposite forces, ' F ' are exerted tangentially to the wheels (Fig. 15.10). If the wheel is pivoted at its centre O it begins to rotate about O in an anticlockwise direction.

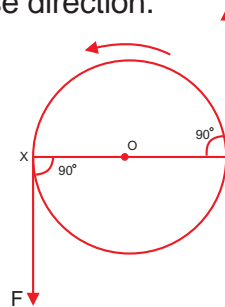


Fig. 15.10

Two equal and opposite forces whose lines of action do not coincide are said to **constitute a couple in mechanics**.

15.9. GRAVITATION



Name	: Isaac Newton
Born	: 4 January 1643
Birth Place	: Woolsthorpe, England
Died	: 20 March 1727
Best Known as :	The genius who explained gravity.

We always observe that an object dropped from a height falls towards the earth. It is said that Newton was sitting under the tree, an apple fell on him. The fall of the apple made Newton start thinking. It is seen that a falling apple is attracted towards the earth. Does the apple attract the earth? If so we do not see earth moving towards an apple. Why?

According to Newton's Third Law of Motion, the apple does attract the earth. But according to Second Law of motion, for a given force, acceleration is inversely proportional to the mass of the object. The mass of an apple is negligibly small compared to that of the earth. So we do not see the earth moving towards the apple. We know that all planets go around the sun. Extend the above argument for all

planets in our solar system. There exist a force between sun and the planets. **Newton concluded that all objects in the universe attract each other. This force of attraction between objects is called the gravitational force.**

ACTIVITY 15.3

Take a piece of thread. Tie a small stone at one end.

Hold the other end of the thread and whirl it round as shown in Fig. 15.11.

Note the motion of the stone.

Release the thread.

Again note the direction of motion of the stone.



Fig 15.11.

It is noted that the stone describes a circular path with a velocity of constant magnitude.

15.9.1. Newton law of gravitation

Every object in the universe attracts every other object with a force which is directly proportional to the product of their masses and inversely proportional to the square of the distance between them. The force acts along the line joining the centers of two objects.

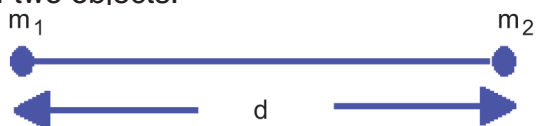


Fig. 15.12

Let two objects A and B of masses m_1 , m_2 respectively lie at a distance 'd' from each other as shown in Fig.15.12. Let the force of attraction between two objects is 'F'. According to above law,

$$F \propto m_1 m_2 \quad (1)$$

$$F \propto \frac{1}{d^2} \quad (2)$$

Combining (1) and (2)

$$F \propto \frac{m_1 m_2}{d^2} \quad (3)$$

or
$$F = \frac{G m_1 m_2}{d^2} \quad (4)$$

Where G is the constant of proportionality and is called the Universal gravitation constant. From eqn (4)

$$G = \frac{F \cdot d^2}{m_1 m_2}$$

Substituting the S.I units in this equation the unit of G is found to be $\text{N m}^2\text{kg}^{-2}$
The value of G is $6.673 \times 10^{-11} \text{ N m}^2\text{kg}^{-2}$

15.9.2 Mass

Mass is the amount of matter present in a body (or) is a measure of how much matter an object has.

15.9.3 Weight

Weight is the force which a given mass feels due to the gravity at its place (or) is a measure of how strongly gravity pulls on that matter.

If you were to travel to the moon, your weight would change because the pull of the gravity is weaker there than on the earth, but your mass would stay the same

because you are still made up of the same amount of matter.

Example 15.4

Mass of an object is 5 kg. What is its weight on the earth?

Solution:

Mass, $m = 5 \text{ kg}$

Acceleration due to gravity, $g = 9.8 \text{ m s}^{-2}$

Weight, $w = m \times g$

$$w = 5 \text{ kg} \times 9.8 \text{ m s}^{-2} = 49 \text{ N}$$

Thus the weight of the object is, **49 N**

Difference between mass and weight

Mass	Weight
1. Fundamental quantity.	Derived quantity.
2. It is the amount of matter contained in a body.	It is the gravitational pull acting on the body.
3. Its unit is kilogram.	It is measured in newton.
4. Remains the same.	Varies from place to place.
5. It is measured using physical balance.	It is measured using spring balance.

15.9.4 Acceleration due to gravity

Galileo was the first to make a systematic study of the motion of a body under the gravity of the Earth. He dropped various objects from leaning tower of Pisa and made analysis of their motion under gravity. He came to the conclusion that “***in the absence of air, all bodies will fall at the same rate***”.

It is the air resistance that slows down a piece of paper or a parachute falling under gravity. If a heavy stone and a parachute are dropped where there is no air, both will fall together at the same rate.

Experiments showed that the velocity of a freely falling body under gravity increases at a constant rate. (i.e.) with a constant acceleration. **The acceleration produced in a body on account of the force of gravity is called acceleration due to gravity.** It is denoted by ***g***. At a given place, the value of ***g*** is the same for all bodies irrespective of their masses. It differs from place to place on the surface of the Earth. It also varies with altitude and depth.

The value of ***g*** at sea-level and at a latitude of 45° is taken as the standard free-fall acceleration (i.e.) **$g=9.8 \text{ m s}^{-2}$**

Acceleration due to gravity at the surface of the earth

Consider a body of mass ‘*m*’ on the surface of the earth as shown in Fig. 15.13.

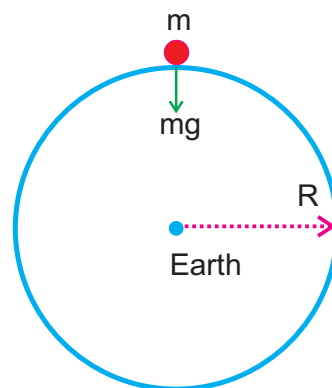


Fig.15.13

Its distance from the centre of the Earth is *R* (radius of the Earth).

The gravitational force experienced by the body is $F = \frac{GMm}{R^2}$ where M is the mass of the earth. From Newton's second law of motion,

$$\text{Force, } F = mg$$

Equating the above two forces,

$$F = \frac{GMm}{R^2} = mg$$

Therefore,

$$g = \frac{GM}{R^2}$$

This equation shows that 'g' is independent of the mass of the body 'm' but, it varies with the distance from the centre of the Earth. If the Earth is assumed to be a sphere of radius R, the value of 'g' on the surface of the Earth is given by

$$g = \frac{GM}{R^2}$$

15.9.5. Mass of earth

From the expression $g = GM/R^2$, the mass of the Earth can be calculated as follows:

$$M = \frac{gR^2}{G}$$

$$M = 9.8 \times (6.38 \times 10^6)^2 / 6.67 \times 10^{-11}$$

$$M = 5.98 \times 10^{24} \text{ kg.}$$

Science today

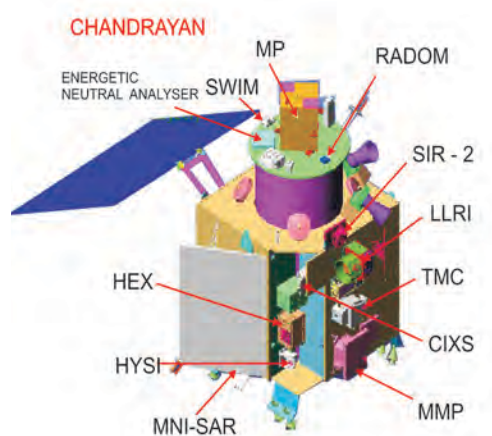
Chandrayaan

Chandrayaan-1 is a moon-traveler or moon vehicle. It was Indian's first unmanned lunar probe. It was launched by Indian Space Research Organization in October 2008 from Srihari Kota in Andrapradesh and operated until August



Mylsamy Annadurai was born on 2nd July 1958 at Kodhavadi, a hamlet near Pollachi in Coimbatore District. Mylsamy and Balasaraswathy are his parents. His father served as a teacher in an Elementary school. Panchayat Union Elementary School in Kothavadi was Mylsamy Annadurai's first school, where he studied from I to V stds. He then moved to Government and Aided schools in and around his native place for continuing and completing his school education upto XI std. His educational journey continued. He finished his PUC in NGM College, Pollachi and B.E degree at Government College of Technology, Coimbatore. In 1982 he pursued his Higher Education and acquired the M.E degree in PSG College of Technology, Coimbatore and the same year he joined in ISRO as a scientist. And later he got Doctorate in Anna University of Technology, Coimbatore..

Annadurai is a leading technologist in the field of satellite system. Currently Annadurai serves as the Project Director of Chandrayaan-1 and Chandrayaan-2. He has made significant contribution to the cost effective design of Chandrayaan. Through his inspiring speeches he has become a motivating force among the Indian students.



2009. The mission included a lunar orbiter and an impactor. It carried five ISRO payloads and six payloads from other space agencies including NASA, European Space Agencies (ESA), and the Bulgarian Aerospace Agency which were carried free of cost.

Chandrayaan operated for 312 days and achieved 95% of its planned objectives. The following are its achievements,

- The discovery of wide spread presence of water molecules in lunar soil.
- Chandrayaan's Moon Mineralogy Mapper has confirmed that moon was once completely molten.
- European Space Agency payload- Chandrayaan-1 imaging X-ray spectrometer (CIXS) detected more than two dozen weak solar flares during the mission.
- The terrain mapping camera on board Chandrayaan-1 has recorded images of the landing site of US space craft Apollo-15, Apollo-11.
- It has provided high-resolution spectral data on the mineralogy of the moon.

- Lunar Laser Ranging Instrument (LLRI) covered both the Lunar Poles and additional lunar region of interest.
- The X-ray signatures of aluminum, magnesium and silicon were picked up by the CIXS X-ray camera
- The Bulgarian payload called Radiation Dose Monitor (RADOM) was activated on the day of launch itself and worked till the mission end.
- More than 40000 images have been transmitted by Chandrayaan Camera in 75 days.
- The Terrain Mapping Camera acquired images of peaks and Craters. The moon consists of mostly of Craters.
- Chandrayaan beamed back its first images of the Earth in its entirety.
- Chandrayaan-1 has discovered large caves on the lunar surface that can act as human shelter on the moon.

Cryogenic techniques

The word cryogenics terms from Greek and means "the production of freezing cold".

In physics cryogenics is the study of the production of very low temperature (below 123K); and the behaviour of materials at those temperature. A person who studies elements under extremely cold temperature is called a cryogenicist. Cryogenics use the Kelvin scale of temperature. Liquefied gases such as liquid nitrogen, liquid helium is used in many cryogenic applications. Liquid nitrogen is the most commonly used element in cryogenics and is legally purchasable around the world. Liquid helium is also commonly used and allows

for the lowest attainable temperature to be reached. These liquids are held in special containers called Dewar flasks which are generally about six feet tall and three feet in diameter.

The field of cryogenics advanced during world war-2. Scientist found that metals frozen to low temperature showed more resistance to wear. This is known as cryogenic hardening. The commercial cryogenic processing industry was founded in 1966 by Ed Busch; and merged several small companies later to form oldest commercial cryogenic company in the world. They originally experimented with the possibility of increasing the life of metal tools.

Cryogens like liquid nitrogen are further used for specially chilling and freezing applications.

(i) Rocket

The important use of cryogenics is cryogenic fuels. Cryogenic fuels mainly liquid hydrogen has been used as rocket fuel.

(ii) Magnetic Resonance Imaging (MRI)

MRI is used to scan inner organs of human body by penetrating very intense magnetic field. The magnetic field is generated by super conducting coils with the help of liquid helium. It can reduce the temperature of the coil to around 4k. At this low temperature very high resolution images can be obtained.

(iii) Power transmission in big cities:

It is difficult to transmit power by overhead cables in cities. So underground cables are used. But underground cables get heated and the resistance of the wire increases leading to wastage of power.

This can be solved by cryogenics. Liquefied gases are sprayed on the cables to keep them cool and reduce their resistance.

(iv) Food Freezing:

Cryogenic gases are used in transportation of large masses of frozen food, when very large quantity of food must be transported to regions like war field, earthquake hit regions etc., they must be stored for.

(v) Vaccines:

The freezing of biotechnology products like vaccines require nitrogen freezing systems.

Space station:

A space station is an artificial structure designed for humans to live and work in outer space for a period of time.

Current and recent-history space stations are designed for medium-term living in orbit, for periods of weeks, months or even years. The only space stations are Almaz and Salyut series, Sky lab and Mir.



Space stations are used to study the effects of long-space flight on the human body. It provides platforms for greater number and length of scientific studies than available on other space vehicles. Space stations have been used for both military and civilian purposes. The last

military-used space station was Salyut 5, which was used by the Almaz program of the Soviet Union in 1976 and 1977.

Broadly speaking the space stations so far launched has been of two types. Salyut and Skylab have been “monolithic.” They were constructed and launched in one piece, and then manned by a crew later. As such, they generally contained all their supplies and experimental equipment when launched, and were considered “expended”, and then abandoned, when these were used up.

Starting with Salyut 6 and Salyut 7, a change was seen. These were built with two docking ports. They allowed a second crew to visit, bringing a new space craft with them.

This allowed for a crew to man the station continually, sky lab was also equipped with two docking ports, but the extra port was never utilized. The presence of the second port on the new station allowed progress supply vehicle to be docked to the station, meaning that fresh supplies could be brought to aid long-duration missions.

The second group, Mir and the International Space Station (ISS), have been modular; a core unit was launched, and additional modules, generally with a specific role, were later added to that.

(on Mir they were usually launched independently, whereas on the ISS most are brought by the Space Shuttle). This method allows for greater flexibility in operation. It removes the need for a single immensely powerful launch vehicle. These stations are also designed from the outset to have their supplies provided by logistical support, which allows for a longer lifetime at the cost of requiring regular support launches.

These stations have various issues that limit their long-term habitability, such as very low recycling rates, relatively high radiation levels and a lack of gravity. Some of these problems cause discomfort and long-term health effects.

Future space habitats may attempt to address these issues, and are intended for long-term occupation. Some designs might even accommodate large number of people, essentially “cities in space” where people would make their homes. No such design has yet been constructed, even for a small station; the current (2010) launch costs are not economically or politically viable.

The People’s Republic of China is expected to launch its space station named Tiangong 1, in the first half of 2011. This would make China the third country to launch a space station.

EVALUATION

PART A

- The acceleration in a body is due to _____.
(balanced force, un-balanced force, electro static force)
- The physical quantity which is equal to rate of change of momentum is (displacement, acceleration, force, impulse)
- The momentum of a massive object at rest is _____.
(very large, very small, zero, infinity)

4. The weight of 50 kg person at the surface of earth is _____.
(50 N, 35 N, 380 N, 490 N)
5. The freezing of biotechnology products like vaccines require _____ freezing systems.
(Helium, Nitrogen, Ammonia, Chlorine)
6. Action and reaction is always acting on the same body.
5. The important use of cryogenics is cryogenic fuels. What do you mean by cryogenic fuels?
6. As a matter of convention, an anticlockwise moment is taken as _____ and a clockwise moment is taken as _____.

PART – B

1. From the following statements write down that which is not applicable to mass of an object
 - a. It is a fundamental quantity
 - b. It is measured using physical balance.
 - c. It is measured using spring balance.
2. Fill in the blanks.
 - a) Force = mass x acceleration, then momentum = _____?
 - b) Liquid hydrogen is for rocket, then _____ for MRI.
3. The name of some organisations which are associated with Chandrayan-I mission are given below. but some of them are not. List out the wrong ones.
(ISRO, BARC, NASA, ESA, WHO, ONGC)
4. Correct the mistakes, if any, in the following statements.
 - a. One newton is the force that produces an acceleration of 1 ms^{-2} in an object of 1 gram mass.

PART – C

1. a) Newton’s first law of motion gives a qualitative definition of force. Justify.



- b) The figure represents two bodies of masses 10 kg and 20 kg and moving with an initial velocity of 10 ms^{-1} and 5 ms^{-1} respectively. They are colliding with each other. After collision they are moving with velocities 12 ms^{-1} and 4 ms^{-1} respectively. The time of collision be .2 s. Then calculate F_1 and F_2 .
2. a) Space stations are used to study the effects of long-space flight on the human body. justify.
- b) $F = G \frac{m_1 m_2}{d^2}$ is the mathematical form of Newton’s law of gravitation, G - gravitational constant, m_1 , m_2 , are the masses of two bodies separated by a distance d, then give the statement of Newton’s law of gravitation.

FURTHER REFERENCE

Books : 1. Advanced physics by : **M. Nelkon and P. Parker**, C.B.S publications
2. College Physics by : **R.L.Weber, K.V. Manning**, Tata McGraw Hill

Websites: www.britannica.com | www.zonalandeducation.com | www.wiki.animers.com
<http://www.khanacademy.org>



ELECTRICITY AND ENERGY

16 ELECTRICITY AND ENERGY

Name	: Michael Faraday
Born	: 22 September 1791
Birth place	: Newington, England
Died	: 25 August 1867
Best known as	: Inventor of the first dynamo



Electricity has an important place in modern society. It is a controllable and convenient form of energy for variety of uses in homes, schools, hospitals, industries and so on. What constitutes electricity?

How does it flow in an electric circuit? What are the factors that regulate electricity through an electric circuit?. In this chapter we shall attempt to answer such questions.

16.1. ELECTRIC CURRENT AND CIRCUIT

We are familiar with air current and water current. We know that flowing water constitute water current in rivers. Similarly if the electric charge flows through a conductor (metallic wire), we say that there is an electric current in the conductor. In a

torch we know that a battery provide flow of charges or an electric current through a torch bulb to glow. We have also seen that it gives light only when it is switched on. What does a switch do? A switch makes a conducting link between the cell and the bulb. ***A continuous and closed path of an electric current is called an electric circuit.*** Now if the circuit is broken anywhere the current stops flowing and the bulb does not glow.

How do we express electric current? ***Electric current is expressed by the amount of charge flowing through a particular area of cross section of a conductor in unit time.*** In other words it is the rate of flow of electric charges. In circuit using metallic wires, electrons constitute flow of charges. The direction of electric current is taken as opposite to the direction of the flow of electrons.

If a net charge Q , flows across any cross-section of a conductor in time t , then the current I through the cross-section is

$$I=Q/t$$

The S.I unit of electric charge is **coulomb**. This is equivalent to the charge contained in nearly 6×10^{18} electrons. The electric current is expressed by a unit called **ampere (A)**, named after the French Scientist.

From the above equation,

When $Q = 1 \text{ C}$, $t = 1\text{s}$, $I=1\text{A}$.

When one coulomb of charge flows in one second across any cross section of a conductor, the current in it is one ampere. An instrument called ammeter is used to measure current in a circuit.

Example 16.1

A current of 0.75 A is drawn by a filament of an electric bulb for 10 minutes. Find the amount of electric charge that flows through the circuit.

Solution:

Given, $I = 0.75 \text{ A}$,
 $t = 10 \text{ minutes} = 600 \text{ s}$

We know, $Q = I \times t$
 $= 0.75 \text{ A} \times 600 \text{ s}$
 $Q = 450 \text{ C}$

The Fig.16.1 shows a schematic diagram of an electric circuit comprising battery, bulb, ammeter and a plug key.

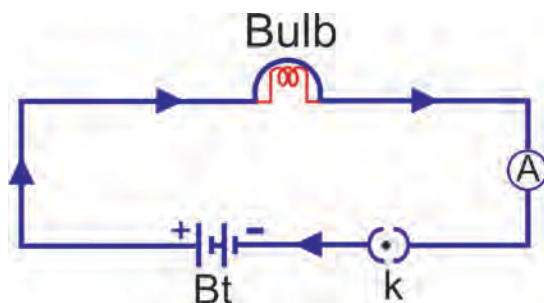


Fig. 16.1 Electric circuit

16.2. ELECTRIC POTENTIAL AND POTENTIAL DIFFERENCE

What makes the electric charge to flow? Charges do not flow in a copper wire by themselves, just as water in a perfectly horizontal tube does not flow. One end of the tube is connected to a tank of water. Now there is a pressure difference between the two ends of the tube. Water flows out of the other end of the tube. For flow of charges in a conducting metallic wire, the electrons move only if there is a difference of electric pressure-called potential difference-along the conductor. This difference of potential may be produced by a battery, consisting of one or more electric cells. When the cell is connected to a conducting circuit element, the potential difference sets the charges in motion in the conductor and produces an electric current.

We define the electric potential difference between two points in an electric circuit carrying current as the work done to move a unit charge from one point to the other.

Potential difference (V) between two points = work done (W)/charge (Q).

$$V = W/Q$$

The S.I Unit of potential difference is volt (V).

$$1 \text{ volt} = 1\text{joule}/1\text{coulomb}$$

One volt is the potential difference between two points in a current carrying conductor when 1 joule of work is done to move a charge of 1 coulomb from one point to the other.

The potential difference is measured by means of an instrument called voltmeter.

16.3. CIRCUIT DIAGRAM

The Schematic diagram, in which different components of the circuit are represented by the symbols conveniently used, is called a circuit diagram. Conventional symbols used to represent some of the most commonly used electrical components are given in table 16.1.

COMPONENTS	SYMBOLS
An electric cell	
A battery or a combination of cells	
Plug key or switch (open)	
Plug key or switch (closed)	
A wire joint	
Wires crossing without joining	
Electric bulb	
A resistor of resistance R	
Variable resistance or rheostat	
Ammeter	
Voltmeter	
Light Emitting Diode	

Table 16.1.

Example 16.2.

How much work is done in moving a charge of 5 C across two points having a potential difference 10 V ?

Solution:

Given charge, $Q = 5 \text{ C}$

Potential difference, $V = 10 \text{ V}$

The amount of work done in moving the charge, $W = V \times Q$
 $W = 10 \text{ V} \times 5 \text{ C} = 50 \text{ J}$

16.4. OHM'S LAW

Name	: George Simon Ohm
Born	: 16 March 1789
Birth place	: Erlangen, Germany
Died	: 06 July 1854
Best known for	: Ohm's law



Is there a relationship between the potential difference across a conductor and the current through it? Let us explore with an activity.

In this activity you will find the ratio V/I is a constant.

ACTIVITY 16.1

- Set up a circuit as shown in Fig. 16.2. consisting of a nichrome wire XY of length say 0.5m, an ammeter, a Voltmeter and four cells of 1.5V each. (Nichrome is an alloy of Nickel, Chromium, Manganese and Iron metals).
- First use only one cell as the source in the circuit. Note the reading in the ammeter I , for the current and reading of the voltmeter V for the potential difference across the nichrome wire XY in the circuit. Tabulate them in the table given.

ACTIVITY

Repeat the above steps using two, three cells and then four cells in the circuit separately.

- Calculate the ratio of V to I for each pair of potential difference V and current I .

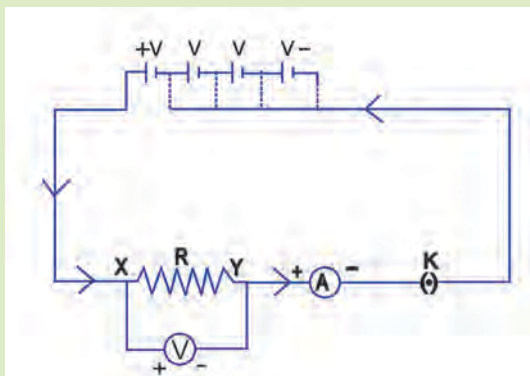


Fig. 16.2

In 1827, a German Physicist George Simon Ohm found out the relationship between the current I flowing in a metallic wire and the potential difference across its terminals.

Ohm's law states that at constant temperature the steady current (I) flowing through a conductor is directly proportional to the potential difference (V) between its ends.

$$V \propto I \text{ (or) } V/I = \text{constant.}$$

Example 16.3

The potential difference between the terminals of an electric heater is 60 V when it draws a current of 5 A from the source. What current will the heater draw if the potential difference is increased to 120 V ?

Solution:

Given the potential difference, $V = 60 \text{ V}$

Current, $I = 5 \text{ A}$

According to ohm's law,

$$R = V/I = 60 \text{ V} / 5 \text{ A} = 12 \Omega$$

When the potential difference is increased to 120 V, the current is given by

$$I = V/R = 120 \text{ V} / 12 \Omega = 10 \text{ A}$$

16.5. RESISTANCE OF A CONDUCTOR

S.No	Number of cells used in the circuit	Current through the nichrome wire I (ampere)	Potential difference across the nichrome wire. V (volt)	V/I (volt/ampere) Ω
1.				
2.				
3.				
4.				
5.				
6.				

From Ohm's law, we know

$$V \propto I, \mathbf{V = IR}$$

R is a constant for a given metallic wire at a given temperature and is called its resistance. It is the property of a conductor to resist the flow of charges through it. Its S.I unit is ohm, represented by the Greek letter Ω .

$$R = V/I, 1 \text{ ohm} = 1 \text{ volt}/1 \text{ ampere}$$

If the potential difference across the two ends of a conductor is 1 volt and the current through it is 1 ampere, then the resistance of the conductor is 1 ohm.

16.6. SYSTEM OF RESISTORS

ACTIVITY 16.2

- Set up the circuit by connecting four dry cells of 1.5V each in series with the ammeter leaving a gap XY in the circuit, as shown in Fig. 16.3.
- Complete the circuit by connecting the nichrome wire in the gap XY. Plug the key. Note down the ammeter reading. Take out the key from the plug.
- Replace the nichrome wire with the torch bulb in the circuit and find the current through it by measuring the reading of the ammeter.

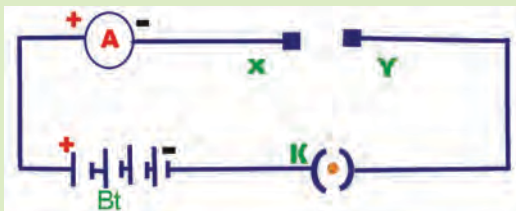


Fig. 16.3

ACTIVITY

- Now repeat the above steps with the LED bulb in the gap XY.
- Are the ammeter readings differ for different components connected in the gap XY?. What do the above observations indicate?

In various electrical circuits we often use resistors in various combinations. There are two methods of joining the resistors together. Resistors can be connected in series or in parallel.

Resistors in series

Consider three resistors of resistances R_1, R_2, R_3 in series with a battery and a plug key as shown in Fig. 16.4.

The current through each resistor is the same having a value I . The total

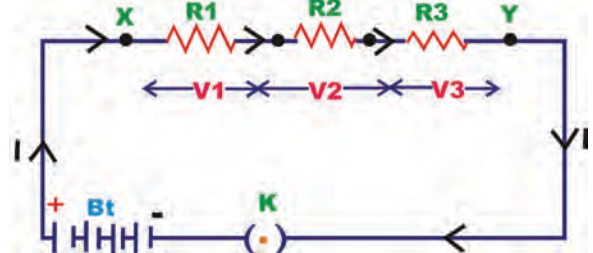


Fig. 16.4

potential difference across the combination of resistors in series is equal to the sum of potential difference across individual resistors. That is,

$$\mathbf{V = V_1 + V_2 + V_3} \quad (1)$$

It is possible to replace the three resistors joined in series by an equivalent single resistor of resistance R_s such that the potential difference V across it, and

the current I through the circuit remains the same.

Applying Ohm's law to the entire circuit we have, $V=IR$

On applying ohm's law to the three resistors

separately we further have

$$V_1 = IR_1, V_2 = IR_2 \text{ and } V_3 = IR_3$$

Substituting these values in equation (1)

$$IR = IR_1 + IR_2 + IR_3$$

$$\text{(or)} \quad R_s = R_1 + R_2 + R_3$$

When several resistors are connected in series, the resistance of the combination R_s is equal to the sum of their individual resistances R_1, R_2, R_3 and is thus greater than any individual resistance.

Example 16.4

Two resistances 18Ω and 6Ω are connected to a 6 V battery in series. Calculate (a) the total resistance of the circuit, (b) the current through the circuit.

Solution:

$$\begin{aligned} \text{(a) Given the resistance, } R_1 &= 18 \Omega \\ R_2 &= 6 \Omega \end{aligned}$$

The total resistance of the circuit $R_s = R_1 + R_2$

$$R_s = 18 \Omega + 6 \Omega = 24 \Omega$$

(b) The potential difference across the two terminals of the battery $V = 6 \text{ V}$

Now the current through the circuit,

$$\begin{aligned} I &= V / R_s = 6 \text{ V} / 24 \Omega \\ &= 0.25 \text{ A} \end{aligned}$$

Resistors in parallel

Consider three resistors having resistances R_1, R_2, R_3 connected in parallel. This combination is connected with a battery and plug key as shown in Fig. 16.5

In parallel combination the potential difference across each resistor is the same having a value V . The total current I is equal to the sum of the separate currents through each branch of the combination.

$$I = I_1 + I_2 + I_3 \quad (1)$$

Let R_p be the equivalent resistance of

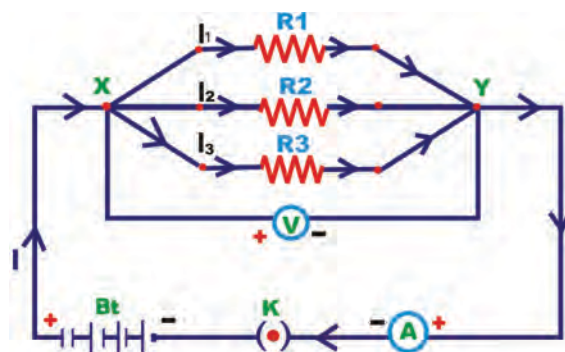


Fig. 16.5

the parallel combination of resistors. By applying Ohm's law to the parallel combination of resistors we have $I = V/R_p$

On applying ohm's law to each resistor We have

$$I_1 = V/R_1, I_2 = V/R_2 \text{ and } I_3 = V/R_3$$

Substituting these values in equation (1)

$$V/R_p = V/R_1 + V/R_2 + V/R_3$$

$$\text{(or)} \quad 1/R_p = 1/R_1 + 1/R_2 + 1/R_3$$

Thus the reciprocal of the equivalent resistance of a group of resistance joined in parallel is equal to the sum of the reciprocals of the individual resistance.

Example 16.5

Three resistances having the values 5Ω , 10Ω , 30Ω are connected parallel with each other. Calculate the total circuit resistance.

Solution:

Given, $R_1 = 5 \Omega$, $R_2 = 10 \Omega$, $R_3 = 30 \Omega$

These resistances are connected parallel

Therefore, $1/R_p = 1/R_1 + 1/R_2 + 1/R_3$

$$\frac{1}{R_p} = \frac{1}{5} + \frac{1}{10} + \frac{1}{30} = \frac{10}{30}$$

$$R_p = \frac{30}{10} = 3 \Omega$$

16.7. HEATING EFFECT OF ELECTRIC CURRENT

ACTIVITY 16.3

- Take an electric cell, a bulb, a switch and connecting wires. Make an electric circuit as shown in Fig. 16.6. By pressing the key allow the current to pass through the bulb.
- The bulb gets heated when current flows continuously for a long time (when the key is on).

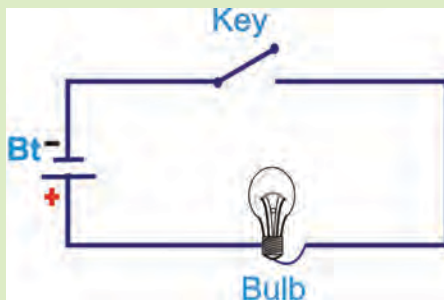


Fig. 16.6

We know that a battery is a source of electrical energy. Its potential difference between the two terminals sets the electrons in motion to flow the current through the resistor. To flow the current, the source has to keep spending its energy. Where does this energy go? What happens when an electric fan is used continuously for longer time? A part of the energy may be consumed into useful work (like in rotating the blades of the fan). Rest of the energy may be expended in heat to raise the temperature of the gadget. If the electric circuit is purely resistive, the source energy continuously gets dissipated entirely in the form of heat. This is known as heating effect of electric current. Heating effect of electric current has many useful appliances. The electric laundry iron, electric toaster, electric oven and electric heater are some of the familiar devices which uses this effect.

16.8. JOULES LAW OF HEATING

Consider a current I flowing through a resistor of resistance R . Let the potential difference across it be V . Let t be the time during which a charge Q flows across. The work done in moving the charge Q through the potential difference V is VQ . Therefore the source must supply energy equal to VQ in time t . Hence the power input to the circuit by the source is

$$P = V (Q/t) = VI$$

or the energy supplied through the circuit by the source in time t is $P \times t$, that is $VI \times t$. What happens to this energy

expended by the source? This energy gets dissipated in the resistor as heat. Thus for a steady current I , the amount of heat H produced in time t is

$$H = V I t$$

Applying Ohm's law we get $H = I^2 R t$.

This is known as Joule's law of heating. The law implies that heat produced in a resistor is (1) directly proportional to the square of current for a given resistance, (2) directly proportional to the resistance for a given current, and (3) directly proportional to the time for which the current flows through the resistor.

Example 16.6

A potential difference 20 V is applied across a 4Ω resistor. Find the rate of production of heat.

Solution:

Given potential difference, $V = 20 \text{ V}$

The resistance, $R = 4 \Omega$

The time, $t = 1 \text{ s}$

According to ohm's law, $I = V / R$

$$I = 20 \text{ V} / 4 \Omega = 5 \text{ A}$$

The rate of production of heat, $H = I^2 R t$

$$H = 5^2 \times 4 \times 1 \text{ J} = 100 \text{ J}$$

16.9. ROLE OF FUSE

A common application of Joule's heating is the fuse used in electric circuits. It consists of a piece of wire made of metal or an alloy (37% Lead, 63% Tin). It has high resistance and low melting point. The fuse is connected in series with the device. During the flow of any unduly high

electric current the fuse wire melts and protects the circuits and appliances.

16.10. DOMESTIC ELECTRIC CIRCUITS

In our homes, we receive supply of electric power through a main supply (also called mains), either supported through overhead electric poles or by underground cables. One of the wires in the supply, usually with red insulation cover, is called live wire (or positive). Another wire, with black insulation, is called neutral wire (or negative). In our country, the potential differences between the two are 220 V.

At the meter-board in the house, these wires pass into an electricity meter through a main fuse. Through the main switch they are connected to the line wires in the house. These wires supply electricity to separate circuits with in the house. Often, two separate circuits are used, one of 5A current rating for appliances with higher power ratings such as geysers, air coolers, etc. The other circuit is of 5A current rating for bulbs, fans, etc. The earth wire which has insulation of green colour is usually connected to a metal plate deep in the earth near the house. This is used as a safety measure, especially for those appliances that have a metallic body, for example, electric press, toaster, table fan, refrigerator, etc. The metallic body is connected to the earth wire, which provides a low-resistance conducting path for the current. Thus, it ensures that any leakage of current to the metallic body of the appliance keep its potential to that of the earth, and the user may not get a severe electric shock.

Fig.16.7 gives a schematic diagram of one of the common domestic circuits. In each separate circuit, different appliances can be connected across the live and neutral wires. Each appliance has a separate switch to ON/OFF the flow of current through it. In order that each appliance has equal potential difference, they are connected parallel to each other.

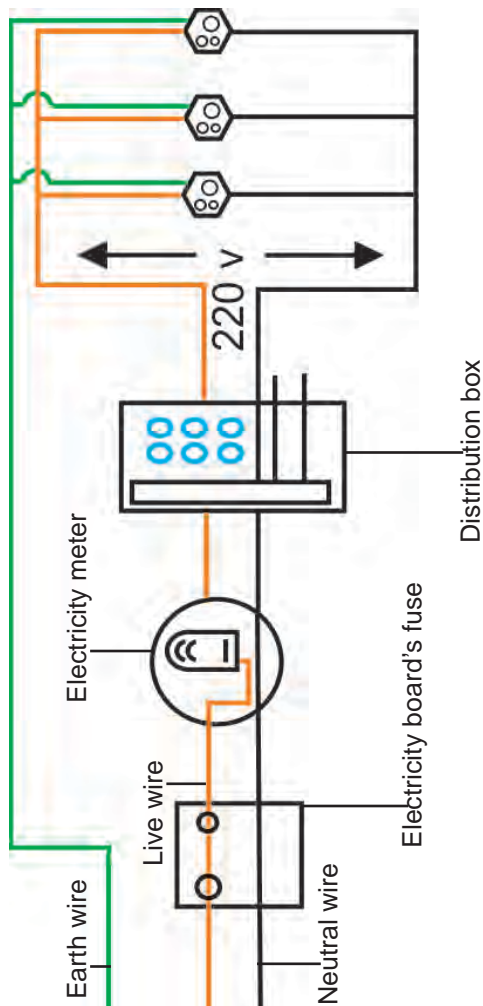


Fig. 16.7

Electric fuse is an important component of all domestic circuits. Over loading can occur when the live wire and the neutral wire come into direct contact. In such a situation the current in the circuit

abruptly increases. This is called short circuiting. The use of an electric fuse prevents the electric circuit and appliance from a possible damage by stopping the flow of unduly high electric current.

16.11. ELECTRIC POWER

We know already that the rate of doing work is power. This is also the rate of consumption of energy. This is also termed as electric power.

The power P is given by $P=VI$

$$\text{(or)} \quad P=I^2 R = V^2/R$$

The SI unit of electric power is watt (W). It is the power consumed by a device that carries 1 A of current when operated at a potential difference of 1 V. Thus,

$$1 \text{ W}=1 \text{ volt} \times 1 \text{ ampere} =1 \text{ V A}$$

The unit watt is very small. Therefore, in actual practice we use a much larger unit called kilowatt. It is equal to 1000 watt. Since electric energy is the product of power and time, the unit of electric energy is, therefore, watt hour (Wh). One watt hour is the energy consumed when one watt of power is used for one hour. The commercial unit of electric energy is kilowatt hour (KWh), commonly known as unit.

$$\begin{aligned} 1 \text{ kWh} &= 1000\text{watt} \times 3600\text{second} \\ &= 3.6 \times 10^6 \text{ watt second} \\ &= 3.6 \times 10^6 \text{ joule (J)} \end{aligned}$$

Example 16.7

An electric bulb is connected to a 220 V generator. The current is 0.50 A. What is the power of the bulb?

Solution:

Electric generator

voltage $V = 220 \text{ V}$, the current $I = 0.50 \text{ A}$

The power of the bulb,

$$P = VI = 220 \times 0.50 = 110 \text{ W}$$

16.12. CHEMICAL EFFECT OF ELECTRIC CURRENT

ACTIVITY 16.4

- Take out carbon rods carefully from two discarded cells.
- Clean their metal caps with sand paper.
- Wrap copper wire around the metal caps of the carbon rods.
- Connect these copper wires in series with a battery and an LED.
- Dip the carbon rods into lemon juice taken in a plastic or rubber bowl.
- Does the bulb glow?
- Does lemon juice conduct electricity?

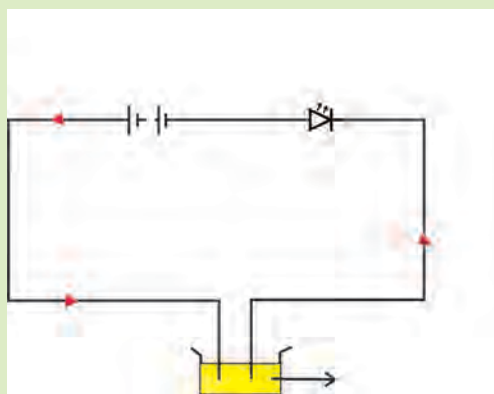



Fig. 16.8

It is observed that lemon juice conduct electricity.

16.13. ELECTROLYSIS- ELECTROCHEMICAL CELLS

When the current is passed through aqueous or molten solutions of inorganic acids, bases and salts, the conduction of electricity is always accompanied by chemical decomposition of the solutions. Such solutions are called electrolytes and the phenomenon of the conduction of electricity through electrolytes and chemical decomposition is called electrolysis.

Electro chemical cell

Name	: Alessandro Volta	
Born	: 18 February 1745	
Birth place	: Como, Italy	
Died	: 05 March 1827	
Best known for	: The Italian who built the first battery	

The cells in which the electrical energy is derived from the chemical action are called electrochemical cells.

Voltaic cell consists of two electrodes, one of copper and the other of zinc dipped in a solution of dilute sulphuric acid in a glass vessel. This is shown in Fig. 16.9.

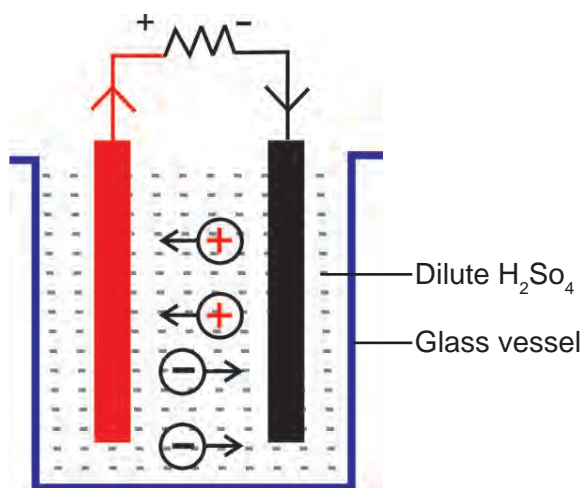


Fig. 16.9

On connecting the two electrodes externally, with a piece of wire, current flows from copper to zinc outside the cell and from zinc to copper inside it. The copper rod of the cell is the positive pole and the zinc rod of the cell is negative pole. The electrolyte is dilute sulphuric acid.

The action of the cell is explained in terms of the motion of the charged ions. At the zinc rod, the zinc atoms get ionized and pass into solution as Zn^{++} ions. This leaves the zinc rod with two electrons more, making it negative. At the same time, two hydrogen ions ($2H^+$) are discharged at the copper rod, by taking these two electrons. This makes the copper rod positive. As long as excess electrons are available on the zinc electrode, this process goes on and a current flows continuously in external circuit. This simple cell is thus seen as a device which converts chemical energy into electrical energy. Due to opposite charges on the two plates, a potential difference is set up between copper and zinc. Copper being at a higher potential than zinc, the difference of potential between the two electrodes is 1.08 V.

16.14. PRIMARY AND SECONDARY CELLS

Primary cell

The cells from which the electric energy is derived by irreversible chemical reaction are called primary cells. The primary cell is capable of giving an emf, when its constituents, two electrodes

and a suitable electrolyte, are assembled together. The main primary cells are Daniel cell and Leclanche cell. These cells cannot be recharged. Leclanche cell is discussed here.

1. Leclanche cell

A Leclanche cell consists of a glass vessel which is filled with ammonium chloride solution. Ammonium chloride solution is acting as electrolyte. In it there stands a zinc rod and porous pot containing a carbon rod which is packed round with a mixture of manganese dioxide and powdered carbon. Therefore the carbon rod forms the positive pole and the zinc rod the negative pole.

At the zinc rod, the atoms get ionised and pass in to the solution as Zn^{++} ions. This leaves the zinc rod with two electrons more making it negatively charged. At the same time,

Ammonium chloride splits into ammonia gas, two Hydrogen ions ($2H^+$) and two chloride ions ($2Cl^-$). Zn^{++} ions and $2Cl^-$ ions recombine to form zinc chloride.

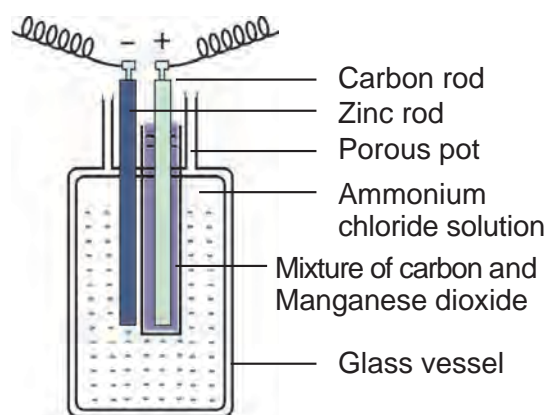


Fig. 16.10

The $2H^+$ ions migrate to the carbon rod and make it positively charged. When the carbon rod and zinc rod are connected by

a wire, the current flows from carbon to zinc through the wire. The e.m.f of the cell is about 1.5V.

Secondary cells

The advantage of secondary cell is that they are rechargeable. The chemical reactions that take place in secondary cells are reversible. The active materials that are used up when the cell delivers current can be reproduced by passing current through the cell in opposite direction. The chemical process of obtaining current from a secondary cell is called discharge. The process of reproducing active materials is called charging. One of the most commonly used secondary cell is lead acid accumulator.

Lead-acid accumulator

In a lead-acid accumulator, the anode and cathode are made of lead dioxide and lead respectively. The electrolyte is dilute sulphuric acid. As power is discharged from the accumulator, both the anode and

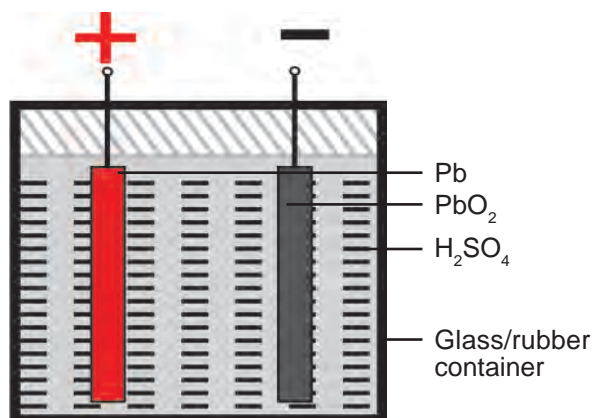


Fig. 16.11

cathode undergoes a chemical reaction that progressively changes them into lead sulphate. When the anode and cathode



are connected by a wire, the current flows from cathode to anode through the wire.

When current is applied to a lead-acid accumulator, the electrochemical reaction is reversed. This is known as recharging of the accumulator. The e.m.f of freshly charged cell is 2.2V.

16.15. SOURCES OF ENERGY

Energy comes from different forms and one can be converted to another. If energy can neither be created nor be destroyed. We should be able to perform endless activities without thinking about energy resources. But we hear so much about the energy crises. What is the reason?

If we drop a plate from a height, the potential energy of the plate is converted mostly to sound energy when it hit's the ground. If we light a candle the chemical energy in the wax is converted to heat energy and light energy on burning.

In these examples we see that energy, in the usable form is dissipated to the surroundings in less usable forms. Hence any source of energy we use to do work is consumed and cannot be used again. We use muscular energy for carrying out physical work, electrical energy for running various appliances, chemical energy for cooking food or running a vehicle, all come from a source. We should know

how to select the source needed for obtaining energy in its usable form, and then only it will be a useful source.

A good source of energy would be one

- Which would do a large amount of work per unit volume of mass?
- Be easily accessible.
- Be easy to store and transport and
- Perhaps most importantly be economical.

16.15.1. Conventional-sources of energy

1. Fossil fuels

In ancient time's wood was the most common source of energy. The energy of flowing water and wind was also used for limited activities. Can you think of some of these uses? The exploitation of coal as a source of energy made the industrial revolution possible. Industrialisation has caused the global demand for energy to grow at a tremendous rate. The growing demand for energy was largely met by the fossil fuels, coal and petroleum. These fuels were formed over millions of years ago and there are only limited reserves. The fossil fuels are non-renewable sources of energy. So we need to conserve them. If we were to continue consuming these sources at such alarming rates we would soon run out of the energy. In order to avoid this alternate source of energy were explored.

Burning fossil fuels has other disadvantages like air pollution, acid rain and production of green house gases.

We will see how various sources of energy can be used to run the turbine and generate electricity in the following sections.

2. Thermal power plant

Large amount of fossil fuels are burnt everyday in power stations to heat up water to produce steam which further runs the turbine to generate electricity. The transmission of electricity is more efficient than transporting coal or petroleum over the same distance. Therefore, many thermal power plants are set up near coal or oil fields. The term thermal power plant is used since fuel is burnt to produce heat energy which is converted into electrical energy.

3. Hydro power plants

Another traditional source of energy was the kinetic energy of flowing water or the potential energy of water at a height. Hydro power plants convert the potential energy of falling water into electricity. Since there are very few water falls which could be used as a source of potential energy, hydro power plants are associated with dams. In the last century, a large number of dams were built all over the world. As we can see, a quarter of our energy requirements in India is met by hydro power plants. In order to produce hydro electricity, high-rise dams are constructed on the river to obstruct the flow of water and there by collect water in larger reservoirs. The water level rises and in this process the kinetic energy of flowing water gets transformed into potential energy. The water from the high level in the dam

is carried through the pipes, to the turbine, at the bottom of the dam Fig.16.12. since the water in the reservoir would be refilled each time it rains(hydro power is a renewable source of energy) we would not have to worry about hydro electricity sources getting used up the way fossil fuels would get finished one day.

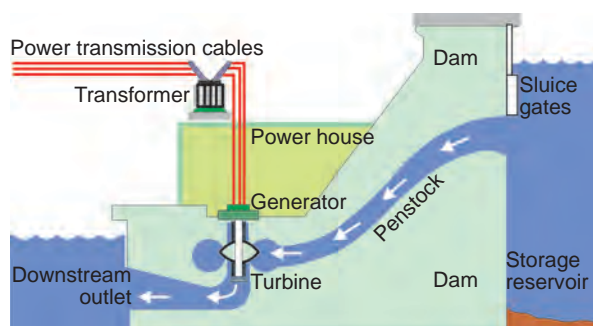


Fig. 16.12

4. Bio-mass

We mentioned earlier that wood has been used as a fuel for a long time. If we can ensure that enough trees are planted, a continuous supply of fire-wood can be assured. You must also be familiar with the use of cow-dung cakes as a fuel. Given the large-stock published in India, this can also assure us a steady source of fuel. Since these fuels are plant and animal products, the source of these fuels is set to be bio-mass. These fuels, however, do not produce much heat on burning and a lot of smoke is given out when they are burnt. Therefore, technological inputs to improve the efficiency of these fuels are necessary. When wood is burnt in a limited supply of oxygen, water and volatile materials present in it get removed and charcoal is left behind as the residue. Charcoal burns without flames, is comparatively smokeless and has higher heat generation efficiency.

Similarly, cow-dung, various plant materials like the residue after harvesting the crops, vegetable wastes and sewage are decomposed in the absence of oxygen to give bio-gas. Since the starting material is mainly cow-dung, it is popularly known as gobar gas. The gobar gas plant structure is shown in Fig. 16.13.

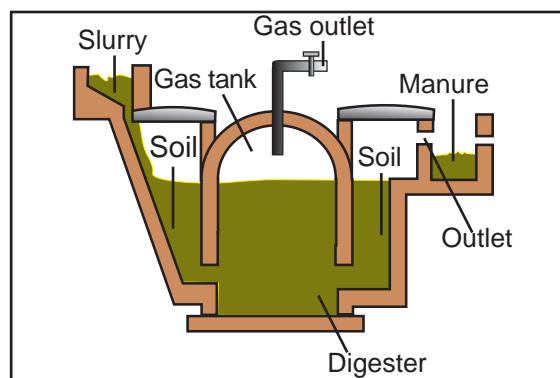


Fig 16.13

5. Wind energy

The kinetic energy of the wind can be used to do work. This energy was harnessed by wind mills in the past to do mechanical work. For example, in a water-lifting pump, the rotatory motion of windmill is utilized to lift water from a well. Today, wind energy is also used to generate electricity. A wind mill essentially consists of a structure similar to a large electric fan that is erected at some height on a rigid support.

To generate electricity, the rotatory motion of the windmill is used to turn the turbine of the electric generator. The output of a single windmill is quiet small and cannot be used for commercial purposes. Therefore, a number of windmills are erected over a large area, which is known as wind energy farm. The energy output of each windmill in a farm is coupled to-

gether to get electricity on a commercial scale.

Wind energy is a environment-friendly and efficient source of renewable energy. It requires no recurring expenses for the production of electricity. The wind speed should be higher



Fig. 16.14

than 15 km per hour to maintain the required speed of the turbine. Fig. 16.14.

16.15.2. Non-conventional sources of energy

Our life-styles are changing; we use machines to do more and more of our tasks. Therefore our demand for the energy increases. We need to look for more and more sources of energy. We could develop the technology to use the available sources of energy more efficiently and also look to new sources of energy. We shall now look at some of the latest sources of energy.

1. Solar energy

The sun has been radiating an enormous amount of energy at the present

ACTIVITY 16.5

- Find out from your grand-parents or other elders
 - (a) How did they go to school?
 - (b) How did they get water for their daily needs when they were young?
 - (c) What means of entertainment did they use?
- Compare the above answers with how you do these tasks now.
- Is there a difference? If yes, in which case more energy from external sources is consumed?

rate for nearly 5 billion years and will continue radiating at that rate for about 5 billion years more. Only a small part of solar energy reaches the outer layer of the earth atmosphere. Nearly half of it is absorbed while passing through the atmosphere and the rest reaches the earth's surface.

A black surface absorbs more heat than any other surface under identical conditions. Solar cookers and solar water heaters use this property in their working. Some solar cookers achieve a higher temperature by using mirrors to focus the rays of the sun. solar cookers are covered with a glass plate.

These devices are useful only at certain times during the day. This limitation of using solar energy is overcome by using solar cells that convert solar energy into electricity. A large number of solar cells are combined in a arrangement called solar

ACTIVITY 16.6

- Take two conical flasks and paint one white and the other black. Fill both with water.
- Place the conical flask in direct sunlight for half an hour to one hour.
- Touch the conical flasks. Which one is hotter? You could also measure the temperature of the water in the two conical flasks with a thermometer.
- Can you think of ways in which this finding could be used in your daily life?

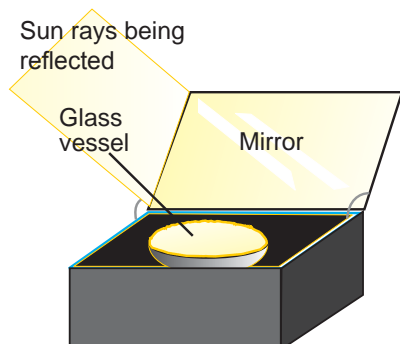


Fig. 16.15

cell panel that can deliver enough electricity for practical use Fig. 16.16. The principal advantages associated with solar cells are that they have no moving part, require little maintenance. Another advantage is that they can be set up in remote areas in which laying of power transmission line may be expensive.

ACTIVITY 16.7

- Study the structure and working of a solar cooker or a solar water-heater, particularly with regard to how it is insulated and maximum heat absorption is ensured.

- Design and build a solar cooker or water-heater using low-cost material available and check what temperature are achieved in your solar system.
- Discuss what would be the advantages and limitations of using the solar cooker or water-heater.

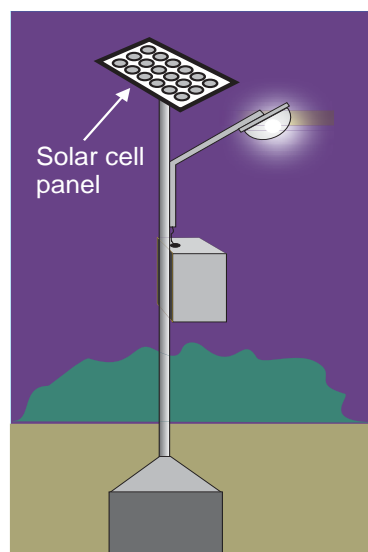


Fig 16.16

16.15.3. Nuclear energy

How is nuclear energy generated? In a process called nuclear fission, the nucleus of a heavy atom (such as uranium, plutonium or thorium), when bombarded with low-energy neutrons, can be split apart into lighter nuclei. When this is done, a tremendous amount of energy is released if the mass of the original nucleus is just a little more than the sum of the masses of the individual products. The fission of an atom of uranium, for example, produces 10 million times the energy produced by the combustion of an atom of carbon from

coal. In a nuclear reactor designed for electric power generation sustained fission chain reaction releases energy in a controlled manner and the released energy can be used to produce steam and further generate electricity.

16.15.4. Radioactivity

Name	: Henry Becquerel
Born	: 15 December 1852
Birth place	: Paris, France
Died	: 25 August 1908
Best known for	: Discovery of radioactivity



The phenomenon of radioactivity was discovered by Henry Becquerel in 1896. He found that a photographic plate wrapped in a black paper was affected by certain penetrating radiations emitted by uranium salt. Rutherford showed later that the radiations from the salt were capable of ionizing a gas. The current produced due to the ions was taken as a measure of activity of the compound.

A few years later Madam Marie Curie and her husband Pierre Curie discovered the highly radioactive elements radium and polonium. The activity of the material has been shown to be the result of the three different kinds of radiations, α , β , and γ .

The phenomenon of spontaneous emission of highly penetrating radiations such as α , β , and γ rays by heavy elements having atomic number greater than 82 is called radioactivity and the substances which emit these radiations are called radioactive elements.

The radioactive phenomenon is spontaneous and is unaffected by any external

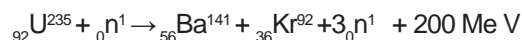
agent like temperature, pressure, electric and magnetic fields etc.

16.15.5. Nuclear fission and nuclear fusion

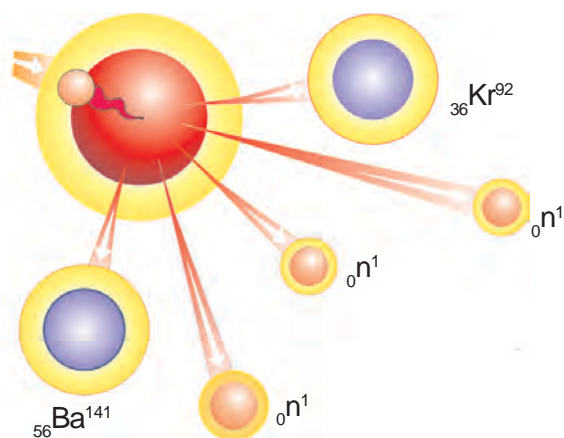
1. Nuclear fission

In 1939, German scientists Otto Hahn and Strassman discovered that when uranium nucleus is bombarded with a neutron, it breaks up into two fragments of comparable masses with the release of energy.

The process of breaking up of the nucleus of a heavier atom into two fragments with the release of large amount of energy is called nuclear fission. The fission is accompanied of the release of neutrons. The fission reactions with ${}_{92}\text{U}^{235}$ are represented as



In the above example the fission reaction is taking place with the release of 3 neutrons and 200 Million electron volt energy.



The process of fission
Fig. 16.17

2. Nuclear fusion

Nuclear fusion is a process in which two or more lighter nuclei combine to form a heavier nucleus. The mass of the product is always less than the sum of the masses of the individual lighter nuclei. According to Einstein's mass energy relation $E = mc^2$, the difference in mass is converted into energy. The fusion process can be carried out only at extremely high temperature of the order of 10^7 K because, only at these very high temperatures the nuclei are able to overcome their mutual repulsion. Therefore before fusion, the lighter nuclei must have their temperature raised by several million degrees. The nuclear fusion reactions are known as thermo nuclear reactions.

A suitable assembly of deuteron and triton is arranged at the sight of the explosion of the atom bomb. Favorable temperature initiates the fusion of light nuclei in an uncontrolled manner. This releases enormous amount of heat energy. This is the hydrogen bomb.

The fusion reaction in the hydrogen bomb is ${}_1\text{H}^2 + {}_1\text{H}^3 \rightarrow {}_2\text{He}^4 + {}_0\text{n}^1 + \text{Energy}$

Example: 16.8

Calculate the energy produced when 1 kg of substance is fully converted into energy.

Solution:

Energy produced,	$E = mc^2$
Mass,	$m = 1 \text{ kg}$
Velocity of light,	$c = 3 \times 10^8 \text{ m s}^{-1}$
	$E = 1 \times (3 \times 10^8)^2$
	$E = 9 \times 10^{16} \text{ J}$

16.15.6. Nuclear Reactivity Advantages

Nuclear reactivity is a measure of the departure of a reactor from criticality. It is a useful concept to predict how the neutron population of a reactor will change over time.

If a reactor is exactly critical, that is, the neutron production is exactly equal to the neutron destruction, then the reactivity is zero. If the reactivity is positive, then the reactor is super critical. If the reactivity is negative, then the reactor is sub critical.

16.15.7. Hazards of nuclear energy

α , β and γ radiations are all ionizing radiations. These radiations cause a change in the structure of molecules in cells, disturbs the normal functioning of the biological system. The extent to which the human organism is damaged depends upon

1. The dose and the rate at which the radiation is given and
2. The part of the body exposed to it. The damage may be either pathological or genetic.

The radiation exposure is measured by the unit called roentgen(R). One roentgen is defined as the quantity of radiation which produces 1.6×10^{12} pairs of ion in 1 gram of air.

Safe limit of receiving the radiation is about 250 milli roentgen per week.

The following precautions are to be taken for those, who are working in radiation laboratories.

- (i) Radioactive materials are kept in thick-walled lead container.
- (ii) Lead aprons and lead gloves are used while working in hazardous area.
- (iii) A small micro-film badge is always worn by the person and it is checked periodically for the safety limit of radiation.
- (iv) Nuclear devices can be operated using remote control system.
- (v) Clean up contamination in the work area promptly.

16.15.8. SCIENCE TODAY - Energy from seas

1. Tidal energy

Due to the gravitational pull of mainly the moon on the spinning earth, the level of the water in the sea rises and falls. If you live near the sea or ever travel to some place near the sea, try and observe how the sea-level changes during the day. The

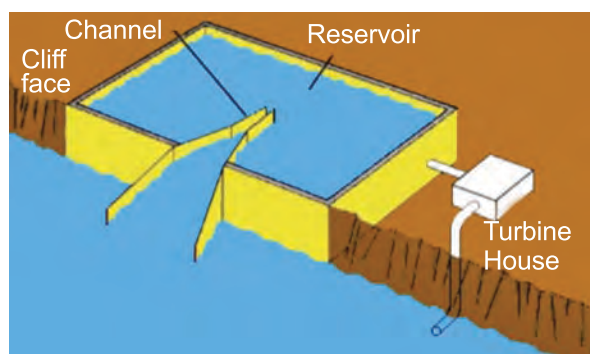


Fig. 16.18

phenomenon is called high and low tides and the difference in sea-levels gives us tidal energy. Tidal energy is harnessed by constructing a dam across a narrow open-

ing to the sea. A turbine fixed at the opening of the dam converts tidal energy to electricity. Fig. 16.18. As you can guess, the locations where such dams can be built are limited.

2. Wave energy

Similarly, the kinetic energy possessed by huge waves near the sea-shore can be trapped in a similar manner that generates electricity. The waves are generated by strong winds blowing across the sea. Wave energy would be a viable proposition only where waves are very strong. A wide variety of devices has been developed to trap wave energy for rotation of turbine and production of electricity. Fig.16.19

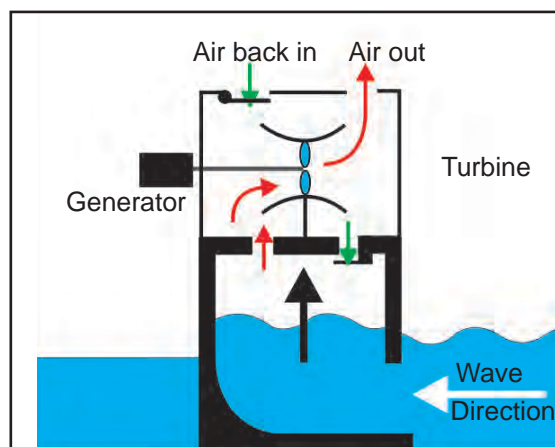


Fig. 16.19

3. Ocean thermal energy

The water at the surface of the sea or ocean is heated by the sun while the water in deeper sections is relatively cooled. This difference in temperature is exploited to obtain energy in ocean-thermal-energy conversion plants. These plants can operate if the temperature difference between the water at the surface and water at depths up to 2 kilometers is 293 K (20° C) or more. The warm surface-water is used

to boil a volatile liquid like ammonia. The vapours of liquid then used to run the turbine of generator. The cooled water from the depth of the ocean is pumped up and condense vapour again to liquid. Fig.16.20.

The energy potential from the sea (tidal energy, wave energy and ocean thermal energy) is quite large, but efficient commercial exploitation is difficult.

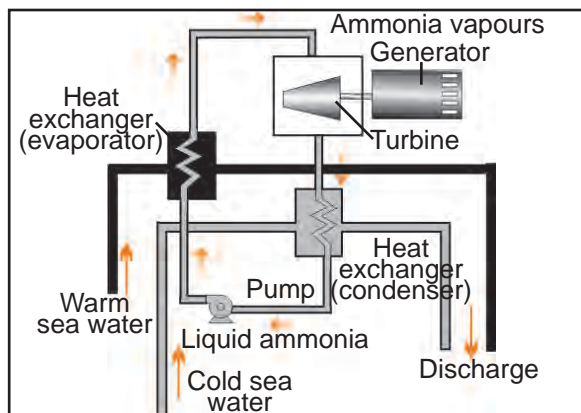


Fig. 16.20

EVALUATION

PART A

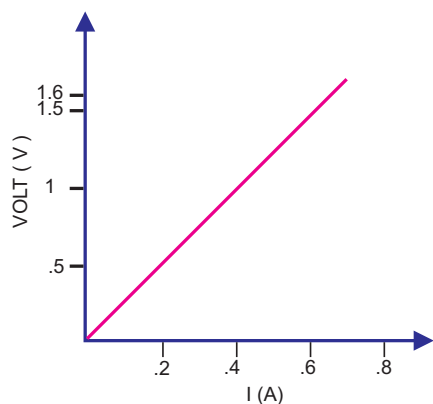
- The potential difference required to pass a current 0.2 A in a wire of resistance 20 ohm is _____. (100 V, 4 V, 0.01 V, 40 V)
- Two electric bulbs have resistances in the ratio 1 : 2. If they are joined in series, the energy consumed in these are in the ratio _____. (1 : 2, 2 : 1, 4 : 1, 1 : 1)
- Kilowatt-hour is the unit of _____. (potential difference, electric power, electric energy, charge)
- _____ surface absorbs more heat than any other surface under identical conditions. (White, rough, black, yellow)
- The atomic number of natural radioactive element is _____. (greater than 82, less than 82, not defined, atleast 92)

PART B

- From the following statements write down that which does not represent ohm's law.
 - current / potential difference = constant
 - potential difference / current = constant
 - current = resistance x potential difference

- Fill in the blanks
 - Potential difference : voltmeter, then: current _____.
 - power plant : conventional source of energy then solar energy _____.
- In the list of sources of energy given below, some of them are wrong. List out the wrong ones. (Wind energy, solar energy, hydro electric power, nuclear energy, tidal energy, wave energy, geo-thermal energy.)
- Correct the mistakes, if any, in the following statements.
 - A good source of energy would be one which would do a small amount of work per unit volume of mass.
 - Any source of energy we use to do work is consumed and can be used again.

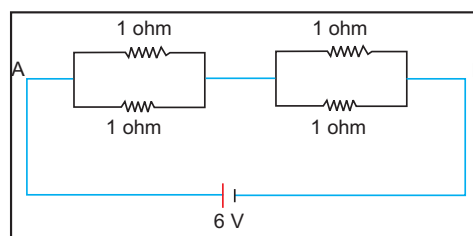
5. The schematic diagram, in which different components of the circuit are represented by the symbols conveniently used, is called a circuit diagram. What do you mean by the term components?
6. Following graph was plotted between V and I values. What would be the values of V/I ratios when the potential difference is 0.8 V and 1.2 V.



7. We know that γ – rays are harmful radiations emitted by natural radio active substances.
- Which are other radiations from such substances?
 - Tabulate the following statements as applicable to each of the above radiations

They are electromagnetic radiation. They have high penetrating power. They are electrons. They contain neutrons.

8. Draw the schematic diagram of an electric circuit consisting of a battery of two cells of 1.5V each, three resistance of 5 ohm, 10 ohm and 15 ohm respectively and a plug key all connected in series.
9. Fuse wire is made up of an alloy of _____ which has high resistance and _____.
10. Observe the circuit given below and find the resistance across AB.



11. Complete the table choosing the right terms from within the brackets. (zinc, copper, carbon, lead, leadoxide, aluminium.)

+ ve electrode	Daniel cell	
- ve electrode	Lechlanche cell	

FURTHER REFERENCE

Books : 1. Electricity and Magnetism, by **D.C Tayal** Himalayam publishing house.

2. Sources of energy, by **C. Walker**, Modern curriculam press.

Website : www.reprise.com, www.wikipedia.org



**MAGNETIC EFFECT OF
ELECTRIC CURRENT
AND LIGHT**

17. MAGNETIC EFFECT OF ELECTRIC CURRENT AND LIGHT



Name : Oersted
Born : 14 August 1777
Birth place : Langeland Denmark
Died : 9 March 1851
Best known for : The study of electromagnetism

17.1. MAGNETIC FIELD AND MAGNETIC LINES OF FORCE

We are familiar with the fact that a compass needle gets deflected when brought near a bar magnet. Why does a compass needle get deflected?

ACTIVITY 17.1

- Fix a sheet of white paper on a drawing board using some adhesive material.
- Place a bar magnet in the centre of it.
- Sprinkle some iron filings uniformly around the bar magnet (Fig 17.1).
- A salt-sprinkler may be used for this purpose.
- Now tap the board gently.
- What do you observe?

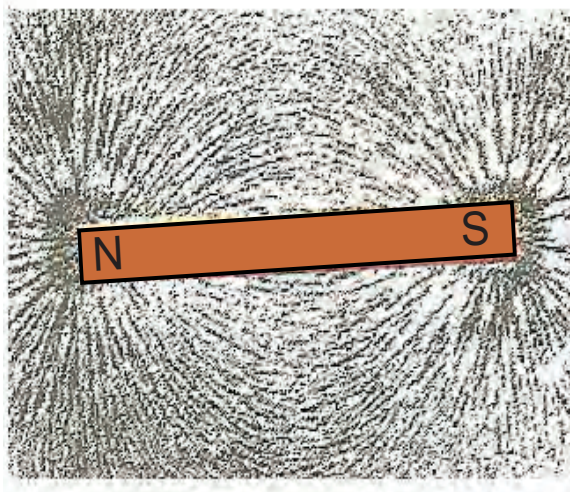


Fig. 17.1

The iron filings arrange themselves in a pattern as shown in Fig. 17.1. Why do the iron filings arrange in such a pattern? What does this pattern demonstrate? The magnet exerts its influence in the region surrounding it. Therefore the iron filings experience a force. The force thus exerted makes iron filings to arrange in a pattern. The region surrounding the magnet, in which the force of the magnet can be detected, is said to have a **magnetic field**. The lines along which the iron filings align themselves represent **magnetic lines of force**.

ACTIVITY 17.2

- Take a small compass and a bar magnet.
- Place the magnet on a sheet of white paper fixed on a drawing board, using some adhesive material.
- Mark the boundary of the magnet.
- Place the compass near the north pole of the magnet. How does it behave? The south pole of the needle points towards the north pole of the magnet. The north pole of the compass is directed away from the north pole of the magnet.
- Mark the position of two ends of the needle.
- Now move the needle to a new position such that its south occupies the position previously occupied by its north pole.
- In this way, proceed step by step till you reach the south pole of the magnet as shown
- Join the points marked on the paper by a smooth curve. This curve represents a field line.
- Repeat the above procedure and draw as many lines as you can. You will get a pattern shown in Fig.17.2. These lines represent the magnetic field around the magnet. These are known as magnetic field lines.
- Observe the deflection of the compass needle as you move it along the field line. The deflection increases as the needle is moved towards the pole.

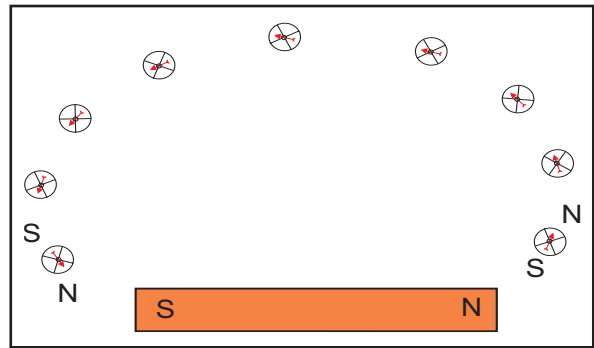


Fig 17.2

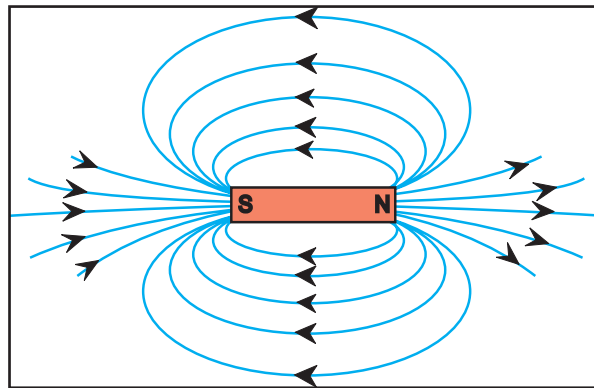


Fig 17.3

Magnetic field is a quantity that has both magnitude and direction. The direction of the magnetic field is taken to be the direction in which a north pole of the compass needle moves inside it. Therefore it is taken by convention that the field lines emerge from the north pole and merge at the south pole as shown in Fig.17.3. Inside the magnet, the direction of field lines is from its south pole to its north pole. Thus the magnetic field lines are closed curves. No two field-lines are found to cross each other.

17.2. MAGNETIC FIELD DUE TO CURRENT CARRYING CONDUCTOR

In the activity 17.3 we have seen that electric current through a metallic conductor

ACTIVITY 17.3

- Take a straight thick copper wire and place it between the points X and Y in an electric circuit, as shown in Fig..17.4. The wire XY is kept perpendicular to the plane of the paper.
- Horizontally place a small compass near this copper wire. See the position of its needle.
- Pass the current through the circuit by inserting the key into the plug.
- Observe the change in the position of the compass needle and the direction of deflection.
- Replace the cell connection in the circuit so that the direction of the current in the copper wire changes.
- Observe the change in the direction of deflection of the needle.

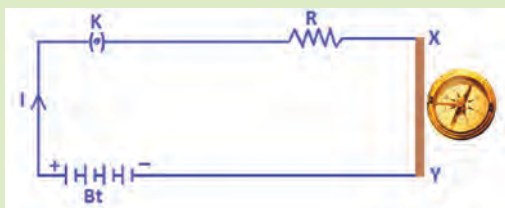


Fig 17.4

produces a magnetic field around it. If the current flows in one direction (from X to Y), the north pole of the compass needle would move towards the east. If the current flows in opposite direction (from Y to X), you will see that the needle moves in opposite direction, that is towards the west. It means that the direction of magnetic field

produced by the electric current depends upon the direction of flow of current.

17.2.1. Magnetic field due to current carrying straight conductor

What determines the pattern of the magnetic field generated by current through a conductor? Does the pattern depend on the shape of the conductor? We shall investigate this with an activity.

ACTIVITY 17.4

- Take a battery (12 V), a variable resistance (rheostat), an ammeter (0-5A), a plug key, and a long straight thick copper wire.
- Insert the thick wire through the centre, normal to the plane of a rectangular cardboard. Take care that the cardboard is fixed and does not slide up or down.
- Connect the copper wire vertically between the points X and Y, as shown in Fig 17.5(a), in series with the battery, a plug key, ammeter and a rheostat.
- Sprinkle some iron fillings uniformly on the cardboard. (you may use a salt sprinkler for this purpose).
- Keep the variable of the rheostat at a fixed position and note the current through the ammeter.
- Close the key so that the current flows through the wire. Ensure that the copper wire placed between the points X and Y remains vertically straight.

- Gently tap the cardboard a few times. Observe the pattern of the iron fillings. You would find that the iron fillings align themselves showing a pattern of concentric circles around the copper wire, Fig 17.5(b).
- What do these concentric circles represent? They represent the magnetic field lines.
- How can the direction of the magnetic field be found? Place a compass at a point (say P) over a circle. Observe the direction of the needle. The direction of the north pole of the compass needle would give the direction of the field lines produced by the electric current through the straight wire at point P. Show the direction by an arrow.
- Does the direction of magnetic field lines get reversed if the direction of current through the straight copper wire is reversed? Check it.

What happens to the deflection of the compass needle placed at a given point if the current in the copper wire is changed? We find that the deflection in the needle also changes. In fact, if the current is increased, the deflection also increases. It indicates that the magnitude of the magnetic field produced at a given point increases as the current through the wire, increases.

What happens to the deflection of the needle if the compass is moved from the copper wire but the current through the wire remains the same? To see this, now place the compass at a farther point from the conducting wire. What change do you observe? We see that the deflection in the needle decreases. Thus the magnetic field produced by the given current in the conductor decreases as the distance from it increases. From Fig.17.5 (b), it can be noticed that the concentric circles representing the magnetic field around a current-carrying straight wire become larger and larger as we move away from it.

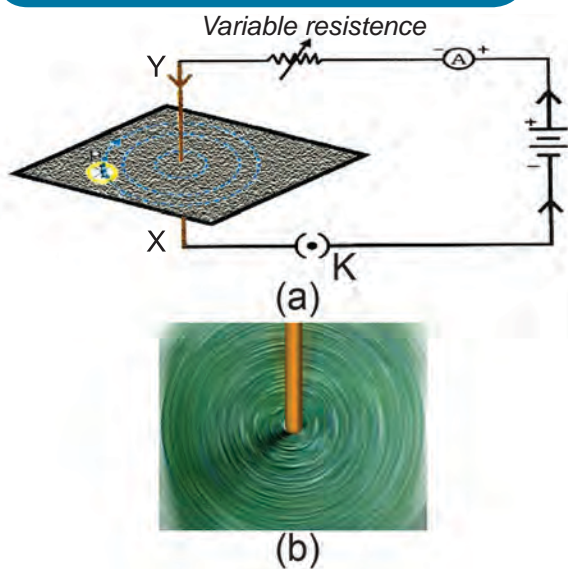


Fig. 17.5

17.2.2. Magnetic field due to current carrying circular loop

We have so far observed the pattern of the magnetic field lines produced around a current-carrying straight wire. Suppose this straight wire is bent in the form of a circular loop and current is passed through it, how would the magnetic field lines look like?

We know that the magnetic field produced by a current-carrying straight wire depends inversely on the distance

ACTIVITY 17.5

- Take a rectangular cardboard having two holes. Insert a circular coil having large number of turns through them, normal to the plane of the cardboard.
- Connect the ends of the coil in series with a battery, a key and rheostat, as shown in Fig.17.6.
- Sprinkle iron fillings uniformly on the cardboard.
- Plug the key.
- Tap the cardboard gently a few times. Note the pattern of the iron fillings that emerges on the cardboard.

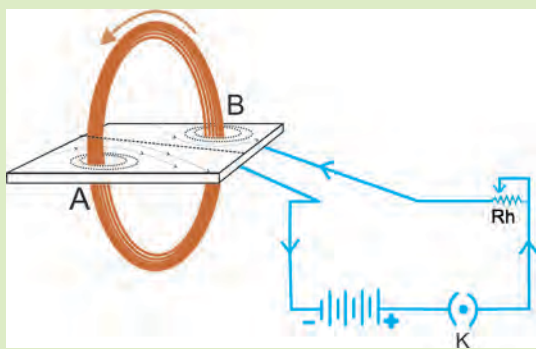


Fig. 17.6

from it. Similarly at every point of a current-carrying circular loop, the concentric circles representing the magnetic field around it becomes larger and larger as we move away from the wire (Fig. 17.7).

By the time we reach the centre of the circular loop, the arcs of these big circles would appear as straight lines. Every point on the wire carrying current would give rise to the magnetic field appearing as straight lines at the centre of the loop.

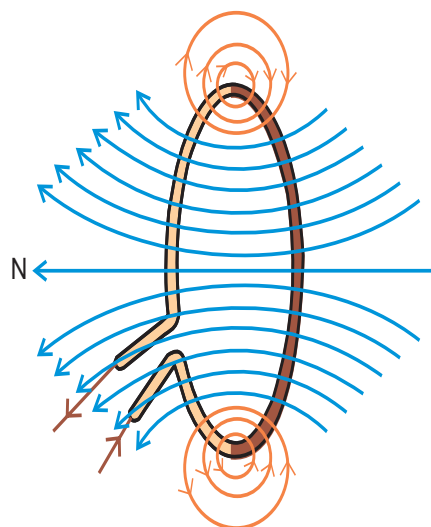


Fig. 17.7

We know that the magnetic field produced by a current-carrying conductor at a given point, depends directly on the current passing through it. Therefore, if there is a circular coil having n turns, the field produced is n times as large as produced by a single turn. This is because the current in each circular turn has the same direction, and the field due to each turn then just adds up.

17.3. FORCE ON A CURRENT CARRYING CONDUCTOR IN A MAGNETIC FIELD

We know that an electric current flowing through a conductor produces a magnetic field. The field so produced exerts a force on a magnet placed in the vicinity of a conductor. French scientist Andre Marie Ampere suggested that the magnet must also exert an equal and opposite force on the current carrying conductor. The force due to a current carrying conductor can be demonstrated through the following activity.

ACTIVITY 17.6

- Take a small aluminium rod AB of about 5 cm. using two connecting wires suspend it horizontally from a stand as shown in Fig. 17.8.
- place a horse-shoe magnet in such a way that the rod lies between two poles with the magnetic field directed upwards. For this put the North Pole of the magnet vertically below and South Pole vertically above the aluminium rod.
- Connect the aluminium rod in series with a battery, a key and a rheostat.
- Now pass a current through the aluminium rod from end B to A.
- What do you observe? It is observed that the rod is displaced towards the left. You will notice that the rod gets displaced.
- Reverse the direction of current flowing through the rod and observe the direction of its displacement. It is now towards the right.
- Why does the rod get displaced?

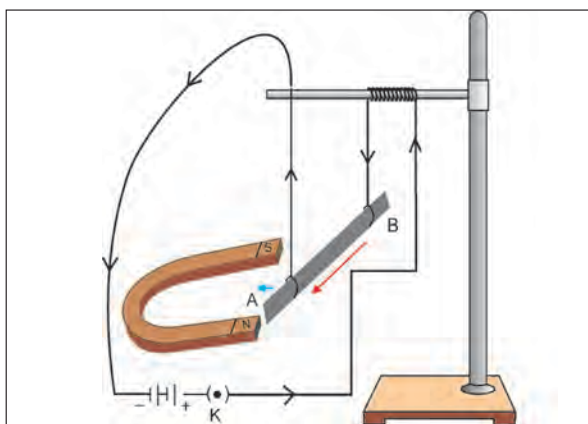


Fig. 17.8

The displacement of the rod in the above activity suggests that a force is exerted on the current-carrying aluminium rod when it is placed on a magnetic field. It also suggests that the direction of force is also reversed when the direction of current through the conductor is reversed. Now change the direction of field to vertically downwards by interchanging the two poles of the magnet. It is once again observed that the direction of force acting on the current-carrying rod gets reversed. It shows that the direction of force on the conductor depends upon the direction of current and the direction of magnetic field. Experiments have shown that the displacement of the rod is largest when the direction of current is at right angles to the direction of magnetic field.

17.3.1. Fleming left hand rule

We considered that the direction of the current and that of the magnetic field perpendicular to each other and found that the force is perpendicular to both of them. The three directions can be illustrated through a simple rule, called Fleming's left hand rule.(Fig.17.9).

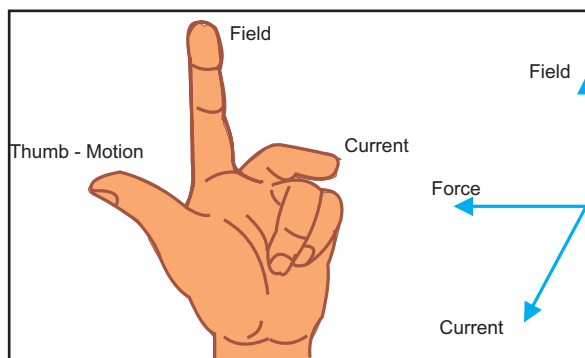


Fig. 17.9

Stretch the thumb, fore finger and middle finger of your left hand such

that they are mutually perpendicular. If the forefinger points in the direction of magnetic field and the middle finger points in the direction of current, then the thumb will point in the direction of motion or the force acting on the conductor.

17.4. ELECTRIC MOTOR

An electric motor is a rotating device that converts electrical energy in to mechanical energy. Do you know how an electric motor works?

An electric motor, as shown in Fig. 17.10, consists of a rectangular coil ABCD of insulated copper wire. The coil is placed between two poles of a magnetic field such that the arm AB and CD are perpendicular to the direction of magnetic field. The ends of the coil are connected to the two halves S_1 and S_2 of a split ring. The inner side of these halves insulated and attached to an axle. The external conducting edges of S_1 and S_2 touch two conducting stationary brushes B_1 and B_2 , respectively.

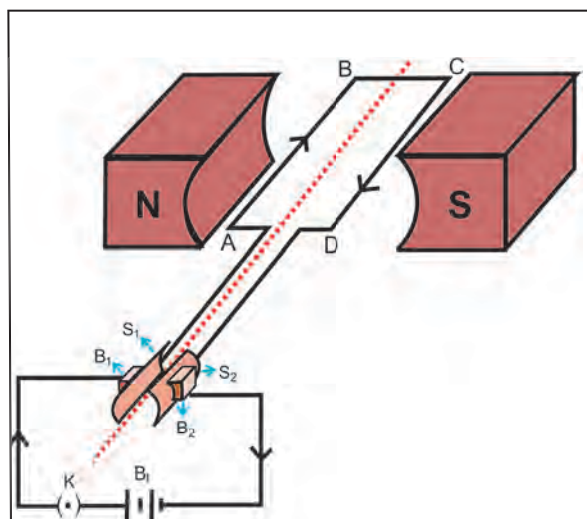


Fig. 17.10

Current in the coil ABCD enters from the source battery through conducting brush B_1 and flows back to the battery through brush B_2 . Notice that the current in arm AB of the coil flows from A to B. In arm CD it flows from C to D, that is, opposite to the direction of current through arm AB. On applying Fleming's left hand rule for the direction of force on a current-carrying conductor in a magnetic field. We find that the force acting on arm AB pushes it downwards while the force acting on arm CD pushes it upwards. Thus the coil and the axle, mounted free to turn about an axis, rotate anti-clockwise. At half rotation S_2 makes contact with the brush B_1 and S_1 with brush B_2 . Therefore the current in the coil gets reversed and flows along the path DCBA. A device that reverses the direction of flow of current through a circuit is called a commutator. In electric motors, the split ring acts as a commutator. The reversal of current also reverses the direction of force acting on the two arms AB and CD. Thus the arm AB of the coil that was earlier pushed down is now pushed up and the arm CD previously pushed up is now pushed down. Therefore the coil and the axle rotate half a turn more in the same direction. The reversing of the current is repeated at each half rotation, giving rise to a continuous rotation of the coil and to the axle.

The commercial motors use (i) an electro magnet in place of permanent magnet; (ii) large number of turns of the conducting wire in the current-carrying coil, and (iii) a soft iron core on which the coil is wound. The soft iron core, on which the coil is wound, is called an armature. This enhances the power of the motor.

17.5. ELECTROMAGNETIC INDUCTION

Faraday in 1831 discovered that an electro motive force is produced in a circuit whenever the magnetic flux linked with a coil changes. He showed that emf is generated in a conductor when ever there is a relative motion between the conductor and a magnetic field. Then emf produced in this way is called an induced emf and the phenomenon is known as electro magnetic induction. The induced emf will cause a current to flow through the conductor. Such a current is known as induced current .Faraday made an important breakthrough by discovering how a magnet can be used to generate electric currents.

17.5.1. Faraday's Experiments

We know that when a current-carrying conductor is placed in a magnetic field, it experiences a force. This force causes the conductor to move. Now let us imagine a situation in which a conductor is moving inside a magnetic field or a magnetic field is changing around a fixed conductor. What will happen? To observe this effect, let us perform the following activity.

ACTIVITY 17.7

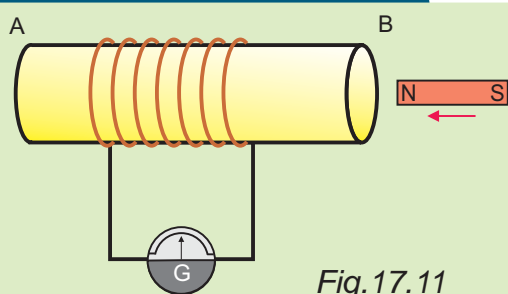


Fig.17.11

- Take a coil of wire AB having a large number of turns.

- Connect the ends of the coil to a galvanometer as shown in Fig.17.11
- Take a strong bar magnet and move its north pole towards the end B of the coil. Do you find any change in the galvanometer reading?
- There is a momentary deflection in the needle of the galvanometer, say to the right. This indicates the presence of a current in the coil AB. The deflection becomes zero, the moment the motion of the magnet stops.
- Now withdraw the north pole of the magnet away from the coil. Now the galvanometer is deflected towards the left, showing that the current is now setup in the direction opposite to the first.
- Place the magnet stationary at the point near to the coil, keeping its north pole towards the end B of the coil. We see that the galvanometer needle deflects towards the right when the coil is moved towards the north pole of the magnet. Similarly the needle moves towards left when the coil is moved away.
- When the coil is kept stationary with respect to the magnet, the deflection of the galvanometer drops to zero. What do you conclude from this activity?

You can also check that if you have moved South Pole of the magnet towards the end B of the coil, the deflections in the galvanometer would just be opposite to the previous case. When the coil and the magnet are both stationary, there is no deflection in the galvanometer. It is thus clear that motion of a magnet with respect to the coil produces an induced electromotive force, which sets up an induced electric current in the circuit.

ACTIVITY 17.8

- Two different coils of copper wire having large number of turns (say 50 and 100 turns respectively). Insert them over a non conducting cylindrical roll as shown in Fig.17.12.
- Connect the coil -1 having large number of turns, in series with a battery and a plug key. Also connect the other coil -2 with a galvanometer.
- plug in the key. Observe the galvanometer. Is there a deflection in its needle?. You will observe that the needle of the galvanometer instantly jumps to one side and just as quickly returns to zero, indicating a momentary current in coil -2.
- Disconnect coil-1 from the battery. You will observe that the needle momentarily moves, but to the opposite side. It means that, Now the current flows in the opposite direction in coil -2.

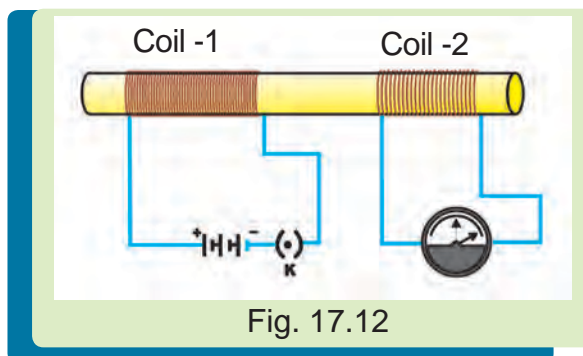


Fig. 17.12

Let us now perform a different activity in which the moving magnet is replaced by a current-carrying coil and the current in the coil can be varied.

In this activity we observe that as soon as the current in coil-1 reaches either a steady value or zero, the galvanometer in coil-2 shows no deflection. From these observations we conclude that a potential difference is induced in coil-2 whenever the current through the coil-1 is changing. Coil-1 is called the primary coil and coil-2 is called the secondary coil. As the current in the first coil changes, the magnetic field associated with it also changes. Thus the magnetic field lines around the secondary coil also change. Hence the change in magnetic field lines associated with the secondary coil is the cause of induced electric current in it. The direction of the induced current can be found using **Fleming's right hand rule**.

Stretch the thumb, forefinger and middle finger of right hand so that they are perpendicular to each other. If the forefinger indicates the direction of the magnetic field and the thumb shows the direction of motion of conductor, then the middle finger will show the direction of induced current.

17.6. ELECTRIC GENERATOR

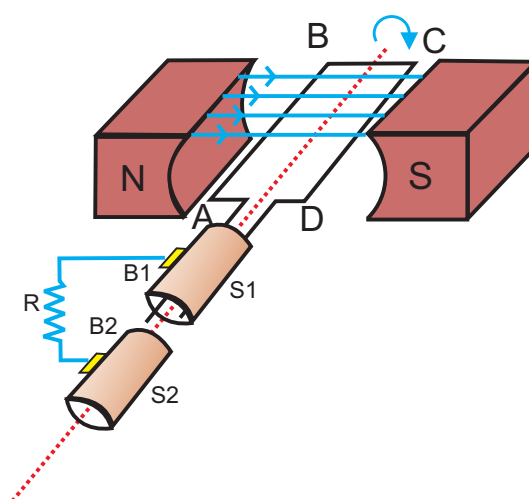
The phenomenon of electro magnetic induction is employed to produce large currents for use in homes and industry. In an electric generator, mechanical energy is used to rotate a conductor in a magnetic field to produce electricity.

An electric generator, as shown in Fig.17.13a, consists of rotating rectangular coil ABCD placed between the two poles of a permanent magnet. The two ends of this coil are connected to the two rings S_1 and S_2 . The inner sides of these rings are made insulated. The two conducting stationary brushes B_1 and B_2 are kept pressed separately on the rings S_1 and S_2 respectively. The two rings S_1 and S_2 are internally attached to an axle. The axle may be mechanically rotated from outside to rotate the coil inside the magnetic field. Outer ends of the two brushes are connected to the external circuit.

When the axle attached to the two rings is rotated such that the arm AB moves up, the arm CD moves down in the magnetic field produced by the permanent magnet. Let us say the coil ABCD is rotated clockwise. By applying Fleming's right-hand rule the induced currents are setup in these arms along the directions AB and CD. Thus an induced current flows in the direction ABCD. If there are large numbers of turns in the coil, the current generated in each turn adds up to give a large current through the coil. This means that the current in the external circuit flows from B_1 to B_2 .

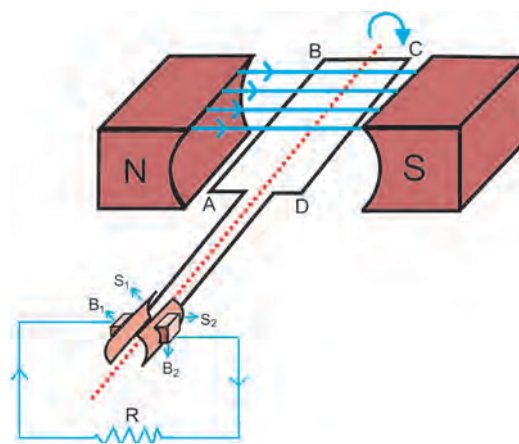
After half a rotation, arm CD starts moving up and AB moving down. As

a result, the directions of the induced currents in both the arms change, giving rise to the net induced current in the direction DCBA. The current in the external circuit now flows from B_1 to B_2 . Thus after every half rotation the polarity of the current in the respective arms changes. Such a current which changes direction after equal intervals of time, is called an alternating current (AC). This device is called an AC generator.



A.C Generator

a



D.C Generator

b

Fig 17.13

To get a direct current (DC), a splitting type commutator must be used with this arrangement, Fig.17.13b, one brush is at all times in contact with the arm moving up in the field, while the other is in contact with the arm moving down. Thus a unidirectional current is produced. The generator is thus called a DC generator.

An important advantage of AC over DC is that electric power can be transmitted over long distances without much loss of energy.

17.7. LIGHT

We see a variety of objects in the world around us. However we are unable to see anything in a dark room. On lighting up the room things becomes visible. What makes things visible? During the day the sunlight helps us to see objects. An object reflects light that falls on it. This reflected light when received by our eyes, enables us to see things.

There are a number of common wonderful phenomena associated with light. In this chapter, we shall study the phenomena of reflection and refraction of light using the straight-line propagation of light.

Reflection of light

A highly polished surface, such as a mirror, reflects most of the light falling on it. You are already familiar with the laws of reflection of light. Let us recall these laws.

- (i) The angle of incidence is equal to the angle of reflection, and
- (ii) The incident ray, the normal to the mirror at the point of incidence and the reflected ray, all lie in the same plane.

These laws of reflection are applicable to all types of reflecting surfaces including spherical surfaces.

Spherical mirrors

ACTIVITY 17.9

- Take a perfect hemispherical spoon. Try to view your face in its curved surface.
- Do you get the image? Is it larger or smaller?
- Move the spoon slowly away from your face. Observe the image. How does it change?
- Reverse the spoon and repeat the activity. How does the image look like now?
- Compare the characteristics of the images on the two surfaces.

The curved surface of a shining spoon could be considered as a curved mirror. The most commonly used type of curved mirror is the spherical mirror. The reflecting surface of a spherical mirror may be curved inwards or outwards. **A spherical mirror whose reflecting surface is curved inwards is called a concave mirror. A spherical mirror whose reflecting surface is curved outwards is called a convex mirror.** The schematic representation of these mirrors is shown in Fig. 17.14.

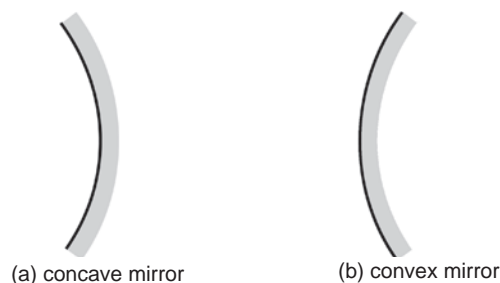


Fig 17.14

You may now understand that the surface of the spoon curved inwards can be approximated to a concave mirror and the surface of the spoon bulged outwards can be approximated to a convex mirror.

Before we move further on spherical mirrors, we need to recognise and understand the meaning of a few terms. These terms are commonly used in discussions about spherical mirrors.

The centre of the reflecting surface of a spherical mirror is a point, called the pole. It is represented by the letter P.

The reflecting surface of a spherical mirror forms a part of a sphere. This sphere has a centre. This point is called the centre of curvature of the spherical mirror. It is represented by the letter C.

The radius of the sphere of which the reflecting surface of a spherical mirror forms a part, is called the radius of curvature of the mirror. It is represented by the letter R.

Imagine a straight line passing through the pole and the centre of curvature of a spherical mirror. This line is called the principal axis.

ACTIVITY 17.10

- Hold a concave mirror in your hand and direct its reflecting surface towards the sun.
- Direct the light reflected by the mirror on to a sheet of paper held close to the mirror.
- Move the sheet of paper back and forth gradually until you find on the paper sheet a bright, sharp spot of light.
- Hold the mirror and the paper in the same position for a few minutes. What do you observe? Why?

Let us understand important terms related to mirrors, through above activity.

The paper at first begins to burn producing smoke. It may even catch fire. Why does it burn? The light from the sun is converged at a point, as a sharp, bright spot by the mirror. In fact, this spot of light is the image of the sun on the sheet of paper. This point is the focus of the concave mirror. The heat produced due to the concentration of the sunlight ignites the paper. The distance of the image from the position of the mirror gives the approximate focal length of the mirror. Observe Fig.17.15(a) closely

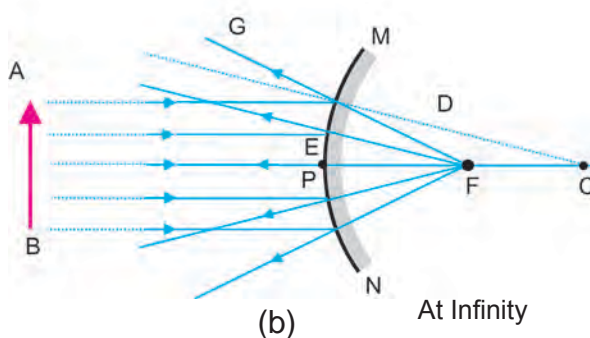
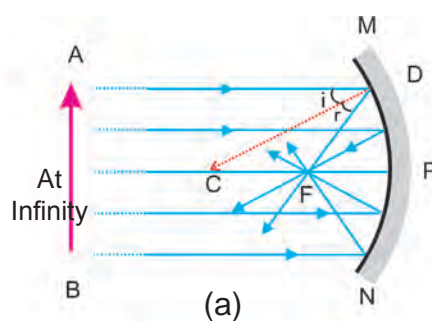


Fig. 17.15

A number of rays parallel to the principal axis are falling on a concave mirror. Observe the reflected rays. They are all meeting at a point on the principal axis of the mirror. This point is called the principal focus of the concave mirror. Similarly

observe Fig. 17.15(b). How are the rays parallel to the principal axis reflected by a convex mirror? The reflected rays appear to come from a point on the principal axis. This point is called the principal focus of the convex mirror. The principal focus is represented by the letter F.

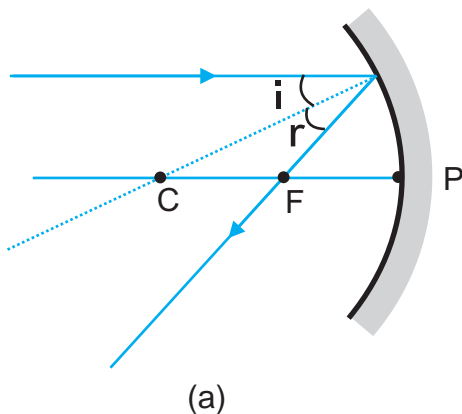
The distance between the pole and the principal focus of a spherical mirror is called the focal length. It is represented by the letter f .

The diameter of the reflecting surface of spherical mirror is called its aperture. In fig 17.15, distance MN represents the aperture. In our discussion we shall consider only such spherical mirrors whose aperture is much smaller than its radius of curvature.

Is there any relationship between the radius of curvature R , and focal length f , of a spherical mirror? For spherical mirrors of small apertures the radius of curvature is found to be equal to twice the focal length. We put this as $R = 2f$.

17.7.1 Reflection of light by spherical mirror

The reflection of light by a spherical mirror takes place according to certain definite rules as follows.



(i) A ray parallel to the principal axis, after reflection, will pass through principal focus in case of a concave mirror or appear to diverge from the principal focus in case of a convex mirror. This is illustrated in Fig. 17.16(a) and (b).

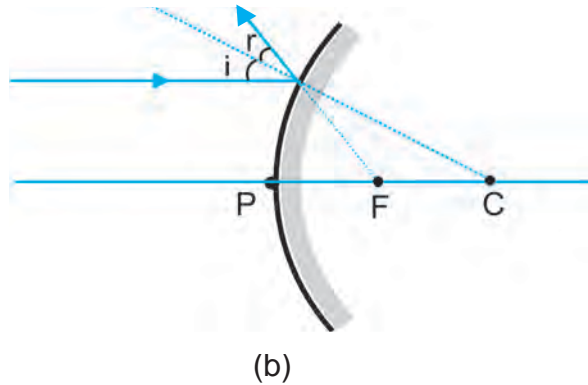
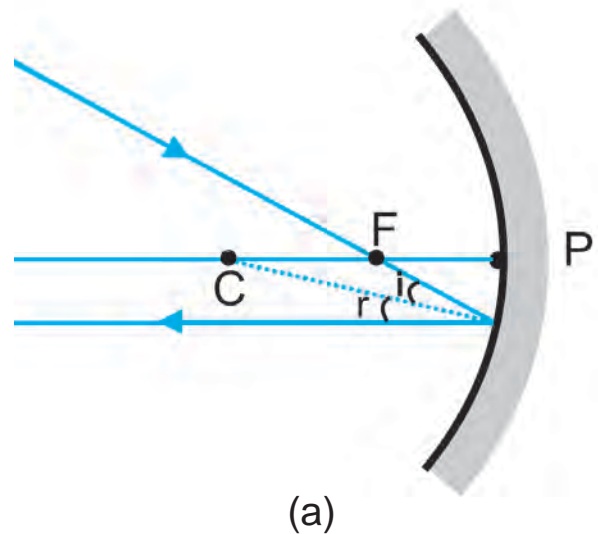
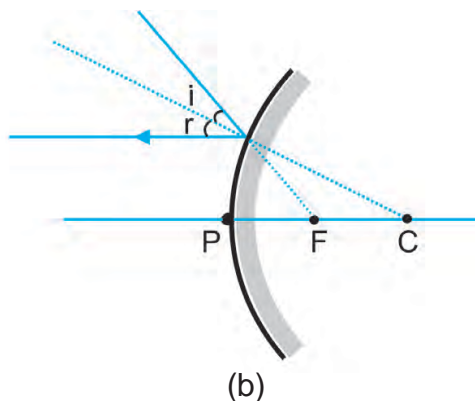


Fig. 17.16

(ii) A ray passing through the principal focus of a concave mirror or a ray directed towards the principal focus of a convex mirror, after reflection, will emerge parallel to the principal axis. This is illustrated in Fig.17.17 (a) and (b).

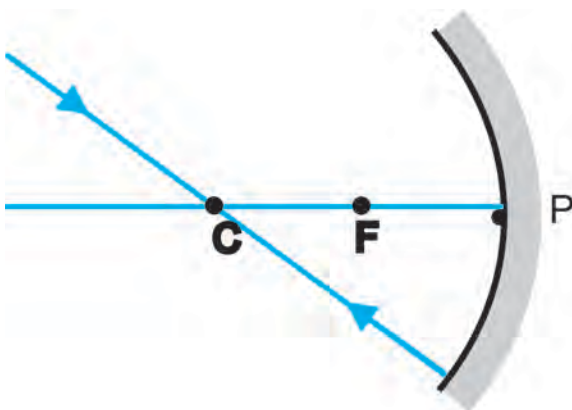




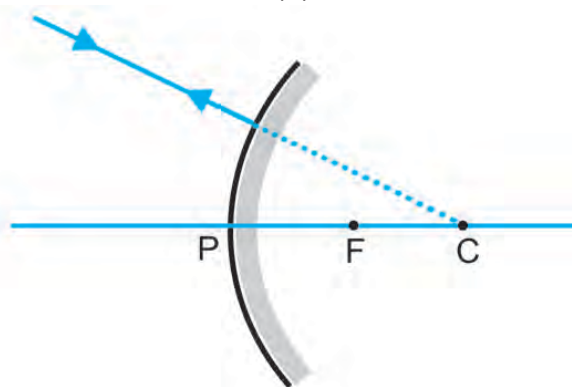
(b)

Fig. 17.17

(iii) A ray passing through the centre of curvature of a concave mirror or directed in the direction of the centre of curvature of a convex mirror, after reflection, is reflected back along the same path. This is illustrated in Fig.17.18 (a) and (b).



(a)



(b)

Fig 17.18

Image formation by concave mirror

How about the images formed by spherical mirrors? How can we locate the image formed by a concave mirror for different positions of the object? Are the images real or virtual? Are the images enlarged, diminished or have the same size?

The nature, position and size of the image formed by a concave mirror depend on the position of the object in relation to point P, F and C. The image formed is real for some positions of the object. It is found to be a virtual image for a certain other position. The image is either magnified, reduced or has the same size, depending on the position of the object.

We can study the formation of image by spherical mirrors by drawing ray diagrams. To construct the ray diagrams, it is more convenient to consider only two rays. These rays are so chosen that it is easy to know their directions after reflection from the mirror. You may take any two of the rays mentioned in the previous section for locating the image. The intersections of the two reflected rays give the position of image of the point object. This is illustrated in the Fig.17.19.

Uses of concave mirror

Concave mirrors are commonly used in torches, search-lights and vehicles head lights to get powerful parallel beams of light. They are used as shaving mirrors to see a large image of the face. The dentists use concave mirrors to see large images of the teeth of patients. Large concave mirrors are used to concentrate sun light to produce heat in solar furnaces.

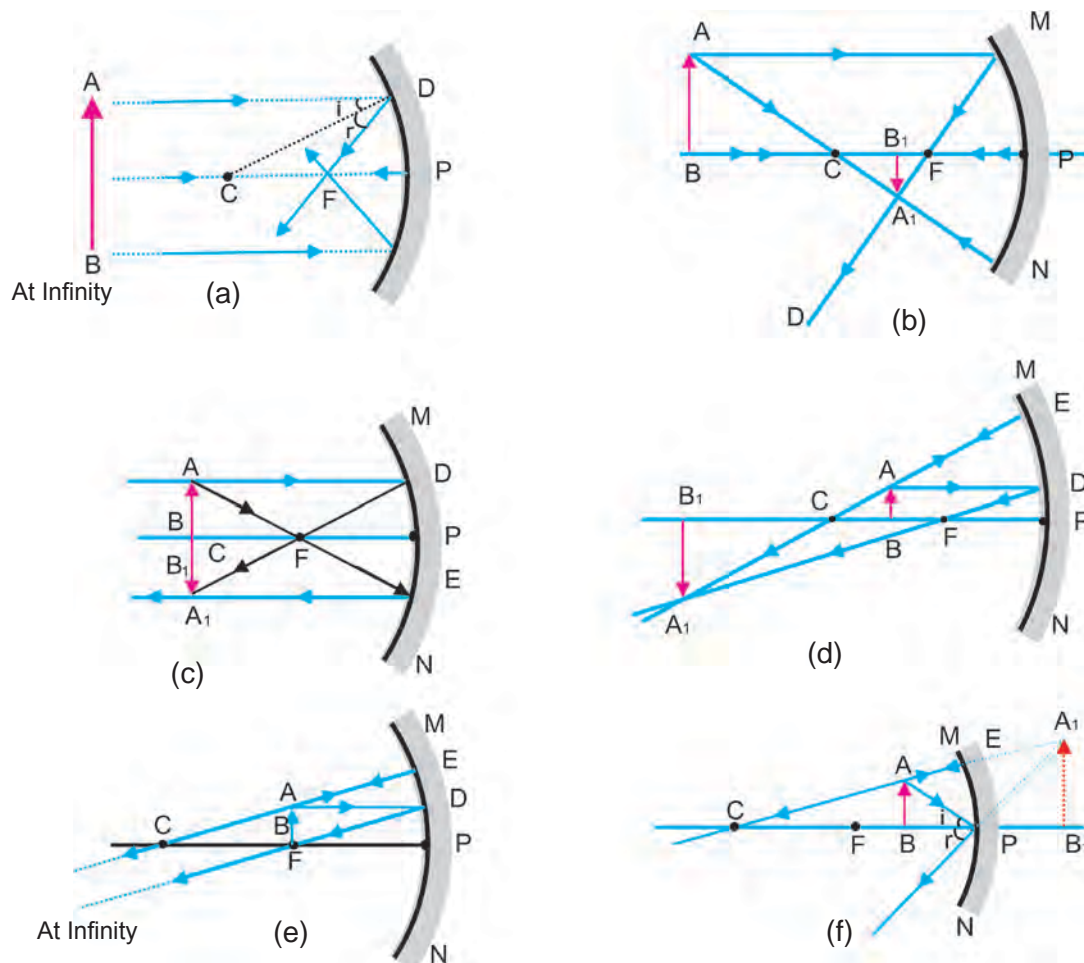


Fig 17.19

A summary of these observations is given in Table: 17.1.

Position of the Object	Position of the image	Relative size of the image	Nature of the image
At infinity	At focus F	Highly diminished, point-sized	Real and inverted
Beyond C	Between F and C	Diminished	Real and inverted
At C	At C	Same size	Real and inverted
Between C & F	Beyond C	Enlarged	Real and inverted
At focus F	At infinity	Highly enlarged	Real and inverted
Between P and F	Behind the Mirror	Enlarged	Virtual and erect

Table 17.1

Image formation by a convex mirror

We consider two positions of the object for studying the image formed by a convex mirror. First when the object is at infinity and the second position is when the object is at a finite distance from the mirror. The ray diagrams for the formation of image by a convex mirror for these two positions of the object are shown in Fig 17.20(a) and (b), respectively.

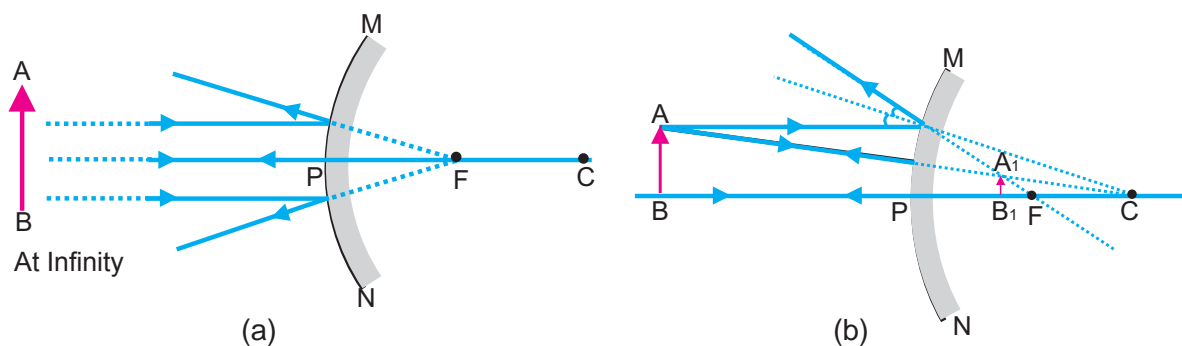


Fig. 17.20

A summary of these observations is given in Table: 17. 2

Position of the object	Position of the image	Relative size of the image	Nature of the image
At infinity	At focus F behind the Mirror	Highly diminished, point-sized	Virtual and erect
Between infinity and Pole P of the Mirror	Between P and F behind the Mirror	Diminished	Virtual and erect

Table 17.2

You have studied the image formation by a concave mirror and a convex mirror, which of these mirrors will give the full image of a large object? Let us explore through an activity.

ACTIVITY 17.11

- Observe the image of a distant tree in a concave mirror.
- Could you see a full length image?
- Repeat this Activity with a convex mirror. Did the mirror show full length image of the object?
- Explain your observations with reason.

You can see a full length image of a tree in a small convex mirror.

Uses of convex mirrors

Convex mirrors are commonly used as rear-view mirrors in vehicles. These mirrors are fitted on the sides of the vehicle, enabling the driver to see traffic behind him/her to facilitate safe driving. Convex mirrors are preferred because they always give an erect image. Also they have a wider field of view as they are curved outwards.

Sign convention for reflection by spherical mirrors

While dealing with the reflection of light by spherical mirrors, we shall follow a set of sign conventions called the **New Cartesian Sign Convention**. In this convention, the pole (P) of the mirror is taken as the origin. The principal axis of the mirror is taken as the X axis (X'X) of the coordinate system. The conventions are as follows.

- (i) The object is always placed to the left of the mirror.
- (ii) All distances parallel to the principal axis are measured from the pole of the mirror.
- (iii) All the distances measured to the right of the origin (along +X-axis) are taken as positive while those measured to the left of the origin (along -X-axis) are taken as negative
- (iv) Distances measured perpendicular to and above the principal axis (along +Y-axis) are taken as positive.
- (v) Distances measured perpendicular to and below the principal axis (along -Y-axis) are taken as negative.

The New Cartesian Sign Convention described above is illustrated in Fig. 17.21.

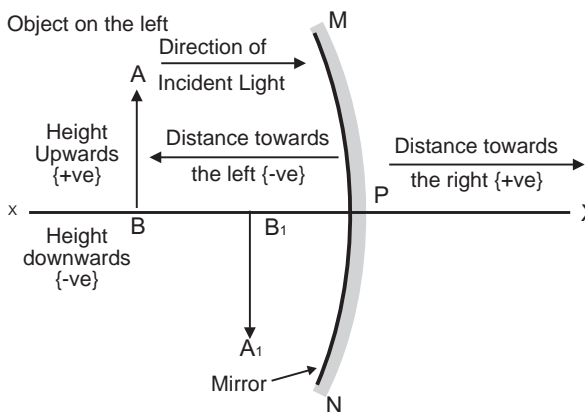


Fig. 17.21

These sign conventions are applied to obtain the mirror formula

Mirror formula

In a spherical mirror, the distance of the object from its pole is called the object distance (u). The distance of the image from the pole of the mirror is called the image distance (v). You already know that the distance of the principal focus from the pole is called the focal length (f). There is a relationship between these three quantities given by the mirror formula which is expressed as

$$1/v + 1/u = 1/f$$

This formula is valid in all situations for all spherical mirrors for all positions of the object. You must use the New Cartesian Sign convention while substituting numerical values for u , v , f , and R in the mirror formula for solving problems.

Example: 17.1

A convex mirror used for rear-view on an automobile has a radius of curvature of 3 m. If a bus is located at 5 m from this

mirror, find the position and nature of the image.

Solution:

Radius of curvature, $R = +3.00 \text{ m}$

Object distance $u = -5.00 \text{ m}$

Image distance $v = ?$

We know,

$$\frac{1}{v} + \frac{1}{u} = \frac{1}{f}$$

or,

$$\begin{aligned} \frac{1}{v} &= \frac{1}{f} - \frac{1}{u} \\ &= \frac{1}{1.5} - \frac{1}{-5.00} = \frac{1}{1.5} + \frac{1}{5.00} \\ &= \frac{5.00 + 1.50}{6.50} \\ &= \frac{7.50}{6.50} = \frac{7.50}{7.50} \\ v &= \frac{7.50}{6.50} = 1.15 \text{ m} \end{aligned}$$

The image is 1.15 m at the back of the mirror. The image is virtual.

17.7.2. Refraction of light

Light seems to travel along straight-line paths in a transparent medium. What happens when light enters from one transparent medium to another? Does it still move along a straight-line path or change its direction? We shall recall some of our day-to-day experiences.

You might have observed that the bottom of a tank or a pond containing water appears to be raised. Similarly, when a thick glass slab is placed over some printed matter, the letters appear raised when viewed through the glass slab. Why does it happen? Have you seen a pencil partially immersed in water in a glass tumbler? It appears to be displaced at the interface of air and water. You might have observed that a lemon kept in water in a glass tumbler appears to be bigger than its actual size, when viewed from the sides. How can you account such experiences?

Let us consider the case of the apparent displacement of the pencil, partly immersed in water. The light reaching you from the portion of the pencil inside water seems to come from a different direction, compared to the part above water. This makes the pencil appear to be displaced at the interface. For similar reasons, the letters appear to be raised, when seen through a glass slab placed over it.

Does a pencil appear to be displaced to the same extent, if instead of water, we use liquids like kerosene or turpentine? Will the letters appear to rise to the same height if we replace a glass slab with a transparent plastic slab? You will find that the extent of the effect is different for different pair of media. These observations indicate that light does not travel in the same direction in all media. It appears that when travelling obliquely from one medium to another, the direction of propagation of light in the second medium changes. This phenomenon is known as refraction of light. Let us understand this phenomenon further by doing an activity.

ACTIVITY 17.12

- Place a coin at the bottom of a bucket filled with water.
- With your eye to a side above water, try to pick up the coin in one go. Did you succeed in picking up the coin?
- Repeat the Activity. Why did you not succeed in doing it in one go?
- Ask your friends to do this. Compare your experience with theirs.

The apparent position of the coin as seen through water differ from its actual position .

Laws of refraction

Refraction of light is due to change in the speed of light as it enters from one transparent medium to another. Experiments show that refraction of light occurs according to certain laws. The following are the laws of refraction of light.

- (i) **The incident ray, the refracted ray and the normal to the interface of two transparent media at the point of incidence, all lie in the same plane.**
- (ii) **The ratio of sine of angle of incidence to the sine of angle of refraction is a constant, for the light of a given colour and for the given pair of media. This law is also known as Snell's law of refraction.**
- If i is the angle of incidence and r is the angle of refraction, then,

$$\sin i / \sin r = \text{constant}$$

This constant value is called the refractive index of the second medium with respect to the first.

17.7.3 Refractive index

We know that a ray of light travels obliquely from one transparent medium into another will change its direction in the second medium. The extent of the change in direction that takes place in a given pair of media is expressed in terms of the refractive index of the second medium with respect to the first medium.

The refractive index can be linked to the relative speed of propagation of light in different media. Light propagates with different speeds in different media. It travels the fastest in vacuum with the highest speed of $3 \times 10^8 \text{ m s}^{-1}$. Its speed reduces considerably in glass.

Consider a ray of light travelling from medium 1 into medium 2 as in Fig.17.22. Let i, r be the angle of incidence and angle of refraction.

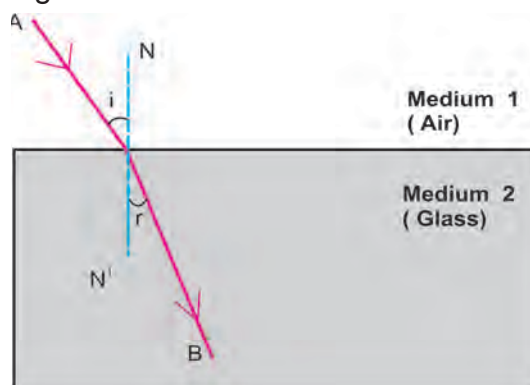


Fig. 17.22

The refractive index of the second medium with respect to the first

$$\mu = \frac{\sin i}{\sin r}$$

$$\mu = \frac{\text{Speed of light in air}}{\text{Speed of light in medium}}$$

17.7.4 Refraction by spherical lenses

Spherical lenses

You might have seen people using spectacles for reading. The watchmakers use a small magnifying glass to see tiny parts. Have you ever touched the surface of a magnifying glass with your hand? Is it plane surface or curved? Is it thicker in the middle or at the edges? The glasses used in spectacles and that by watchmaker are examples of lenses. What is a lens? How does it bend light rays? Let us discuss in this section.

A transparent material bound by two surfaces, of which one or both surfaces are spherical, forms a lens. This means that a lens is bound by at least one spherical surface. In such spherical lenses, the other surface would be plane. A lens may have two spherical surfaces, bulging outwards. Such a lens is called a double convex lens. It is simply called a convex lens. It is thicker at the middle as compared to the edges. Convex lens converges light rays. Hence it is called converging lens. Similarly, a double concave lens is bounded by two spherical surfaces, curved inwards. It is thicker at the edges than at the middle. Such lenses diverge light rays and are called diverging lenses. A double concave lens is simply called a concave lens.

Let us understand the meaning of a few terms which are commonly used in discussions about spherical lenses. A lens has two spherical surfaces. Each of these surfaces forms a part of a sphere. The centres of these spheres are called **centres of curvature of the lens**. The

centre of curvature of a lens is usually represented by the letter C. Since there are two centres of curvature, we may represent them as C_1 and C_2 .

An imaginary straight line passing through the two centres of the curvature of a lens is called its **principal axis**.

The central point of a lens is called its **optical centre**. It is represented by the letter O. A ray of light through the optical centre of a lens passes without suffering any deviation.

The effective diameter of the circular outline of a spherical lens is called its **aperture**. Lenses whose aperture is much less than its radius of curvature are called thin lenses with small aperture. What happens when parallel rays of light are incident on a lens?

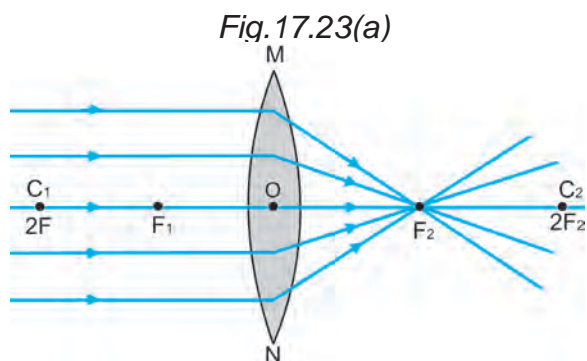
ACTIVITY 17.13

- **CAUTION:** Do not look at the sun directly or through a lens while doing this Activity or otherwise. You may damage your eyes if you do so.
- Hold a convex lens in your hand. Direct it towards the sun.
- Focus the light from the sun on a sheet of paper. Obtain a sharp bright image of the sun.
- Hold the paper and the lens in the same position for a while. Keep observing the paper. What happened? Why?

The light from the sun constitutes parallel rays. These rays were converged by the lens as a sharp bright spot. This is the real image of the sun. The

concentration of the sun light at this spot generated heat. This caused the paper to burn.

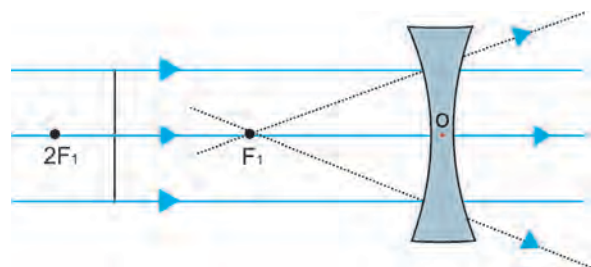
Observe Fig.17.23(a) carefully.



Several rays of light parallel to the principal axis are falling on a convex lens. These rays after refraction from the lens are converging to a point on the principal axis. This point is called the **principal focus** of the lens.

Observe Fig. 17.23(b) carefully,

Fig.17.23(b)



Several rays of light parallel to the principal axis are falling on a concave lens. These rays after refraction from the lens, are appearing to diverge from a point on the principal axis. This point is called the **principal focus** of the concave lens.

If you pass parallel rays from the opposite surface of the lens, you will get another principal focus on the opposite side. Letter F is usually used to represent principal focus. However, a lens has two

principal foci. They are represented by F_1 and F_2 .

The distance of the principal focus from the optical centre of a lens is called its **focal length**. The letter f is used to represent the focal length.

17.7.5 Image formation by lenses

We can represent image formation by lenses using ray diagrams. Ray diagrams will also help us to study the nature, position and relative size of the image formed by the lenses. For drawing ray diagrams in lenses, we consider any two of the following rays.

(i) A ray of light from the object, parallel to the principal axis, after refraction from a convex lens, passes through the principal focus on the other side of the lens, as shown in Fig.17.24(a). In case of a concave lens, the ray appears to diverge from the principal focus located on the same side of the lens, as shown in Fig.17.24(b)

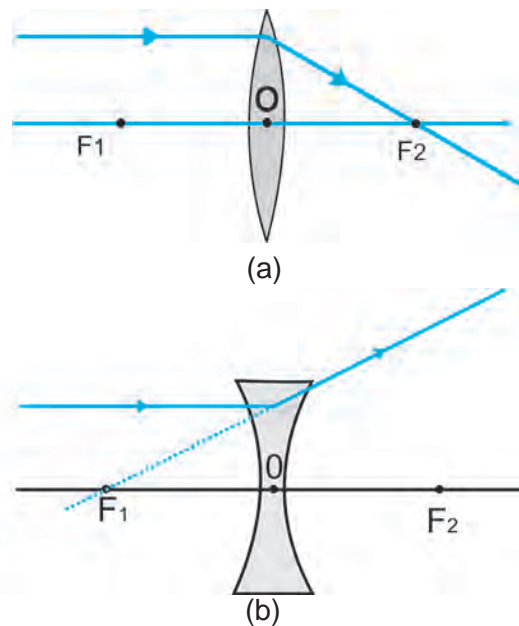


Fig. 17.24

(ii) A ray of light passing through a principal focus after refraction from a convex lens will emerge parallel to the principal axis. This is shown in Fig 17.25(a). A ray of light appearing to meet at the principal focus of a concave lens, after refraction, will emerge parallel to the principal axis. This is shown in Fig. 17.25(b).

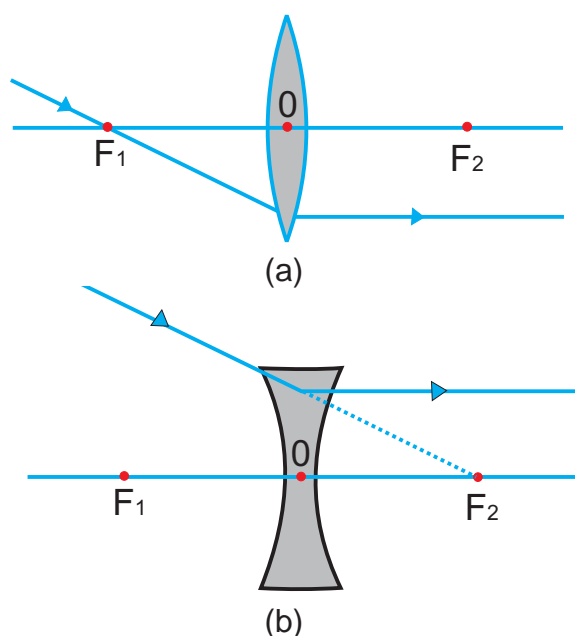


Fig. 17.25

(iii) A ray of light passing through the optical centre of a lens will emerge without any deviation. This is illustrated in Fig 17.26(a) and (b).

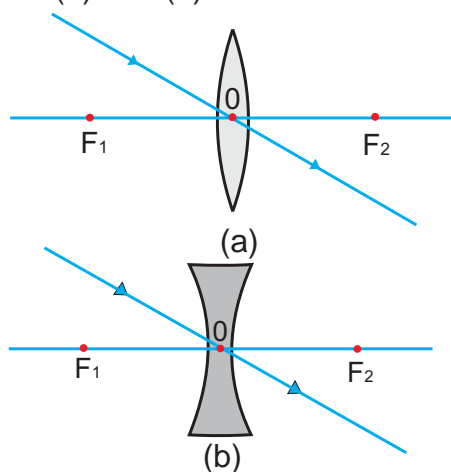


Fig. 17.26

Sign convention for spherical lenses:

All measurements are taken from the optical centre of the lens. According to the convention, the focal length of a convex lens is positive and that of a concave lens is negative. We must take care to apply appropriate signs for the values of u , v , f , object height h and image height h' .

17.7.6 Lens formula

This formula gives the relation between object-distance (u), image-distance (v) and the focal length (f). The lens formula is expressed as

$$\frac{1}{f} = \frac{1}{v} - \frac{1}{u}$$

The lens formula given above is general and is valid in all situations for any spherical lenses.

Example: 17.2

A concave lens has focal length of 15 cm. At what distance should the object from the lens be placed so that it forms an image 10 cm from the lens?

Solution:

$$v = -10 \text{ cm}, \quad f = -15 \text{ cm}, \quad u = ?$$

$$\frac{1}{v} - \frac{1}{u} = \frac{1}{f} \quad \text{Or,}$$

$$\frac{1}{u} = \frac{1}{v} - \frac{1}{f}$$

$$\frac{1}{u} = \frac{1}{-10} - \frac{1}{-15}$$

$$\frac{1}{u} = \frac{-3 + 2}{30} = \frac{-1}{30}$$

$$u = -30 \text{ cm}$$

Thus, the object distance is 30 cm.

The ray diagrams for the image formation in a convex lens for a few positions of the object are shown in Fig. 17.27.

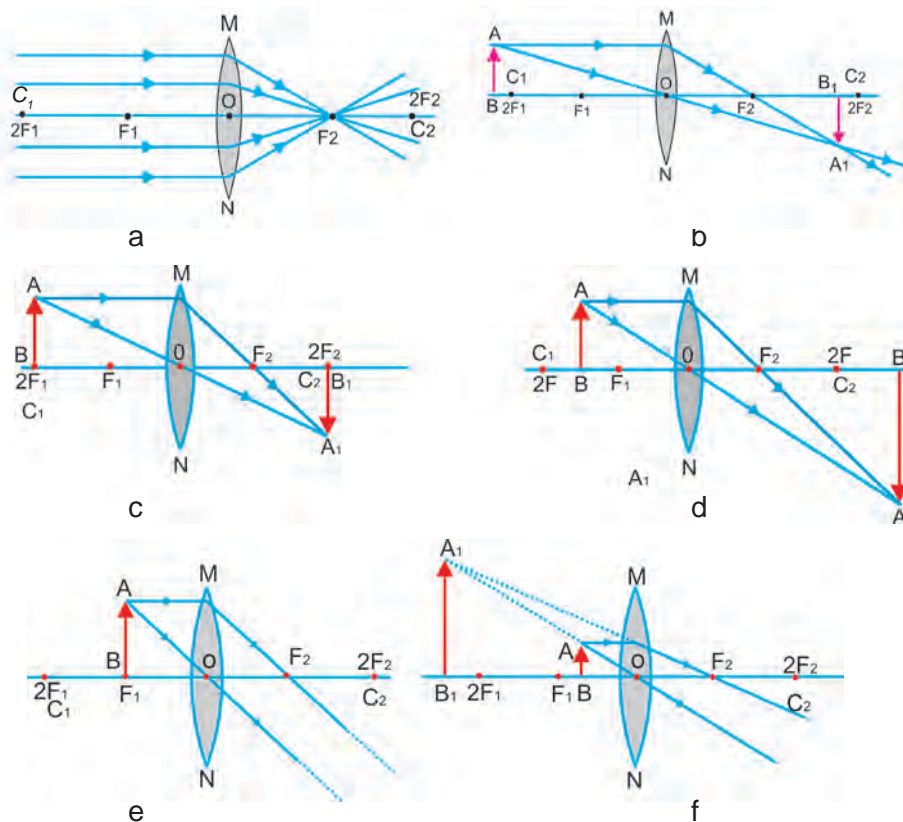


Fig. 17.27

A summary of these observations is given in Table 17.3.

Position of the object	Position of the image	Relative size of the image	Nature of the image
At infinity	At focus F_2	Highly diminished, point-sized	Real and inverted
Beyond $2F_1$	Between F_2 and $2F_2$	Diminished	Real and inverted
At $2F_1$	At $2F_2$	Same size	Real and inverted
Between F_1 and $2F_1$	Beyond $2F_2$	Enlarged	Real and inverted
At focus F_1	At infinity	Infinitely large or highly enlarged	Real and inverted
Between focus F_1 and optical centre O	On the Same side of the lens as the object	Enlarged	Virtual and erect

Table 17.3

The ray diagrams for the image formation in a concave lens for various positions of the object are shown in Fig. 17.28.

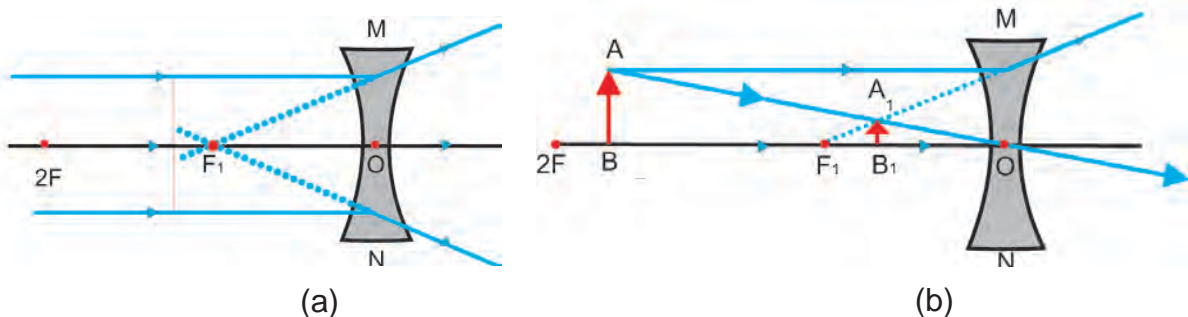


Fig. 17.28

A summary of these observations is given in Table. 17.4.

Position of the object	Position of the image	Relative size of the image	Nature of the image
At infinity	At focus F_1	Highly diminished, point-sized	Virtual and erect
Between infinity and optical centre O of the lens	Between focus F_1 and optical centre O	Diminished	Virtual and erect

Table 17.4

Magnification

The magnification produced by a lens is defined as the ratio of the height of the image to the height of the object

It is represented by the letter m . If h is the height of the object and h' is the height of the image given by the lens, then the magnification produced by the lens is given by,

$$m = \frac{\text{Height of the image (} h' \text{)}}{\text{Height of the object (} h \text{)}} = \frac{v}{u}$$

Example: 17.3

An object is placed at a distance of 30 cm from a concave lens of focal length 15 cm. An erect and virtual image is formed at a distance of 10 cm from the lens. Calculate the magnification.

Solution:

Object distance, $u = -30$ cm

Image distance, $v = -10$ cm

Magnification, $m = v/u$

$$m = \frac{-10 \text{ cm}}{-30 \text{ cm}} = \frac{1}{3} = +0.33$$

17.7.7. Power of lens

The degree of convergence or divergence of light rays achieved by a lens is expressed in terms of its power. **The power of a lens is defined as the reciprocal of its focal length.** It is represented by the letter P . The power P of a lens of focal length f is given by

$$P = \frac{1}{f}$$

The SI unit of power of a lens is 'dioptre'. It is denoted by the letter D . If f is expressed in meter, then, power is expressed in dioptries. Thus 1 dioptre is the power of a lens whose focal length is 1 meter. The power of a convex lens is positive and that of a concave lens is negative.

Example: 17.4

The focal length of a concave lens is 2m. Calculate the power of the lens.

Solution:

Focal length of concave lens, $f = -2$ m
Power of the lens,

$$p = \frac{1}{f}$$

$$p = \frac{1}{-2\text{m}}$$

$$p = -0.5 \text{ dioptre}$$

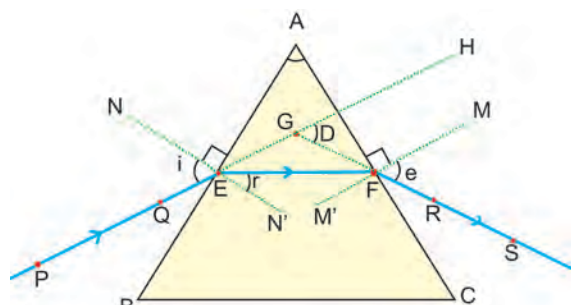
17.7.8. Refraction of light through a prism

Consider a triangular glass prism. It has two triangular bases and three rectangular lateral surfaces. These surfaces are inclined to each other. The angle between its lateral faces is called the angle of the prism. Let us now do an activity to study

the refraction of light through a triangular glass prism.

ACTIVITY 17.14

- Fix a sheet of white paper on a drawing board using drawing pins.
- Place a glass prism on it in such a way that it rests on its triangular base. Trace the outline of the prism using a pencil.
- Draw a straight line PE inclined to one of the refracting surfaces, say AB , of the prism.
- Fix two pins, say at points P and Q , on the line PE as shown in Fig 17.29
- Look for the images of the pins, fixed at P and Q , through the other face AC .
- Fix two more pins, at points R and S , such that the pins at R and S lie on the same straight line.
- Remove the pins and the glass prism.
- The line PE meets the boundary of the prism at point E (see Fig 17.29). Similarly, join and produce the points R and S . Let these lines meet the boundary of the prism at E and F , respectively. Join E and F .
- Draw perpendicular to the refracting surfaces AB and AC of the prism at points E and F , respectively.
- Mark the angle of incidence ($\angle i$), the angle of refraction ($\angle r$) and the angle of emergence ($\angle e$) as shown in Fig 17.29.



PE - Incident ray $\angle i$ - Angle of incident
 EF - Refracted ray $\angle r$ - Angle of refraction
 FS - Emergent ray $\angle e$ - Angle of emergence
 $\angle A$ - Angle of the Prism $\angle D$ - Angle of deviation

Fig.17.29

Here PE is the incident ray. EF is the refracted ray. FS is the emergent ray. You may note that a ray of light is entering from air to glass at the first surface AB. The light ray on refraction has bent towards the normal. At the second surface AC, the light ray has entered from glass to air. Hence it has bent away from normal. Compare the angle of incidence and angle of refraction at each refracting surface of the prism. The peculiar shape of prism makes the emergent ray bent at an angle to the direction of the incident ray. This angle $\angle r$ is called the angle of refraction. In this case $\angle D$ is the angle of deviation. Mark the angle of deviation in the above activity and measure it.

17.7.9. Dispersion of white light by a glass prism

You must have seen and appreciated the spectacular colours in a rainbow. How could the white light of the sun give us various colours of the rainbow? The prism has probably split the incident white light into a band of colours. Note the colours that appear at the two ends of the colour band. What is the sequence of

ACTIVITY 17.15

- Take a thick sheet of cardboard and make a small hole in its middle.
- Allow sunlight to fall on the narrow slit. This gives a narrow beam of white light.
- Now, take a glass prism and allow the light from the slit to fall on one of its faces.
- Turn the prism slowly until the light that comes out of it appear on a near by screen.
- What do you observe? You will find a beautiful band of colours.
- Why does this happen?

colours that you see on the screen? The various colours seen are Violet, Indigo, Blue, Green, Yellow, Orange and Red. As shown in Fig.17.30.

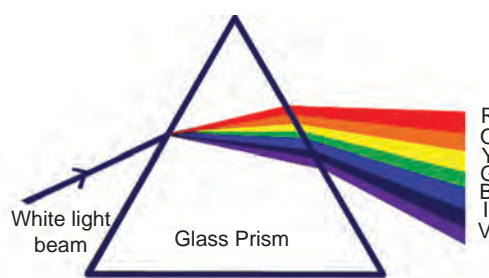


Fig. 17.30

The acronym **VIBGYOR** will help you to remember the sequence of colours.

The band of the coloured component of a light beam is called its spectrum. You might not be able to see all the colours separately. Yet something makes each colour distinct from the other. The splitting of light into its component colours is called dispersion.

You have seen that white light is dispersed into its seven-colour components by a prism. Why do we get these colours? Different colours of light bend through different angles with respect to the incident ray as they pass through the prism. The red light bends the least while the violet the most. Thus the rays of each colour emerge along different paths and thus become distinct. It is the band of distinct colours that we see in a spectrum.

17.7.10. Atmospheric refraction

You might have observed the apparent random wavering or flickering of objects seen through a turbulent stream of hot air rising above a fire. The air just above the fire becomes hotter than the air further up. The hotter air is lighter (less dense) than the cooler air above it, and has a refractive index slightly less than that of the cooler air. Since the physical conditions of the refracting medium (air) are not stationary, the apparent position of the object, as seen through the hot air fluctuates. This wavering is thus an effect of atmospheric refraction (refraction of light by the earth's atmosphere) on a small scale in our local environment. The twinkling of stars is a similar phenomenon on a much larger scale.

17.7.11. Human eye

The human eye is one of the most valuable and sensitive sense organs. It enables us to see the wonderful worlds and colours around us. Of all our sense organs, the human eye is the most significant one as it enables us to see the beautiful, colourful world around us.

The human eye is like a camera. Its lens system forms an image on a light-sensitive screen called the **retina**. Light

enters the eye through the thin membrane called the **cornea**. It forms the transparent bulge on the front surface of the eye ball as shown in Fig. 17.31.

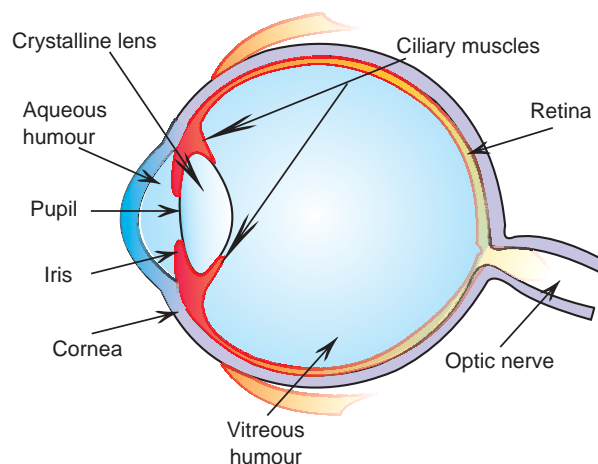


Fig 17.31

The eye ball is approximately spherical in shape with a diameter of about 2.3cm. Most of the refraction for the light rays entering the eye occurs at the outer surface of the cornea. The crystalline lens merely provides the finer adjustment of focal length required to focus objects at different distances on the **retina**. We find a structure called **iris** behind the cornea. Iris is a dark muscular diaphragm that controls the **pupil**. The pupil regulates and controls the amount of light entering the eye. The eye lens forms an inverted real image of the object on the retina. The retina is a delicate membrane having enormous number of **light-sensitive cells**. The light sensitive cells get activated upon illumination and generate **electrical signals**. These signals are sent to the brain via the **optic nerves**. The brain interprets these signals, and finally, processes the information so that we perceive objects as they are.

Defects of vision and rectification

There are mainly three common refractive defects of vision. These are (i) **Myopia** or near - sightedness, (ii) **Hypermetropia** or far-sightedness, and (iii) **Presbyopia**. These defects can be corrected by the use of suitable spherical lenses.

(a) Myopia

Myopia is also known as near-sightedness. A person with myopia can see near by objects clearly but cannot see the distant objects distinctly. A person with this defect has the far point nearer than infinity. Such a person may see clearly up to a distance of a few meters. In a myopic eye, the image of a distant object is formed in front of the retina [Fig. 17.32(a)] and not at the retina itself.

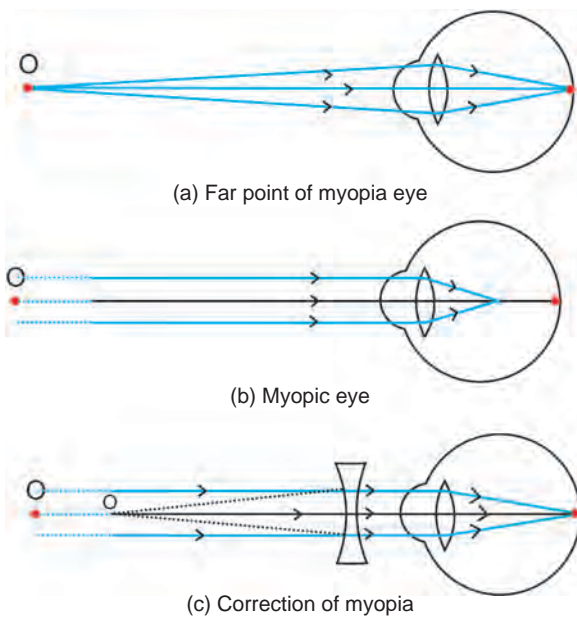


Fig. 17.32

This defect may arise due to (i) excessive curvature of the eye lens, or (ii) elongation of the eyeball. This defect can be corrected by using a concave lens of suitable power. This is illustrated in Fig.17.32(c). A concave lens of suitable power will bring the image back on to the retina and thus the defect is corrected.

(b) Hypermetropia

Hypermetropia is also known as far-sightedness. A person with hypermetropia can see distant objects clearly but cannot see near by objects distinctly. The near point, for the person, is further away from the normal near point (25 cm). Such a person has to keep a reading material such beyond 25cm from the eye for comfortable reading. This is because the light rays from a close by object are focussed at a point behind the retina as shown in Fig.17.33 (b)

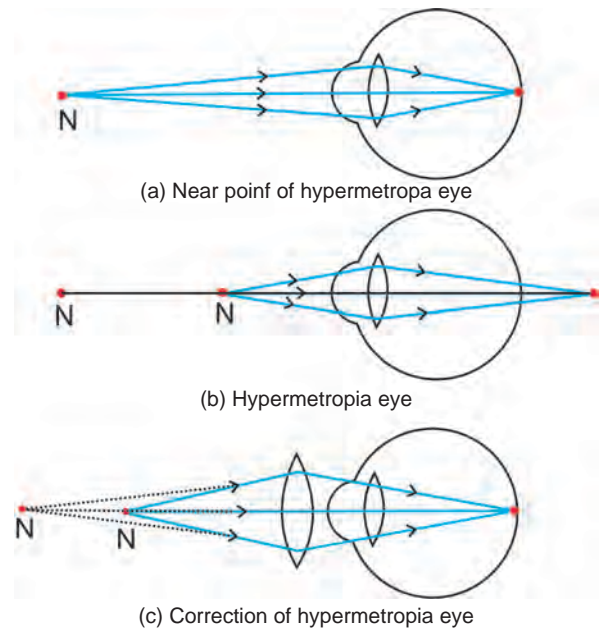


Fig. 17.33

This defect either because (i) the focal length of the eye lens is too long or (ii) the eyeball has become too small. This defect can be corrected by using a convex lens of appropriate power. This is illustrated in Fig.17.33(c). Eye-glasses with converging lenses provide the additional focusing power required for forming the image on the retina.

(c) Presbyopia

The power of accommodation of the eye usually decreases with ageing. For most people, the near point gradually recedes away. They find it difficult to see near by objects comfortably and distinctly without corrective eye-glasses. This defect is called **Presbyopia**. It arises due to the gradual weakening of the **ciliary muscles** and diminishing flexibility of the eye lens. Sometimes, a person may suffer from both myopia and hypermetropia. Such people often require by-focal lenses. A common type of by-focal lenses consists of both concave and convex lenses. The upper portion consists of a concave lens. It facilitates distant vision. These days, it is possible to correct the refractive defects with contact lenses.

17.12. Science today - Hubble space telescope (H.S.T)

Hubble telescope is a space telescope that was carried into orbit by a space shuttle in April 1990. It is named after the American astronomer Edwin Hubble. It becomes a most popular research tool for astronomy. The H.S.T is collaboration between NASA and the European Space Agency, and is one of NASA's great observatories.

Hubble is the only telescope ever designed to be serviced in space by astronauts. The H.S.T design with two hyperbolic mirrors is known for good imaging performance over a wide field of view. During the launch scientist found that the main mirror had been ground incorrectly, which severely affect the telescopes capabilities. After a servicing mission in 1993, the telescope was restored to its intended quality. Four servicing missions where performed from 1993-2002. But the fifth was completed in 2009. The telescope is now expected to function until at least 2014.



Fig.17.34

Hubble's orbit outside the distortion of earth's atmosphere allows it to take extremely sharp images with almost no background light. Hubble's Ultra Deep Field image is the most detailed visible-light image ever made of the universe's most distant object. Hubble Deep field and Hubble ultra Deep field images reveals that galaxies are billions of light years away.

Many Hubble observations accurately measure the rate at which the universe is expanding. It constrain the value of

Hubble's constant and estimates the age of the Universe.

Hubble's images of planets were crucial in studying the dynamics of the collision of a comet with Jupiter, an event believed to occur once every few centuries.

Hubble's observations found that black holes are common to the centers of all galaxies.

The astronomers used the telescope to observe distant supernovae.

EVALUATION

PART A

- The magnification produced by a mirror is $\frac{1}{3}$, then the type of mirror is _____
(concave, convex, plane)
- An electric current through a metallic conductor produces _____ around it.
(heat, light, magnetic field, mechanical force)
- The field of view is maximum for _____
(plane mirror, concave mirror, convex mirror)
- An object is placed 25 cm from a convex lens whose focal length is 10 cm. The image distance is _____. (50 cm, 16.66 cm, 6.66 cm, 10 cm)

PART B

- From the following statement write down that which is applicable to a commutator.
 - galvanometer uses commutator for deadbeat
 - transformer uses commutator to step up voltage
 - motor uses commutator to reverse the current

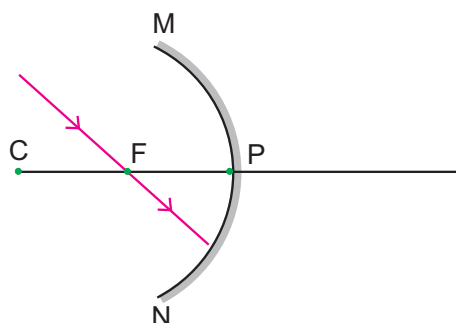
2. Fill in the blanks

- For a motor : a permanent magnet, then commercial motor : _____
- Focal length of a lens; metre, then for power of a lens _____

3. Correct the mistakes, if any, in the following statements.

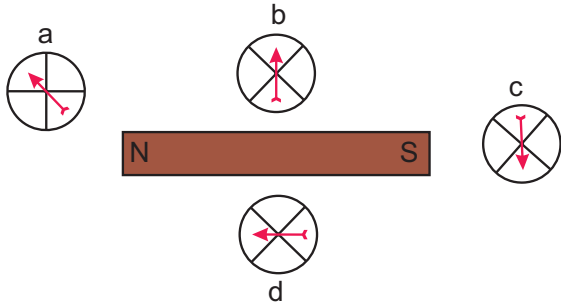
- Magnetic field is a quantity that has magnitude only.

- The magnetic field lines emerge from the south pole and merge at the north pole.
- The ray diagram shown below is introduced to show how a concave mirror forms an image of an object.
 - identify the mistake and draw the correct ray diagram.
 - Write the justifications for your corrections.

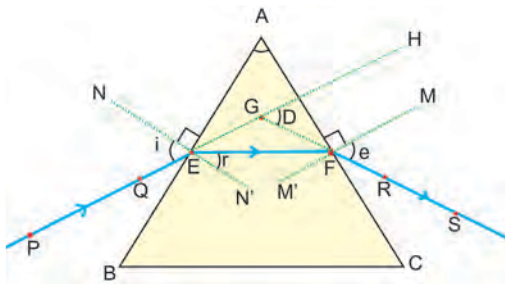


- In traffic signals _____ colour light is used to stop vehicles because it is having _____ wave length.
- Considering this write down the names of the parts in human eye.
 - Dark muscular diaphragm that controls the pupil.
 - The screen at where the image is formed by eye lens.
- You know that myopia is a common refractive defects of vision. Person with this defect can see only nearby objects clearly. Using concave lens of suitable power this defect is corrected.
 - mention other two types of defects like this.

- b) explain how can we correct it.
8. (a) Which of the compass needle orientations in the following diagram might correctly describe the magnet's field at that point?



- (b) To an astronaut sky appears dark instead of blue. Give the reason.

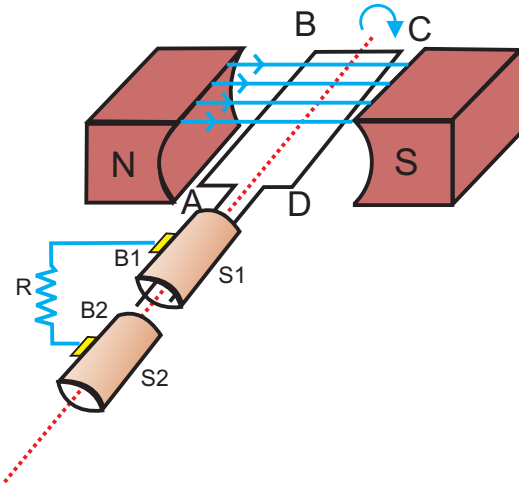


PART – C

1. (a) Label the following in the given diagram given below.

- a) Incident ray b) Refracted ray
c) Emergent ray d) Angle of refraction
e) Angle of deviation f) Angle of emergence

1. (b) The refractive index of diamond is 2.42. What is the meaning of this statement in relation to speed of light?



2. a) Re draw the above diagram.
b) This diagram represents _____
c) Label the parts of the diagram.
d) Write the principle of the name of the device denoted by this diagram.

FURTHER REFERENCE

- Books:** 1. Fundamentals of optics by **D.R. Khanna and H.R. Gulati R.Chand & Co**
2. Magnetism by **Joy Frisch** - Schnoll published by Creative Education.
3. Advanced physics by **Keith Gibbs** Cambridge University press

- Website:** [www.physics about.com](http://www.physics.about.com)
www.opticalsres.com
www.newdn.com
<http://www.khanacademy.org>

SYLLABUS

1. Applied Biology	Heredity and Evolution :- Heredity –Variations-Evolution-Speciation-Human evolution-Evolution tree-Genetic engineering-Bio technology and cloning-Stem cell-Organ culture-Microbial production-Biosensor – Bio chips-Science today – Gene therapy
2. Health and Hygiene	Immune System:- Health and its significance-Diseases and causes-Diseases caused by microbes and prevention-Modes of transmission-Immunization-Treatment and prevention-Biotechnology in Medicine-HIV and Prevention
3. My Body	Structure & Function of the Human Body – Organ System:- Nervous system-Endocrine system-Cell division-Stages of Meiosis-Heredity
4. World of Plants	Reproduction in Plants:- Modes of reproduction - vegetative, asexual and sexual reproduction in plants-Pollination-Fertilization-Fruits and seeds formation-Seed dispersal
5. World of Animals	A Representative Study of Mammals- Morphology-Habitats-Adaptations-Basic Physiological Functions.-Circulatory system in man-Excretory system in man.-Relationship of structure to functions-Animal Behaviour - Behaviour (social, reproductive, parental care) -Some case studies from researchers(animals behavior)
6. Life Process	Life Processes:- Definition-Types of nutrition and human digestive system-Respiration -Transportation in plants-water and minerals and animals - blood circulation-Excretion in plants and animals-Nervous system-Coordination in plants-Movement due to growth-Hormones in animals
7. Environmental Science - Ecology	Conservation of Environment:- Bio degradable and non bio degradable wastes-Water management-Wild life sanctuaries-Balance in Ecosystem-Coal and petroleum-Green chemistry-Science today – Towards a global village
8. Environmental Science – Resource use and Management	Waste Water Management:- Journey of water-Sewage -Treatment -Domestic practices -Sanitation and diseases-Alternate arrangement for sewage disposal -Sanitation in public places-Energy Management-Energy audit (home, school)- Renewable sources (solar, hydrogen, wind)- Non-renewable sources—(coal, petroleum, natural gas)- Bio-fuels-generation & use-Energy Conservation & How we can help.
9. Matter	Solutions:- Solute and Solvent-Types of Solutions-Solubility-Factors affecting – Solubility-Problems
10. Atomic Structure	Atoms and Molecules:- Modern atomic theory-Avogadro Hypothesis-Atomicity-Relation between vapour density and molecular mass of a gas- Difference between-Atom and Molecules-Relative Atomic Mass-Relative Molecular mass-Mole Concepts- Mole- Definition-Problems based on mole concept

11. Exploring Chemical Changes and Formulation	Chemical Reactions:- Types of chemical reactions -Rate of chemical reaction-Factors influencing the rate of the chemical reaction-Acids-Classification of acids- Chemical properties of acids-Uses of acids-Bases-Classification of bases-Chemical properties of bases- uses of bases-Identification of acids and bases-pH scale-pH paper-Importance of pH in everyday life-Salts- Classification of salts-Uses of salts
12. Exploring Chemical Families	Periodic Classification of Elements:- Modern periodic law-Modern periodic table-Characteristics of modern periodic table-Metallurgy –Introduction-Terminologies in metallurgy-Differences between Minerals and Ores-Occurrence of metals- Metallurgy of Al, Cu and Fe- Metallurgy of Aluminium-Metallurgy of Copper- Metallurgy of iron- Alloys- Methods of making alloys-Copper Aluminium and Iron alloys-Corrosion -Method s of preventing corrosion
13. Exploring the World	Carbon and its Compounds:- Introduction-Compounds of carbon-Modern definition of organic chemistry-Bonding in carbon and its compounds-Allotropy- Physical nature of carbon and its compounds- Chemical- properties of carbon compounds-Homologous series-Hydrocarbons and their types -Functional groups- Classification of organic compound based on functional group-Ethanol-Ethanoic acid
14. Matter and Measurement	Measuring Instruments:- Screw Gauge-Measuring long distances – Astronomical distance, light year
15. Forces and Movement	Laws of Motion and Gravitation -Balanced and imbalanced forces-First law of motion-Inertia and mass-Momentum-Second law of motion- $F=ma$ -Third law of motion-Conservation of momentum and proof-Moment of force and couple-Gravitation Newton’s law of gravitation –Mass- Weight-Acceleration due to gravity-Mass of Earth-Science Today- Chandrayan, Cryogenic Techniques and Manned Space Station
16. Exploring Energy	Electricity and Energy:- Electric current and circuit-Electric potential and potential difference-Circuit diagram-Ohm’s law-Resistance of a conductor-System of resistors -Heating effect of electric current-Joules law of heating-Role of fuse-Domestic electric circuits-Electric power-Chemical effect of electric current-Electrolysis electro chemical cells-Primary and Secondary cells-Sources of Energy-Conventional sources of energy-Non- conventional source of energy- Nuclear energy-Radioactivity-Nuclear fission and nuclear fusion-Nuclear reactivity advantages-Hazards of nuclear energy-Science today – Energy from seas.
17. Exploring Phenomena	Magnetic Effect of Electric Current and Light :- Magnetic field and magnetic lines of force-Magnetic field due to current carrying conductor-Magnetic field due to current carrying Straight conductor- Magnetic field due to current carrying Circular loop-Force on a current carrying conductor in a magnetic field-Fleming left hand rule -Electric motor-Electromagnetic induction- Faraday’s experiments-Electric generator –Light-Reflection of light by Spherical mirrors – image formation and Mirror Formula - Refraction – Laws of refraction - Refractive index-Refraction by spherical lenses- Image formation by lenses-Lens formula and magnification-Power of lens-Refraction of light through a prism-Dispersion by a glass prism-Atmospheric refraction- Human eye –Defects and rectification-Science today –Hubble space telescope
18. Technology	Practical and Projects

Design of Question Paper – X Std Science (Theory)

Time: 2½ Hours

Max. Marks: 75

The weightage of marks allotted for the design of question paper shall be as under:

A. Weightage to Learning Outcome

Sl.No	Categories	Mark	PERCENTAGE
1	Knowledge	17	15
2	Understanding	52	45
3	Application	35	30
4	Skill	11	10
	Total	115	100

Note: (1) Total Marks is 115 inclusive of choice. (2) While preparing the question paper, there may be variations in weightage to the extent from 2 % to 5 %.

B. Weightage given to various types of question

S.No	Types of Questions	Marks for Each Question	Total No. of Questions	No. of Questions to be answered	Total Marks
1	Section A Objective Type (OT)	1	15	15	15x 1=15
2	Section B Short Answer (SA)	2	30*	20	20x2 = 40
3	Section C Long Answer (LA)*	5	8	4	4 x 5 = 20
	Total		53	39	75

* Each Question may be split into 2 or 3 sub-divisions carrying 1, 2 or 3 marks. But the questions shall be from each area (Botany, Zoology, Chemistry, Physics). Choices will be internal (Either - or)

*Short Answer split up

Sl.No.	Very Short Answer Type of Questions	To be asked
1	To Match	3
2	To spot the error / mistake in the given statements	3
3	Reason and Assertion	3
4	To Raise questions	3
5	To label the parts in the given diagram	3
6	To copy a diagram & to identify /mark the parts	3
7	To calculate the required value(Problem solving)	3
8	To fill up the blanks (from the given pair of answers)	3
9	To interpret what happens in the given situations	3
10	To find the odd one out	3
	Total Number of Questions given	30
	Total Number of Questions to be answered	20

C. Weightage given to the higher order of questions

SI.No	Estimated higher order of questions	% Percentage
1	Easy	50
2	Average	40
3	Difficult	10

D. Weightage to Content Unit

Units		No. of Questions				Total Marks
		OT	SA	LA		
1. Heredity and Evolution	Botany and Zoology	1(1)	1(2)	1(5)	23	8
2. Immune System		1(1)	1(2)	1(5)		8
3. Structure & Function of the Human Body – Organ System		-	3(2)	-		6
4. Reproduction in Plants		1(1)	1(2)	1(5)		8
5. A representative Study of Mammals		-	3(2)	-		6
6. Life Processes		1(1)	1(2)			3
7. Conservation of Environment		1(1)	1(2)	1(5)		8
8. Waste Water Management		-	3(2)	-		6
9. Solutions	Chemistry	1(1)	2(2)	-	15	5
10. Atoms and Molecules		-	1(2)	1(5)		7
11. Chemical Reaction		1(1)	2(2)			5
12. Periodic Classification of Elements		2(1)	2(2)			6
13. Carbon and its Compounds		1(1)	1(2)	1(5)		8
14. Measurements					-	
15. Laws of Motion and Gravitation	Physics	1(1)	2(2)	1(5)	15	10
16. Electricity and Energy		2(1)	3(2)			8
17. Magnetic Effect of Electric Current and Light		2(1)	3(2)	1(5)		13
Total Number of Questions given		15(15)	30(60)	8(40)	53	115
Total Number of Questions to be answered		15(15)	20(40)	4(20)	39	75

() Indicates the marks

BLUE PRINT

Unit No.	Content Unit	Related Subject	Knowledge			Understanding			Application			Skill			Total No. of Questions	Total Marks
			OT	SA	LA	OT	SA	LA	OT	SA	LA	OT	SA	LA		
1	Heredity and Evolution	Zoo	1(1)				1(2)				1(5)				3	8
2	Immune System	Zoo	1(1)					1(5)				1(2)			3	8
3	Structure & Function of the	Zoo													3	6
	Human Body			1(2)					1(2)				1(2)			
4	Human Body – Organ System	Bot													3	8
	Reproduction in Plants			1(1)				1(5)						1(2)		
5	A Representative Study of Mammals	Zoo		1(2)					1(2)						3	6
6	Life Processes	Bot & Zoo							1(1)						2	3
7	Conservation of Environment	Bot							1(1)						3	8
8	Waste Water Management	Bot													3	6
9	Solutions	Che							1(1)						3	5
10	Atoms and Molecules	Che													2	7
11	Chemical Reaction	Che		1(2)					1(1)						3	5
12	Periodic Classification of Elements	Che	1(1)						1(1)						4	6
13	Carbon and its Compounds	Che							1(1)						3	8
14	Measurements	Phy													-	-
15	Laws of Motion and Gravitation	Phy		1(2)											4	10
16	Electricity and Energy	Phy		1(2)					1(1)						5	8
17	Magnetic Effect of Electric Current and Light	Phy							1(1)						6	13
	Total		4(4)	4(8)	1(5)	8(8)	12(24)	4(20)	2(2)	9(18)	3(15)	1(1)	5(10)	-	53	115

SCIENCE PRACTICALS

S.No	Content
1	To find out the presence of starch in the given food samples of A and B by using iodine solution.
2	To find out the rate of heart beat of human beings by using stethoscope under normal physical conditions.
3	To find out the body temperature by using clinical thermometer and to compare with surrounding temperature.
4	To calculate the Body Mass Index (BMI) of a person, by using the BMI formula and to compare the value with BMI chart.
5	To dissect and display the androecium and gynoecium of any locally available flowers.
6	To classify the fruits, separating the pericarps and writing the edible parts.
7	To identify the structure of ovule.
8	To prove the anaerobic respiration (Fermentation).
9	To find the pH of a given solution using pH paper.
10	To identify the presence of acids and bases in a given solution.
11	Preparation of true solution, colloidal solution and suspension.
12	To predict whether the reaction is exothermic or endothermic.
13	Screw Gauge-measuring small dimension.
14	Resistance of a coil of wire.
15	To map the magnetic field of a bar magnet when its north pole is pointing geographic north.
16.	Focal length of a convex lens by distance object method.
<p><i>Note : Students to record their findings in the appropriate table provided.</i></p> <p><i>No need to keep a separate observation or record note book.</i></p>	

Zoology

Ex. No : 1

Date :

To find out the presence of starch in the given food samples of A and B by using Iodine solution.

Aim:

To find out the presence of starch in the given food samples of A and B by using iodine solution.

Requirements:

Test tubes, Iodine solution.

Procedure:

Take 1 ml of food sample A and B in separate test tubes.

Add one drop of Iodine solution in both the test tubes.

Observe the colour change and record.

Indication : Appearance of dark blue colour indicates the presence of starch.

Table:

Sl.No	Food Sample	Observation	Presence / Absence of Starch
1	A		
2	B		

Result:

The food sample _____ contains starch.

Ex. No : 2

Date :

To find out the rate of heart beat of human beings by using stethoscope under normal physical conditions.

Aim:

To find out the rate of heart beat of a person by using stethoscope.

Requirements:

Stethoscope, stop watch.

Procedure:

Use the stethoscope and hear Lubb and Dubb sound which make up a heart beat.

Count the number of heart beats per minute and record.

Table:

Sl. No	Persons	No. of heart beat per minute
1	A	
2		
3		
4		
5		
Average :		

Inference:

Under normal conditions the average human heart beat is found to be _____ per minute.

Ex. No : 3

Date :

To find out the body temperature by using clinical thermometer and to compare with surrounding temperature.

Aim:

To find out the body temperature of human being using clinical thermometer.

Requirement:

Clinical thermometer, lab thermometer

Procedure:

Find out the room temperature by using lab thermometer.

Clean the Clinical thermometer in dilute dettol soaked cotton.

Shake the clinical thermometer at least four times.

Keep the mercury bulb of the clinical thermometer at the arm pit in boys or elbow in girls for a minute and record the temperature.

Repeat the same outside the room and record your findings for atleast three three of your friends.

Table:

S.No	Test	Room / External Temperature °F	Body Temperature °C	$C = \frac{F - 32}{9} \times 5$
1	Inside the room			
	Outside the room			
2	Inside the room			
	Outside the room			
3	Inside the room			
	Outside the room			

Inference:

Under normal conditions the body temperature of human beings is _____°F, _____°C.

The body temperature of human beings remains the or same/ varies irrespective of surroundings.

Ex. No : 4

Date :

To calculate the Body Mass Index (BMI) of a person, by using the BMI formula and comparing the value with BMI chart.

Aim:

To calculate the BMI of any one of your classmates by using the BMI formula.

Requirements:

Weighing machine, measuring tape.

Procedure:

Find out the weight in kg of your calssmate by using weighing machine.

Find out the height of the same person and convert into meter²

By using the formula

$$\text{BMI} = \frac{\text{weight in kg}}{\text{height in m}^2}$$

Find out the BMI and record.

Note:

BMI 19-25 is normal , 26 and above is obese, below 19 is lean.

Table:

Sl. No	Persons	weight in kg	Height in meter	Height in meter ²	BMI
1					
2					
3					

Inference:

The BMI of my classmate Selvan/Selvi is _____ and so he/she is normal/obese/ lean.

Botany

Ex. No : 5

Date :

To dissect and display the androecium and gynoecium of any locally available flowers.

To dissect and display the androecium and gynoecium of any locally available flowers.

Androecium

- 1) Androecium is the male reproductive part.
- 2) It has two parts, the filament and anther.
- 3) Pollen grains develop inside the anther.

Gynoecium

- 1) Gynoecium is the female reproductive part.
- 2) It has three parts, the ovary style and stigma.
- 3) Ovules are seen inside the ovary.

Separate the Androecium and Gynoecium of a given flower and paste in a separate sheet. Record your observations with regard to number of stamen shape of anther and shape of stigma in the given table.

Sl.no	Name of the flower	Androecium	Gynoecium
1.			
2.			
3.			
4.			
5.			

Ex. No : 6

Date :

To classify the fruits. Separate the pericarps and write the edible parts and fill in the blanks

Simple fleshy fruits

Berry - Tomato

- 1) The pericarp is divided into _____ and _____.
- 2) The mesocarp and endocarp remain _____.
- 3) The edible part in tomato is _____.

Berry - Banana

- 1) The pericarp is divided into _____ and _____.
- 2) The epicarp is _____ and the mesocarp is _____.
- 3) The edible part in banana is _____.

Hesperidium - Orange/Lemon.

- 1) The pericarp is differentiated into _____ layers.
- 2) The outer glandular skin is _____.
- 3) A middle thin whitish layer is _____.
- 4) An inner membranous part is _____.
- 5) The juicy hairs or out growths are _____.

Pepo - Cucumber/ivy gourd (Kovai)

- 1) The pericarp is _____ and _____.
- 2) The mesocarp is _____.
- 3) The edible part is _____.

Drupe – Mango

- 1) The number of seeds in mango is _____.
- 2) Pericarp is differentiated into epicarp, _____, _____ and _____.
- 3) Epicarp is s _____, mesocarp is _____ and endocarp is _____ in nature.
- 4) Edible part of the mango is _____.

Drupe – Coconut

- 1) The pericarp is differentiated into _____, _____ and _____.
- 2) The epicarp is thick, the mesocarp is _____, and the endocarp is hard.
- 3) The endosperm seen inside the _____ is edible.

Classify the given fruits, record your observations in the given table.

Sl. No.	Type of fruit of fruit	Nature of pericarp	Edible part
1.			
2.			
3.			

Ex. No : 7

Date :

To identify the structure of ovule.

The given slide kept for identification is L.S. of ovule

The characteristics of ovule :

- 1) The ovule has _____ layers of walls called as integuments.
- 2) Inner to the integuments, _____ is present.
- 3) The embryo sac has _____, _____ and _____

Observe the given slide and record your observations in the table :

Sl.No	Observation
1.	
2.	
3.	

Ex. No : 8

Date :

To prove the anaerobic respiration (fermentation).

Aim:

To prove the anaerobic respiration(fermentation)

Materials required:

Test tube, sugar solution, yeast.

Procedure:

Sugar solution in a test tube is taken. A little quantity of yeast is added.

The tube is placed in a warm place—sunlight.

Record your observations and inference in the table given below :

Observation	Inference

Results: The alcohol smell indicates that the sugar is converted into alcohol in the fermentation process

Chemistry

Ex. No : 9

Date :

To find the pH of a given solution using pH paper.

Aim:

To find the pH of the given solution using pH paper.

Materials and Apparatus Required:

Test tubes, test tube stand, test tube holder, pH paper, dil. HCl, dil. NaOH, lemon juice, water, baking soda solution, vinegar etc.

Procedure:

Take about 10 ml of the given samples in different test tubes and label them as A, B, C,D,E and F. Dip the pH paper into the test tubes and compare the colour of pH paper with the colour chart of pH reference. Note the approximate value of pH.

Table:

Test tubes	Sample	pH paper		Nature of solution
		Colour produced	Approximate pH	Acidic / Basic / Neutral
A				
B				
C				
D				
E				
F				

Ex. No :10

Date :

To identify the presence of acids and bases in a given solution

Aim

To identify the presence of an acid or a base in a given sample.

Materials and apparatus

Test tubes, test tube stand, glass rod, litmus paper (both red and blue), acids, bases, phenolphthalein, methyl orange solution.

Note:

- All acidic solutions are colourless in phenolphthalein, pink in methyl orange and turn blue litmus paper to red.
- All basic solutions are pink in phenolphthalein, straw yellow in methyl orange and turn red litmus paper to blue.

S.No	Experiment	Observation (Colour change)	Inference (Acid/base)
1	Take 5 ml of the test solution in a test tube, add phenolphthalein in drops to this content.		
2	Take 5 ml of the test solution in a test tube and add methyl orange in drops		
3	Take 10 ml of the test solution solution in a test tube and dip litmus paper into the test tube.		

Ex. No :11

Date :

Preparation of true solution, colloidal solution and suspension

Aim :

To prepare true solution, colloidal solution and suspension

Materials and apparatus required

Beakers, common salt, table sugar, starch, chalk powder, sand, egg albumin,

Table:

Experiment	Observation	Inference
Take 20ml of water in three different beakers and label them as A, B, C. Add common salt in A, starch in B, and chalk powder in C. Stir the contents of three different beakers gently. Record your observations.		

Note :

- i. If the particles do not settle down at the bottom and pass through the filter paper easily the solution is said to be a true solution.
- ii. If the particles do not settle down but they form turbid solution then the solution is said to be a colloidal solution.
- iii. If the particles settle down to form sediments leaving behind residue on the filter paper then the solution is said to be a suspension.

Result :

True solution is in beaker _____

Colloidal solution is in beaker _____

Suspension is in beaker _____

Ex. No :12

Date :

To predict whether the reaction is exothermic or endothermic.

Aim.

To predict whether a reaction is exothermic or endothermic using the given chemicals

Materials and apparatus required

Test tubes, test tube stand, water, glass rod, sodium hydroxide (pellets), ammonium chloride etc.

Note:

- **Exothermic reaction evolves heat**
- **Endothermic reaction absorbs heat**

S.No	Experiment	Observation(hot/cold)	Inference (exo/endo)
1	Take water in a test tube. Add sodium hydroxide pellets one by one followed by stirring. Touch the test tube and note the observation.		
2	Take water in a test tube. Add ammonium chloride salt and stir well .Touch the test tube and note the observation.		

PHYSICS

Ex. No :13

Date :

SCREW GAUGE - Measuring small dimensions of the object

Aim:

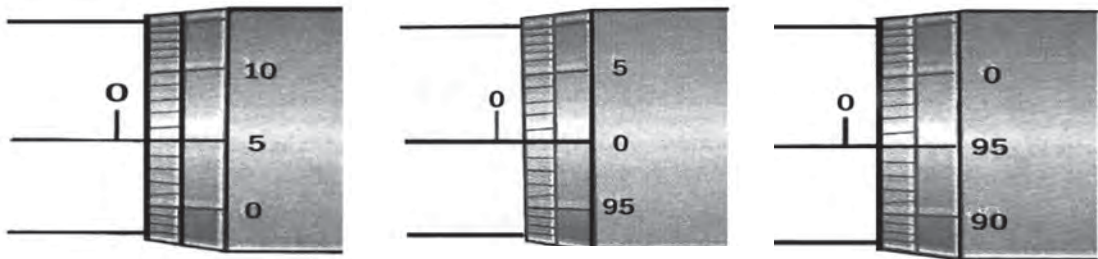
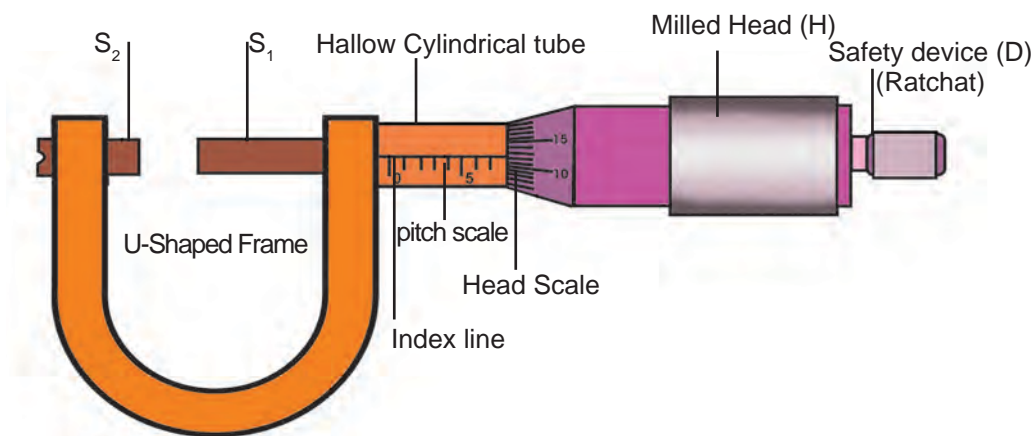
To find the radius of the given wire.

Apparatus required :

Screw gauge, a uniform thin metal wire.

Formula :

Radius of the wire $r = d/2$, d – diameter of the wire.



Procedure :

The least count of the screw gauge is found. Zero error of the screw gauge is found in the following way. The plane surface of the screw and the opposite

plane stud on the frame are brought into contact. If zero of head scale coincides with the pitch scale axis, there is no zero error.

If the zero of the head scale lies below the pitch scale axis, the zero error is positive. If the n^{th} division of the head scale coincides with the pitch scale axis

$$ZE = + (n \times LC)$$

Then the zero correction $ZC = - (n \times LC)$

If the zero of the head scale lies above the pitch scale axis, the zero error is negative. If the n^{th} division of the head scale coincides with the pitch scale axis

$$ZE = - (100 - n) \times LC$$

Then the zero correction $ZC = + (100 - n) \times LC$

Place the wire between two studs. Rotate the head until the wire is held firmly but not tightly. Note the pitch scale reading (PSR) and the head scale division which coincides with the pitch scale axis (HSC). The diameter of the wire is given by $PSR + (H.S.C \times LC) \pm ZC$. Repeat the experiment for different portions of the wire. Tabulate the readings. The average of the last column readings gives the diameter (d) of the wire. The value $d/2$ gives the radius of the wire.

Table:

L.C = **mm** **Z.E =** **mm** **Z.C =** **mm**

S.No	P.S.R (mm)	H.S.C	H.S.C × L.C (mm)	Total reading P.S.R +(H.S.C ×L.C) ± Z.C (mm)
1				
2				
3				
Mean =				

The radius of given wire $r = d/2$ mm

Result :

The radius of the given wire = mm

Ex. No :14

Date :

RESISTANCE OF A WIRE

Aim

To determine the resistance of the given wire .

Apparatus required

A battery(2 V), ammeter(1.5 A), voltmeter(1.5 V), key, rheostat, experimental wire(1 Ω or 2 Ω) and connecting wires.

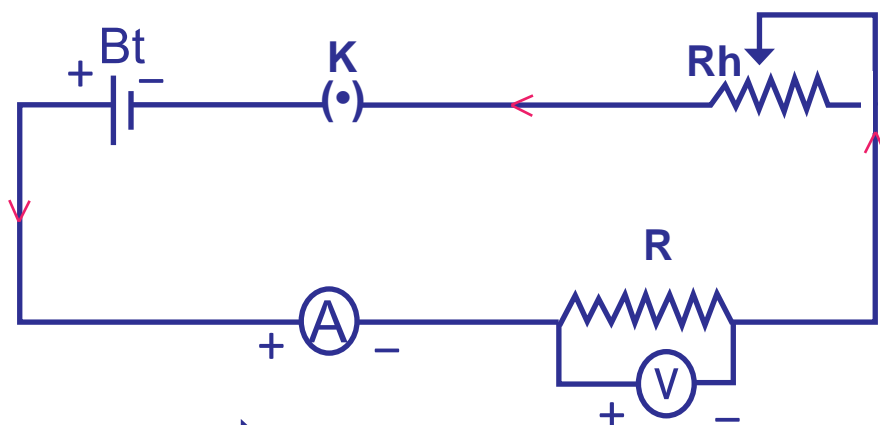
Formula

$$\text{Resistance of the wire } R = \frac{V}{I}$$

V is the potential difference across the wire.

I is the strength of the current through the experimental wire.

Circuit diagram:



Bt	➔	Battery
K	➔	Key
Rh	➔	Rheostat
A	➔	Ammeter
R	➔	Resistance (unknown)
V	➔	Voltmeter

Procedure

Connect the battery eliminator, ammeter, the given wire, rheostat and key in series. The voltmeter is connected in parallel connection across the given wire. The circuit is closed and the rheostat is adjusted such that a constant current flows through the given coil of wire. The current is noted as 'I' from the ammeter and the potential difference across the wire V is noted from the voltmeter. The value V/I gives the resistance of the wire. The experiment is repeated for different values of the current.

The average value of $\frac{V}{I}$ gives the resistance of the wire R.

Tabulation

Trial No	Ammeter reading I (ampere)	Voltmeter reading V (volt)	Resistance R = V/I (ohm)
1			
2			
3			
4			
5			

Mean R =

Result

Resistance of the given wire R = _____ ohm.

Ex. No :15

Date :

MAPPING OF MAGNETIC FIELD

Aim:

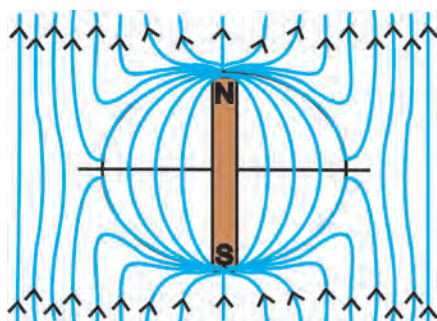
To map the magnetic lines of force when the bar magnet is placed with its north pole facing geographic north

Apparatus required:

Drawing board, drawing pins, bar magnet, small magnetic compass needle and white sheet.

Procedure:

A sheet of paper is fixed on a drawing board. Using a compass needle, the magnetic meridian is drawn on it. A bar magnet is placed on the magnetic meridian such that its north pole points towards geographic north. The north and south poles of the compass are marked by pencil dots. The compass needle is shifted and placed so that its south pole touches the pencil dot marked for the north pole. The process is repeated and a series of dots are obtained. The dots are joined as a smooth curve. This curve is a magnetic line of force. In the same way several magnetic lines of force are drawn around the magnet as shown in figure. The magnetic lines of force is due to the combined effect of the magnetic field due to bar magnet and the Earth.



Result:

The magnetic lines of force are mapped when the bar magnet is placed with its north pole facing geographic north. The mapped sheet is attached.

Ex. No :16

Date :

FOCAL LENGTH OF CONVEX LENS

Aim

To determine the focal length of convex lens by distant object method

Apparatus required

The given convex lens, lens stand, white screen and meter scale

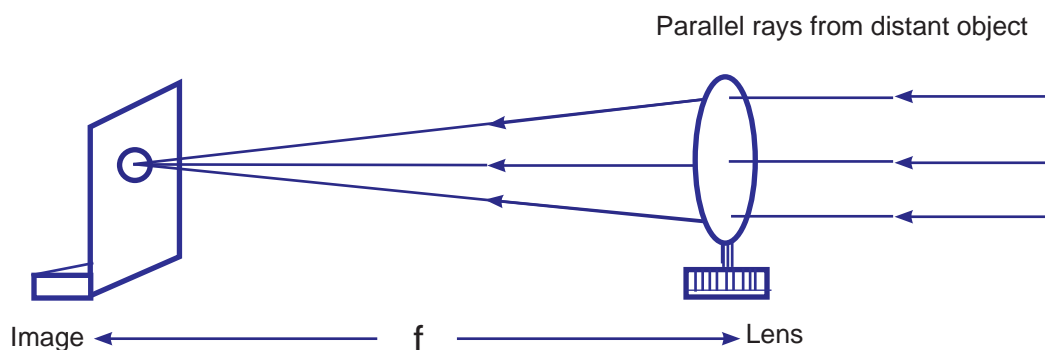
Procedure

$$\text{Formula : Focal length } f = \frac{f_1 + f_2 + f_3}{3}$$

f_1, f_2, f_3 , are the focal lengths measured by focusing different distant objects.

Distant object method

The convex lens is mounted on the stand and is kept facing a distant object (may be a tree or a building). The white screen is placed behind the convex lens and its position is adjusted to get a clear, diminished and inverted image of the object. The distance between the convex lens and the screen is measured. This gives an approximate value of the focal length of the convex lens.



S.No	Distant object	Distance between the convex lens and the screen
1	Tree	f_1
2	Building	f_2
3	Electric pole	f_3
Mean =		

Write down the observations in observation table.

Result:

Focal length of the given convex lens $f =$ _____cm